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Research Article

Keywords: environmental, hydrogen evolution reaction (HER)

Posted Date: November 25th, 2020

DOI: <https://doi.org/10.21203/rs.3.rs-109340/v1>

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Version of Record: A version of this preprint was published at Scientific Reports on February 8th, 2021.

See the published version at <https://doi.org/10.1038/s41598-021-83066-7>.

Metal Doped Layered MgB₂ Nanoparticles as Novel Electrocatalysts for Water Splitting

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ABSTRACT

Growing environmental problems along with the galloping rate of population growth have raised an unprecedented challenge to look for an ever-lasting alternative source of energy for fossil fuels. The eternal quest for sustainable energy production strategies has culminated in the electrocatalytic water splitting process integrated with renewable energy resources. The successful accomplishment of this process is thoroughly subject to highly efficient, earth-abundant, and cost-effective electrocatalysts to drive the oxygen evolution reaction (OER) and hydrogen evolution reaction (HER), preferably, in the same electrolyte. The present contribution has been dedicated to studying the synthesis, characterization, and electrochemical properties of newfangled electrocatalysts with the formal composition of Mg_{1-x}Tm_xB₂ ($x = 0.025, 0.05, \text{ and } 0.1$; Tm (transition metal) = Fe and Co) primarily in HER as well as OER under 1 M KOH medium. The electrochemical tests revealed that among all the metal-doped MgB₂ catalysts, Mg_{0.95}Co_{0.05}B₂ has the best HER performance showing an overpotential of 470 mV at -10 mA cm⁻² and a Tafel slope of 80 mV dec⁻¹ on account of its high purity and fast electron transport. Further investigation shed some light on the fact that Fe concentration and overpotential for HER have adverse relation meaning that the highest amount of Fe doping ($x = 0.1$) displayed the lowest overpotential. This contribution introduces not only highly competent electrocatalysts composed of low-cost precursors for the water-splitting process but also a facile scalable method for the assembly of highly porous electrodes paving the way for further stunning developments in the field.

Introduction

Disruption of the ecological balance by fossil fuels, their volatile prices, and the ever-growing energy demands have provided tremendous impetus to look for new, environmentally benign, abundant, and zero-emitting energy sources¹⁻⁴. Recent studies show that it will be unavoidable to launch and accelerate energy development from conventional energy systems to novel and viable alternatives⁵. Hydrogen, a potential energy carrier, has been addressed as the future energy owing to its higher energy density (~120 MJ/kg) compared to petroleum-based energy sources (~45 MJ/kg)^{4,6,7}, low weight, high abundance, and without leaving any carbon footprints of the combustion products⁸⁻¹⁰.

The current massive-scale production of hydrogen through steam reforming of natural gas and/or methanol consumes an enormous amount of energy and produces CO₂, making this process less viable from sustainability perspective.¹¹⁻¹³ Unconventional and environmentally-benign H₂ generation approaches encompass photo/electrolysis of water supported by renewable energy source (e.g., solar, geothermal)^{3,14,15}. In fact, in large-scale hydrogen production, the advantages of electrochemical water splitting outweigh those of other conventional systems, including coal gasification and steam reforming of methane, because it has enormous amount of sources and easy to use without CO₂ emission^{13,16}.

Since the invention of the voltaic pile, scientists have made plenty of attempts to ameliorate the approach of energy production from water¹⁷, but mass scale implementation based on photo/electrochemical water splitting processes is the major bottleneck mainly because of the inherently sluggish kinetics of OER and HER^{14,18,19}. Electrolysis of water is simply composed of two reactions, so-called, hydrogen evolution reaction (HER) and oxygen evolution reaction (OER). The reduction of H⁺ ions occurs at the cathode ($2\text{H}^+(\text{aq}) + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$) and the oxidation of water takes place at the anode ($2\text{H}_2\text{O}(\text{l}) \rightarrow \text{O}_2(\text{g}) + 4\text{H}^+(\text{aq}) + 4\text{e}^-$)^{5,20}. The thermodynamic voltage of water splitting is 1.23 V at 25 °C and 1 atm which corresponds to an energy input of $\Delta G^\circ = 237.1 \text{ kJ mol}^{-1}$ ^{12,21}, but much higher potentials—referred as overpotential—are required to initiate practical water splitting^{22,23}.

To minimize the overpotentials at the cathode and improve energy conversion efficiency, effective electrocatalysts must be employed^{3,24}. To date, Pt-based materials have been proven the most promising electrocatalysts for HER^{5,25,26}. Nevertheless, the shortage of sources on earth and high cost limit their widespread industrial applications^{20,26,27}. Thus, developing competent, stable, and low-cost catalytic materials for HER is crucial, but so far remains a great challenge^{28,29}.

In the last few years, plenty of non-precious transition metal compounds, namely sulfides (NiS₂, FeS₂, CoS₂), phosphides (Ni₂P, FeP, CoP), and selenides (CoSe₂, FeSe₂) have been found with excellent HER activities^{10,15,16}. Recently, transition metal (Ni, Co, Fe, etc)-based borides (TMBs) have sparked considerable attention in the fields of energy conversion and storage. An growing number of studies evince that some traits introduce TMBs as exceptional electrocatalysts for water splitting^{29–31}. As compared to other TM-based components, TMBs possess some special attributes such as high chemical, thermal, and mechanical stability, significant catalytic activity and conductivity for water splitting^{5,6,8,19,32,33}. In this sense, the results of several recent studies confirm the potential of transition metal borides as dual-functional catalysts for water splitting³⁴; yet, among TMBs, metal diborides (MDBs) have been poorly explored for this application.

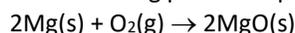
Metal diborides known by the general formula of MB₂, feature some intrinsic singular characteristics including hardness, chemical durability, high thermal conductivity, and high electrical conductivity^{35,36}. So far, MgB₂, as a representative of the large family of MDBs has been a well-known material for its superconductivity at low temperatures^{37–40}. MgB₂ has other attractive characteristics as regards practical use, such as the ability of doping effect, low cost and abundance of its raw materials, all of which makes it an auspicious candidate for doping and subsequently utilizing in water splitting³⁸. Moreover, a singular property of MgB₂ is the evacuation of B 2p-s bands on the basis of a unique Mg-B interaction. Accordingly, for any MDB, the correlation between boron and the metal atoms plays a decisive role on effectiveness of its properties leading to a firm premise for effective application in water electrolysis⁴¹.

Herein we examine and report a very successful approach to substitute the divalent transition metals in Mg_{1-x}Tm_xB₂ ($x = 0.025, 0.05, \text{ and } 0.1$; $Tm = \text{Fe and Co}$) for the Mg²⁺ as primarily enhanced HER electrocatalysts under alkaline medium. It should be pointed out that the aforementioned electrocatalysts have also been tested as OER catalysts and the data are provided in the supporting information.

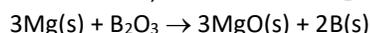
Results and Discussion

XRD patterns of pure MgB₂, Co-doped MgB₂, and Fe-doped MgB₂ are illustrated in Figure 1. For comparison, the theoretical diffraction pattern of magnesium diboride has also been included^{40,42}. Reflections shown in the theoretical pattern have appeared in both undoped MgB₂ as well as Mg_{1-x}Tm_xB₂ which confirms the successful synthesis technique applied in this study. However, there are trace amounts of labeled intrinsic impurities such as MgO (marked by #) at 2θ of 36.7°, 62.38° and Mg (marked by *) at 32.28°, 34.52°, 47.86°. Based on the literature⁴⁰, two main paths that lead to the formation of MgO are as follows.

- If Mg precursor powder is sintered under oxygen it will produce MgO according to the following reaction:



- Also, the reaction of B₂O₃ with Mg forms MgO based on the following reaction:



Because all the synthesis procedure was carried out in a glovebox, the probable occurrence of above-mentioned reactions can be attributed to the impurity of precursor materials. Therefore, it is ironically interesting that the raw materials of Mg powder have surface oxide to some extent that produces MgO which is utterly inevitable to prevent from happening.

The X-ray diffractograms of Co-doped MgB₂ showed that there was no sign of secondary phases containing Co that is an obvious indication of successful substitution of Co for Mg in the crystal structure (see Figure S1 and Table S1). Nevertheless, weak peaks corresponding to Mg and MgO are still present with much lower intensity compared to MgB₂.

The X-ray diffraction for the samples with formal composition Mg_{1-x}Fe_xB₂ revealed that Mg and MgO peaks were roughly negligible. However, there were secondary phases at 2θ of 32.54°, 37.7°, and 45.08° which were assigned to the FeB phase. It is remarkable to mention that the impurity peak at 45.08° (labeled by ●) can also be related to Fe₂B, especially at high concentration doping of Fe ($x = 0.1$)⁴³.

The minor amount of impurities with infinitesimal peaks indicated that the substitution of Co and Fe for Mg has significantly suppressed the amount of Mg and MgO formation. Moreover, with increasing Co and Fe content from $x = 0.025$ to 0.1, the amount of impurities decreases as well, leading to finer structures that is obviously as a result of lowering the amount of Mg in the composition. Altogether, Mg_{0.95}Co_{0.05}B₂ contains the least amount of impurities.

The full-width-at-half-maximum (FWHM) values for synthesized catalysts were calculated by Program STOE WinXPOW (Darmstadt: Stoe & Cie GmbH, 2016) and subsequently crystallite size computed and summarized in Table 1. It is noteworthy to state that the FWHM of the XRD (101) peak has direct relation with the concentration of Fe and Co dopants. This demonstrates that nano Fe- and Co-doped samples contain tiny grains and imperfect crystallinity. The Scherrer's equation is applied to calculate the average particle size (D) of the nanoparticles for each sample (peak 101)

$$D = (k \lambda) / (\beta \cos\theta)$$

where $k = 0.94$ is the shape factor of a particle, presuming shape of the nanoparticles as sphere, $\lambda = 1.5406 \text{ \AA}$ the wavelength of Cu $K\alpha$ radiation, β is the FWHM of the XRD (101) peak, and θ is the diffraction angle of the peak⁴³. A glance at the table points out that the average grain size of the parent MgB_2 is about 29 nm. These results based on the calculation of the grain size of doped samples unveils that doping concentration and grain size have reverse relation. By increasing doping amounts in both Fe and Co compositions, grain size shrinks.

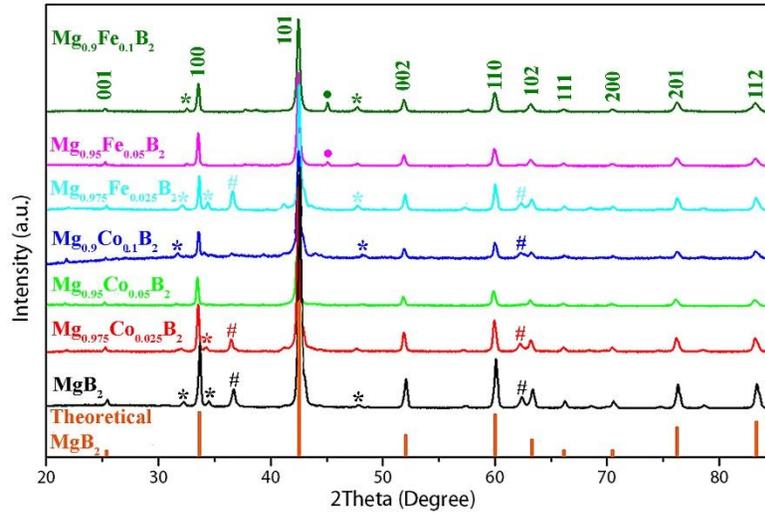


Figure 1. XRD patterns of MgB_2 and $\text{Mg}_{1-x}\text{Tm}_x\text{B}_2$ ($x = 0.025, 0.05,$ and 0.1 ; $\text{Tm} = \text{Fe}$ and Co): impurity peaks of Mg, MgO, and FeB/Fe₂B are marked by *, #, and •, respectively.

Regarding previous experimental and theoretical⁴⁴ results, MgB_2 crystallizes in the well-known AlB_2 -type structure with $P6/mmm$ symmetry. As can be seen from Figure 2, Mg layers are positioned between two honeycomb layers of B atoms. This structure resembles graphite regarded as that of completely intercalated graphite with carbon substituted by boron. Notwithstanding similarities, the B–B bonding is strongly anisotropic for the reason that intralayer B–B bonds are much shorter than the interlayer distance—which is not observed in graphite⁴⁴.

Sample	FWHM (101)	D (nm)
MgB_2	0.3087	28.82945
$\text{Mg}_{0.975}\text{Fe}_{0.025}\text{B}_2$	0.2233	39.85514
$\text{Mg}_{0.95}\text{Fe}_{0.05}\text{B}_2$	0.2447	36.36965
$\text{Mg}_{0.9}\text{Fe}_{0.1}\text{B}_2$	0.2909	30.59351
$\text{Mg}_{0.975}\text{Co}_{0.025}\text{B}_2$	0.2232	39.87299
$\text{Mg}_{0.95}\text{Co}_{0.05}\text{B}_2$	0.2298	38.72782
$\text{Mg}_{0.9}\text{Co}_{0.1}\text{B}_2$	0.3062	29.06483

Table 1. FWHM values and crystallite size for pure $\text{Mg}_{1-x}\text{Tm}_x\text{B}_2$ ($x = 0.025, 0.05,$ and 0.1 ; $\text{Tm} = \text{Fe}$ and Co).

Partial filling of two σ bands associated with strongly covalent sp^2 -hybrid bonding within the 2D boron layers is another attractive feature of MgB_2 . The crystal structure manifests that the holes at the top of these σ bands possess mainly two-dimensional properties and are restricted within the boron sheets. On the contrary, electrons and holes in the π bands mostly represent three-dimensional characteristic which are distributed through the whole crystal. Because of the extensive overlapping that takes place between all p orbitals (in-plane as well as out-of-plane) for adjacent boron atoms, σ and π bands encompass strong in-plane dispersion. However, the interlayer overlap, particularly for p_{xy} orbitals, are not large enough in such a way that the k_z dispersion of σ bands does not surpass 1 eV. It should be noted that unfilled σ bands in conjunction with boron p_{xy} character sufficiently maintain the covalent structure throughout the whole crystal⁴⁵. Conducting covalent bonds stand for a distinctive feature of MgB_2 turning it into an exceptional material probably existing at the threshold of structural instability⁴⁴.

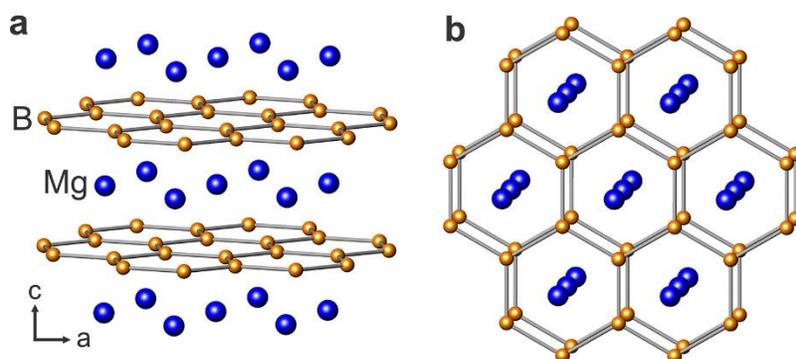


Figure 2. Schematic design of layered MgB_2 structure. a) along the a-axis, b) along the c-axis.

Figure 3 and S2 show XPS studies of pure MgB_2 , Co-doped MgB_2 , as well as Fe-doped MgB_2 nanoparticles. For this, Mg 1s, B 1s, Co 2p, and Fe 2p spectra have been analyzed. Besides, O 1s spectra have also been considered explaining the trace amount of surface impurities such as MgO. Regarding the fact that the Ar atmosphere was applied in all synthesizing steps, the reaction of B_2O_3 with Mg (as mentioned in the previous section) is inevitable. So, MgB_2 is somewhat vulnerable to oxidation as observed in Figure 3 and S2 O 1s spectra.

Fitting and deconvolution of Mg 1s spectra of pure MgB_2 , reveals a peak at 1303.7 eV. Based on the literature⁴⁶, metallic Mg shows a peak of around 1303 eV. This positive shift of Mg 1s toward higher binding energies in MgB_2 is related to donating electrons to boron; thereby, leading to stronger bonding in MgB_2 compared to metallic Mg. Besides, in MgB_2 , there is probability of some reverse charge transfer from B to Mg (+2) leading to a diminution of the formal charge of Mg and hence the partial covalency in Mg-B bond. This charge transfer results in the fragmentary covalency in the Mg-B bond. As written above, in MgB_2 structure—an example of boron-rich system—the electrons transfer from Mg to B layers²⁹. The electrons move to B π -bands as the σ -bands are filled, Then, some electrons shift from B π to Mg (+2). Consequently, the charge state of Mg is reduced from (+2) to lower states⁴⁵.

A general trend can be observed in Figure 3 and S2 that doping shifts Mg 1s toward higher binding energies. Specifically, in the low doping amount of Co and Fe ($x = 0.025$), Mg 1s peak shifts to even higher peaks, 1305.5 eV (refer to Figure S2). There are differences of > 1 eV in binding energies between doped and undoped samples. These large differences are ascribed to the effect of neighboring atoms. As stated by⁴⁶, non-exchangeable and exchangeable Mg ions exhibit peaks around 1303 and 1306 eV, respectively. In low amounts of Co and Fe dopant ($x = 0.025$), Mg ions act as exchangeable ions, whereas, in higher contents of Co and Fe ($x = 0.05$ and 0.1), Mg ions are considered as non-exchangeable ones. For this case, the flow of electronic charge to Mg ion, and the extra-atomic relaxation from surrounding oxygen atoms are larger.

Figure 3 and S2 depict the examined B 1s spectra. B 1s spectra of pure MgB_2 display peaks around 185.2, 186.4, 188.8, and 191.3 pertaining to B(OX), MgB_2 (boron as boride anion), B as metallic and B_2O_3 , respectively. These are consistent with the binding energy values reported for Mg diboride⁴⁵. MgB_2 is viewed as $Mg^{2+}(B_2)^{2-}$ with an average charge of (-1) at B. As it was mentioned previously, in MgB_2 structure, Mg charge is less than (+2), so the average charge on B is less than (-1). Doping of MgB_2 with metallic elements such as Co and Fe shifts binding energies of B to higher values, especially in the low amount of dopant as $x = 0.025$. This slight increase of binding energy stipulates the partial electron transfer from B to the vacant d-orbital of the metallic dopant.

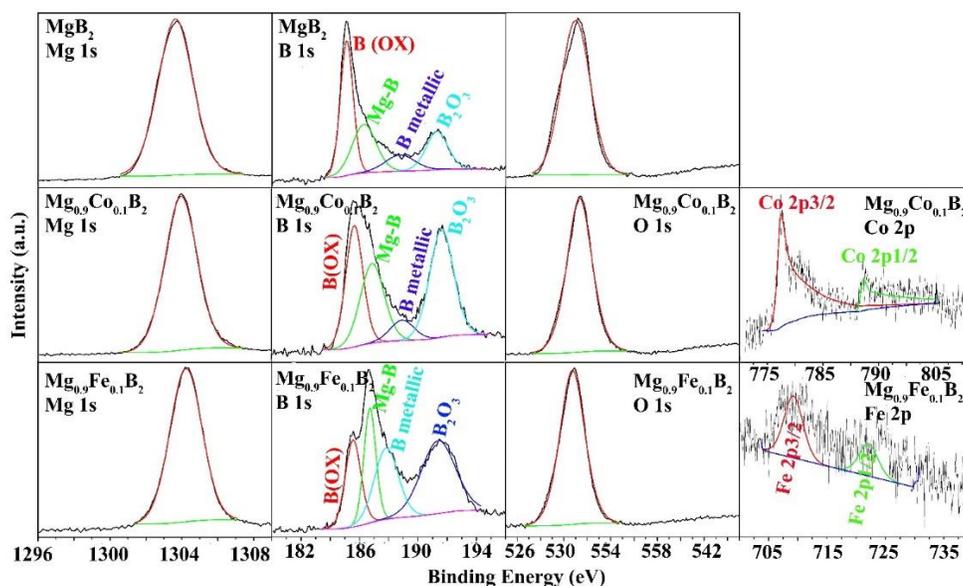


Figure 3. XPS spectra of Mg 1s, B 1s, O 1s, Co 2p, and Fe 2p in MgB_2 , $\text{Mg}_{0.9}\text{Co}_{0.1}\text{B}_2$, and $\text{Mg}_{0.9}\text{Fe}_{0.1}\text{B}_2$.

To investigate the oxygen-containing compounds at the surface, the O 1s spectra have also been illustrated in Figure 3 and S2. The prominent peak around 531 eV corresponds to MgO and surface oxygen possibility. Finally, Co 2p and Fe 2p peaks were observed in all doped samples, confirming the presence of Co and Fe species in the MgB_2 structure. For Fe-doped $\text{Mg}_{1-x}\text{Fe}_x\text{B}_2$ ($x = 0.025$), Fe 2p_{3/2} peak is also possible, but it could not be deconvoluted due to its low intensity.

Generally, the particle morphology studied by FESEM images (Figure 4 and S3) disclosed that for pure MgB_2 , nanoparticles sizes are in the range of 40–300 nm, Fe-doped MgB_2 29–500 nm, and Co-doped MgB_2 39–450 nm. Needless to mention that larger particles with sizes as big as 1–2 micrometers were also detected that can be primarily because of agglomeration. The morphology of resulting particles demonstrated that pure MgB_2 has been mostly dominated by spherical shape particles. Nonetheless, samples with a formal composition of $\text{Mg}_{1-x}\text{Fe}_x\text{B}_2$ displayed quite distinctive particle shapes (e.g., needle-like) that can be ascribed to interaction between metallic and boron layers as well as high pressure applied during their synthesis procedure. Comparing the individual crystallite of Fe and Co substituted MgB_2 signifies the fact that Co-doped samples have slightly better incorporation of Co into the MgB_2 structure. On the contrary, Fe-doped samples lack this benefit most probably due to the formation of secondary phases such as FeB or Fe_2B . The mentioned conclusions were also witnessed in ref. ⁴⁷.

According to the previous literature ^{8,10,21,48}, catalysts with attributes such as large specific surface area, open porous structure, and high accessibility of active sites are expected to exhibit outstanding electrocatalytic performance. Thanks to the FESEM image of the surface of the working electrode (Figure 5), uneven morphology, pores, and cracks can be observed that facilitates the accessibility of the electrolyte molecules into the far deeper parts of the electrode: thereby, drawing on remarkable catalytic activity.

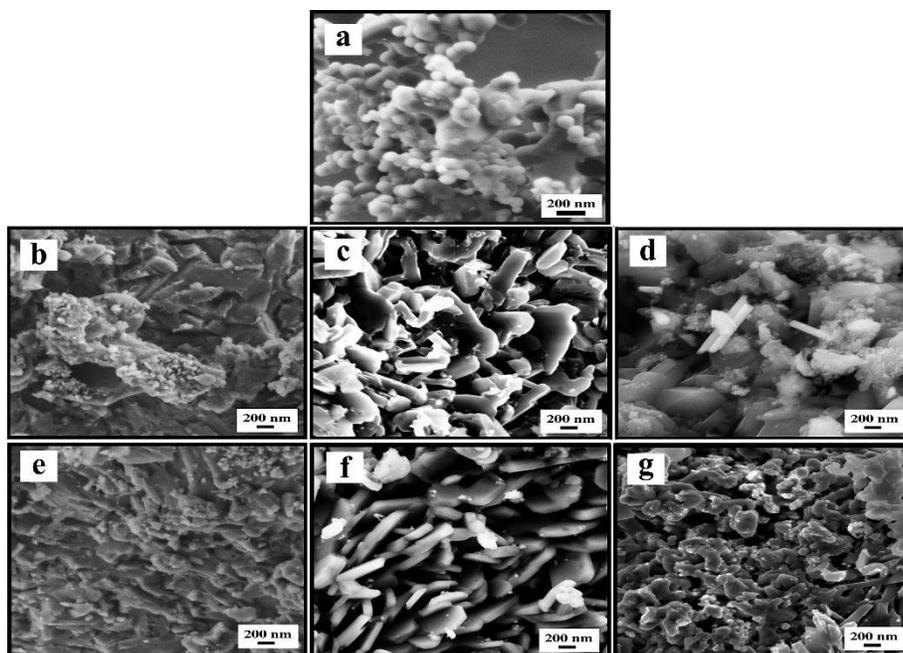


Figure 4. SEM images of a) pure MgB_2 , b) $\text{Mg}_{0.975}\text{Co}_{0.025}\text{B}_2$, c) $\text{Mg}_{0.95}\text{Co}_{0.05}\text{B}_2$, d) $\text{Mg}_{0.9}\text{Co}_{0.1}\text{B}_2$, e) $\text{Mg}_{0.975}\text{Fe}_{0.025}\text{B}_2$, f) $\text{Mg}_{0.95}\text{Fe}_{0.05}\text{B}_2$, and g) $\text{Mg}_{0.9}\text{Fe}_{0.1}\text{B}_2$.

Electrocatalytic features of undoped MgB_2 and transition metal doped MgB_2 were evaluated using electrochemical tests. Both HER and OER were examined carefully. The recorded voltammograms of the OER are given in Figure S4. As it is noticeable from this figure, $\text{Mg}_{0.9}\text{Fe}_{0.1}\text{B}_2$ exhibits the superior reactivity amongst all the doped MgB_2 samples. This decent performance should be assigned to the incorporation of Fe atoms into the MgB_2 structure. In addition, as reported by Li et al.⁴⁹ on boride-based electrodes, Fe dopants can enhance the local electronic structure owing to its superconductivity, resulting in enhanced OER performance.

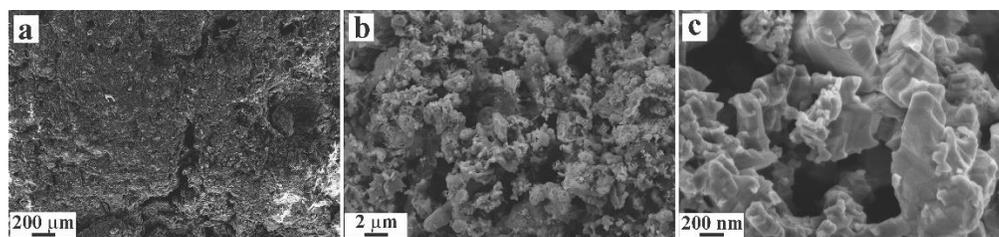


Figure 5. SEM images of the working electrode surface.

The voltammograms of the HER polarization curves are depicted in Figure 6a. For this, LSV electrochemical tests were recorded over a range of potential from -0.2 to -1.6 V under the basic 1 M KOH medium. It is striking that doping of MgB_2 with divalent metals affects the HER reaction significantly, as in doped electrodes, the overpotential is extremely lower at -10 mA cm^{-2} current density compared to that of pure MgB_2 (see Figure 6a). In this regard, $\text{Mg}_{0.95}\text{Co}_{0.05}\text{B}_2$ showed the best results in terms of the HER performance with the lowest overpotential at -10 mA cm^{-2} current density. This suggests that the HER reaction over $\text{Mg}_{0.95}\text{Co}_{0.05}\text{B}_2$ proceeds in an efficient manner which is in accordance with the literature—in that—Co-based borides are favorable catalysts in the water reduction reactions²¹.

As reported by Rafieezad et al.⁴⁰, the presence of MgO impurity can impede the connection between the grains and thereby, undermine the superconductivity of MgB_2 . Moreover, it has been emphasized that coverage of MgB_2 grains by the MgO layer can diminish the critical current density considerably⁴⁰. With these in mind, $\text{Mg}_{0.95}\text{Co}_{0.05}\text{B}_2$ and $\text{Mg}_{0.9}\text{Fe}_{0.1}\text{B}_2$ displayed the highest HER performance respectively on account of the least amount of impurity in terms of MgO (refer to Figure 1).

As it was mentioned previously, TMBs possess remarkable features such as reverse electron transfer from boron to metals resulting in fast electrochemical reactions, surface oxidation of TMBs on the account of the formation of active chemicals such borates, boron oxides, TM oxides, TM (oxy)hydroxides leading to effective catalytic reactions, and heat-treatment of TMBs that can promote catalytic performance as a result of phase change⁵. In addition to these aspects, the effect of hetero-metal doping on TMBs has also been widely reviewed. In this context, it was found that metal doping can expand the specific surface area (SSA) through inhibition of the agglomeration that reflects in the exposure of more available active sites, engineering the crystal structures, and tuning the electronic structures^{5,33,50}.

It was established that in boride-rich compounds such as FeB_2 , Fe as cation holds a partial positive charge, and B as anion carries a partial negative charge. Here, electron transfer from Fe to B expedites the adsorption of OH^- on positively charged Fe sites and H^+ on negatively charged B sites⁴⁹. Regarding MgB_2 as parent electrocatalyst utilized in this study, Fe and Co dopants—more electronegative compared to Mg—have changed surface morphology of the samples, regulated electronic structures, and induced fairly higher electron density transferred from cation site to anion, facilitating both HER and OER reactions.

For better interpretation, EIS analyses were performed in open circuit potential to predict charge transfer resistance through the electrodes. Figure 6b and c present the Nyquist plots of all the electrodes in full and zoom screen, respectively. As revealed by FESEM images from the electrode surface (Figure 5), the fabricated electrodes possess high porosity. To this end, the impedances were modeled using a so-called two-CPE model, $R(R\text{CPE})(R\text{CPE})$, composed of a solution resistance and a connection between two paralleled R-CPE elements in series—the high-frequency R-CPE element is correlated with the porosity whereas the low-frequency element is related to the kinetics of the HER.

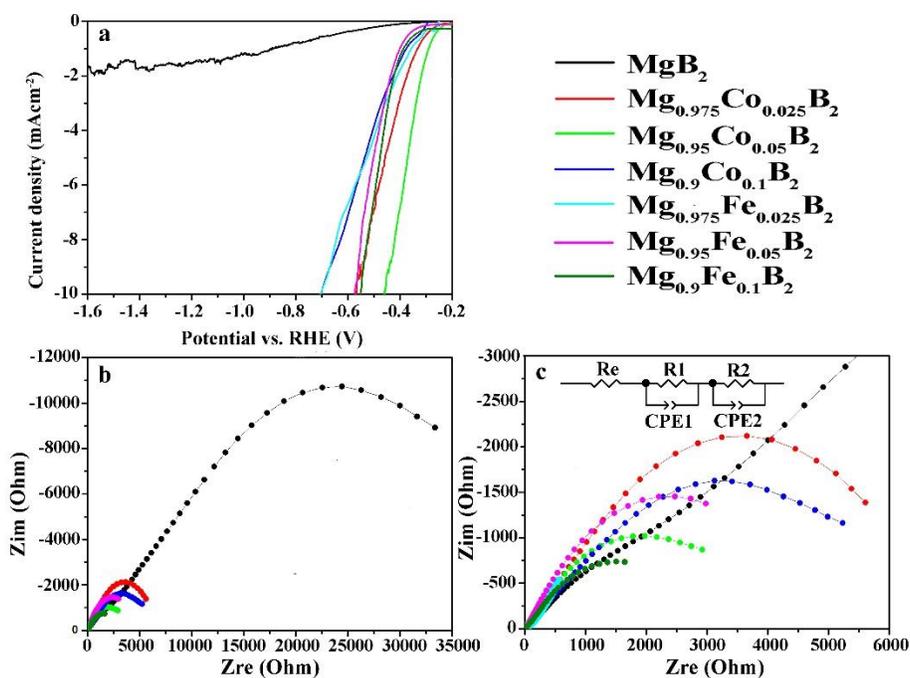


Figure 6. Polarization curves of a) the HER reaction of pure and doped MgB_2 , b) and c) Nyquist plots measured at OCP in supporting electrolyte of 1 M KOH.

In principle, the formation of two semicircles on the Nyquist plot corresponds to the geometric kinetic origin. The first semicircle is related to geometric and the second one is related to kinetic. The diameter of the semicircles represents the charge transfer resistance. Both the transfer of electronic charge on the electrode/electrolyte interface and the charge transfer in the host lattice control the electrochemical reactions at the electrodes. The acquired results for the Nyquist equation by Z-view software are summarized in Table 2 (detailed data in Table S2). The results support that metal doping of MgB_2 has dramatically decreased both R1 (charge-transfer resistance of the electrode/electrolyte interface) and R2 (charge-transfer resistance of the host lattice), and therefore increased the electrochemical catalytic activity.

As Mg is a highly reactive metal, it can readily get oxidized and converted into Mg(OH)₂ on the electrode surface in the KOH solution. The very plain interpretation would be, the formation of Mg(OH)₂ on the surface can block available active sites, increase the reaction resistance and restrain the diffusion of H atoms into the inner phase. Taken this interpretation into account, the high amount of R1 (10.993 kΩ cm⁻²) in undoped MgB₂ explains itself. On the contrary, the substitution of Co²⁺ and Fe²⁺ for Mg²⁺ in the alloy reduces the formation of Mg(OH)₂ and improves the life cycle of the electrode. Additionally, doping amplifies the electron mobility and conductivity of the electrodes, yielding much lower amounts of R2 for the doped electrocatalysts (see Table 2). Overall, Mg_{0.9}Fe_{0.1}B₂ (7.980 kΩ cm⁻²) and Mg_{0.95}Co_{0.05}B₂ (10.409 kΩ cm⁻²) have the lowest charge-transfer resistance in terms of R2 among the doped electrodes; as a result, they outperform all others in the sense of electrocatalytic performance.

Sample	Re (Ωcm ⁻²)	R1 (kΩ cm ⁻²)	R2 (kΩ cm ⁻²)
MgB ₂	130.5	10.993	127.114
Mg _{0.975} Co _{0.025} B ₂	12.2	1.436	18.331
Mg _{0.95} Co _{0.05} B ₂	13.6	0.409	10.409
Mg _{0.9} Co _{0.1} B ₂	13.2	3.539	13.712
Mg _{0.975} Fe _{0.025} B ₂	7.52	0.454	710
Mg _{0.95} Fe _{0.05} B ₂	13.5	0.277	14.325
Mg _{0.9} Fe _{0.1} B ₂	11.9	0.332	7.980

Table 2. Equivalent-circuit element values for EIS data in the 1 M KOH electrolyte.

To assess the probable degradation of the electrocatalysts after electrochemical reactions, FESEM coupled with EDX analyses were conducted on the samples. Figure S5 presents the EDX results of the samples after electrochemical reactions at different spots and the outcomes have been reported with and without oxygen content. Based on semi-quantitative analysis results in Tables S1 and S3, it can be observed that there are slight changes in the atomic percent of all elements after HER and OER, implying that the fabricated electrodes have remained largely intact. Considering that reactions have taken place in 1 M KOH, increasing the atomic percent of oxygen on the surface of both best-performing electrodes from 3.63 to 10.59 % and 3.12 to 14.93 % corresponding to Mg_{0.95}Co_{0.05}B₂ and Mg_{0.9}Fe_{0.1}B₂ is not surprising. It is assumed that the surface metal hydroxide/oxyhydroxide produced in the course of OER is the major ground for the high water oxidation performance⁴⁹. As a result, high oxygen contents for the post-OER electrodes can be ascribed to Mg-, Fe-, and Co- hydroxide/oxyhydroxide. In addition to this, higher alterations in the atomic percent of metals—Mg, Fe, and Co—can be witnessed compared to nonmetal B. For the case of Mg, it should be noted that inevitably unreacted Mg before HER and OER reactions have been consumed as Mg- hydroxide/oxyhydroxide leading to less Mg contents after electrochemical reactions for both electrodes. For the practical applications in the electrocatalytic process, overpotential should be kept to the lowest amounts to deliver a benchmark current density of 10 mA cm⁻² during the process. For this reason, metal doping is not only an efficient way to improve the sluggish kinetics of HER but also is beneficial to lower the kinetic energy resistances of electrocatalytic reaction. The overpotentials of Co- and Fe-doped MgB₂ at -10 mA cm⁻² are shown in Figure 7a indicating the HER performance of the electrodes. Mg_{0.95}Co_{0.05}B₂ has a low overpotential of 470 mV at -10 mA cm⁻² amidst all the investigated samples owing to its high purity. However, the other Co-doped electrodes did not follow any specific activity trend. In contrast, the activity trend in Fe-doped electrodes showed an inverse relation with Fe content. By increasing Fe amount, overpotential diminished, giving rise to the following trend Mg_{0.9}Fe_{0.1}B₂ < Mg_{0.95}Fe_{0.05}B₂ < Mg_{0.975}Fe_{0.025}B₂.

Concerning the fact that there has been a limited number of reports on layered MDbs in HER, a brief comparison of overpotentials with the present work and recently reported boride based electrocatalysts would be intriguing. For this, Mazánek et al¹. stated HER performance of several layered MDbs including MgB₂ in 0.5 M H₂SO₄. According to this report, the best performing MDb was ZrB₂ with an overpotential of over 950 mV at a current density of -10 mA cm⁻². Moreover, an elaborate review was recently published by Gupta et al¹⁷. Here, a few results on overpotential of metal borides in 1 M KOH

solution are given: Mo–B \approx 240 mV (20 mA cm⁻²), NiB_x film \approx 135 mV (10 mA cm⁻²), Ni–B \approx 125 mV (20 mA cm⁻²), Co₂B–CoSe₂ \approx 300 mV (10 mA cm⁻²), Co–Ni–B \approx 205 mV (10 mA cm⁻²), and FeB₂ \approx 61 mV (10 mA cm⁻²). Even though at first glance the investigated electrocatalysts in this study appear less competent compared to the above-written non-precious electrocatalysts, it should be overemphasized that the present results emerged with trace amounts of dopants and fairly economical precursors, showing a great promise for future works.

Figure 7b presents Tafel curves for all the synthesized electrodes. The Tafel slope is a crucial indication of reaction kinetics and rate in the HER process⁵¹. In other words, it is an intrinsic factor for the assessment of the rate-determining steps in the course of HER⁵². According to the previous studies^{18,53,54}, two general paths are proposed for hydrogen evolution in alkaline solutions; either based on Volmer–Heyrovsky mechanism (a and b) or the Volmer-Tafel mechanism (c and d):

- $\text{H}_2\text{O}(1) + \text{e}^- + * \rightarrow \text{H}^* + \text{OH}^-(\text{aq})$, Volmer reaction
- $\text{H}^* + \text{H}_2\text{O}(1) + \text{e}^- \rightarrow \text{H}_2(\text{g}) + \text{OH}^-(\text{aq}) + *$, Heyrovsky reaction
- $\text{H}_2\text{O}(1) + \text{e}^- + * \rightarrow \text{H}^* + \text{OH}^-(\text{aq})$, Volmer reaction
- $\text{H}^* + \text{H}^* \rightarrow \text{H}_2(\text{g}) + 2*$, Tafel reaction

In the foregoing reactions, * represents available active sites and H* hydrogen atoms bound to active sites. Moreover, it has been estimated that the required Tafel slopes for Volmer reaction, Heyrovsky reaction, and Tafel reaction are \sim 120 mV dec⁻¹, \sim 40 mV dec⁻¹, and \sim 30 mV dec⁻¹, respectively¹⁸. Regarding the Tafel slopes shown in Figure 7b, Mg_{0.95}Co_{0.05}B₂ has the smallest Tafel slope, 80 mV dec⁻¹ which indicates that it falls within the range of 40-120 mV dec⁻¹, demonstrating the Volmer–Heyrovsky mechanism. This superior performance accords with overpotential results as well, suggesting facile kinetics and faster electron transport^{18,52,55,56} on Mg_{0.95}Co_{0.05}B₂ as compared to all other samples. Interestingly, Mg_{0.975}Co_{0.025}B₂ with a Tafel slope of 112 mV dec⁻¹, Mg_{0.95}Fe_{0.05}B₂ with that of 117 mV dec⁻¹, and Mg_{0.9}Fe_{0.1}B₂ with a slope of 131 mV dec⁻¹ have relatively comparable catalytic activity, proceeding via Volmer mechanism. Other doped samples that displayed rather higher Tafel slopes likely suffered from insufficient conductivity and low accessible active sites^{52,55}.

Despite the fact that too high the Tafel slope necessitates the electrocatalyst to provide excessive overpotential to meet the required current density⁵⁴, we cannot draw a clear conclusion based on the relation between overpotential and Tafel slope. Nonetheless, a meaningful point is revealed—that is—at low overpotential (Mg_{0.95}Co_{0.05}B₂, Figure 7a) the Volmer–Heyrovsky reaction is dominated, while at high overpotential (Fe-doped samples, Figure 7a) the Volmer–Tafel reaction prevails. The preceding remark is substantiated by the literature as well⁵³.

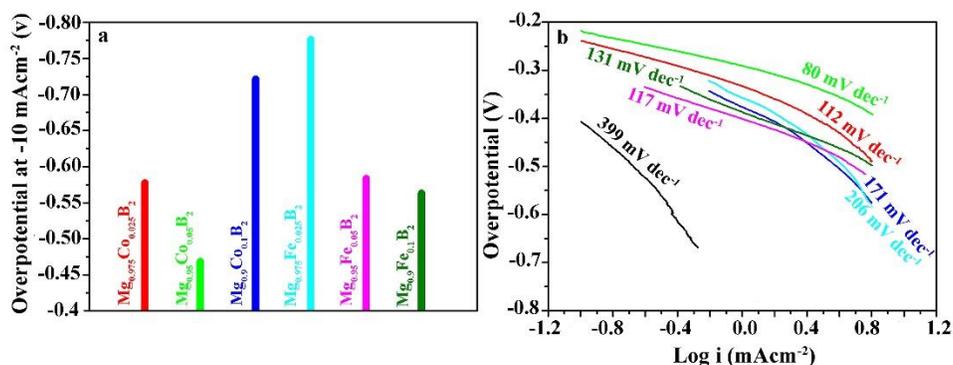


Figure 7. Fundamental electrochemical properties for HER. a) Overpotential at -10 mA cm^{-2} of metal-doped MgB₂, b) The corresponding Tafel plots of undoped and doped MgB₂. The colors indicating the same samples pictured in Figure 6.

The stability tests of the Mg_{0.95}Co_{0.05}B₂ and Mg_{0.9}Fe_{0.1}B₂ electrodes were performed by the cyclic voltammetry (CV) scanning method. Specifically, the current density was measured over a scanning rate of 50 mV s⁻¹ in a voltage window of -1 V to -0.2 V (E vs. RHE) for HER under 1 M KOH solution. (see Figure S6). Altogether, for HER stability tests, even though the current density did not decrease appreciably, yet further investigations are needed to excel in the stability of these very new electrocatalysts.

Conclusion

In this work, we performed a straightforward scalable method to synthesize a series of novel metal-doped MgB₂ with earth-abundant elements, Fe and Co, with a loading concentration of 0.025, 0.05, and 0.1. The as-prepared materials anchored on

a copper foil were employed as electrodes for the water-splitting process to function as HER electrocatalysts. Meanwhile, they were examined as OER catalysts as well. The overall outcomes substantiated that introduction of Co and Fe into the structure of layered MgB_2 meaningfully enhanced both OER and HER performance. In this way, a remarkable improvement of OER activity for Fe-doped MgB_2 electrodes was ascribed to the dramatic reduction of charge-transfer resistance both at electrode/electrolyte surface and within the host lattice. As a result, increasing Fe content gave rise to outstanding OER activity corresponding to $\text{Mg}_{0.9}\text{Fe}_{0.1}\text{B}_2$, by virtue of its optimum values of charge-transfer resistance. Besides, the best-performing catalyst in terms of HER was obtained for $\text{Mg}_{0.95}\text{Co}_{0.05}\text{B}_2$ displaying an overpotential of 470 mV at -10 mA cm^{-2} and a small Tafel slope of 80 mV dec^{-1} . This notable performance arose from several factors including the efficacious integration of Co in the crystal structure, facile kinetics, and fast electron transport that resulted in a drastic abatement of charge-transfer resistance in the host lattice compared to the undoped sample ($10.409 \text{ vs. } 127.114 \text{ k}\Omega \text{ cm}^{-2}$). This comprehensive electrochemical study conducted on Fe and Co-doped MgB_2 may provide further insights into designing new earth-abundant and efficient materials that rival and exceed the performance of conventional electrocatalysts.

Methodology

Synthesis of Electrocatalysts

Preparation of MgB_2

In general, there are two methods to synthesize magnesium diboride, namely “in situ” and “ex-situ” techniques. In this study, the former method has been preferred to the latter due to the conspicuous benefits. In the “in situ” technique, precursor materials are mixed with stoichiometric compositions and annealed under the inert condition. Compared to the “ex-situ” technique, “in situ” is more effective for doping purposes, since already reacted powders of MgB_2 are used in the “ex-situ” technique which does not leave enough room to develop metal doping⁴⁰. Further, according to the literature⁴⁰, “ex-situ” fabrication of MgB_2 is usually subjected to poor grain conductivity which hinders its application in electrolysis. In brief, the synthesis of pure MgB_2 powders was carried out via the “in situ” process. A specified amount of magnesium powder (Mg, 99.8%; Alfa Aesar) and amorphous nano boron (B, 98.5%; Pavezyum) were mixed in stoichiometric amounts in a glove box under Ar atmosphere and transferred into a steel tube. Afterward, the steel tube was sealed by arc melting under Ar. Then, the annealing process was applied at $850 \text{ }^\circ\text{C}$ for 6 h. It should be mentioned that following the annealing process, powders were ground in an agate mortar to obtain fine and homogenous powders for the electrochemical measurements.

Preparation of $\text{Mg}_{1-x}\text{Tm}_x\text{B}_2$ ($x = 0.025, 0.05, \text{ and } 0.1$; Tm = Fe and Co)

For the substitution experiments, cobalt powder (Co, 99.8%; Alfa Aesar), and reduced fine powder of iron (Fe, $\geq 99\%$, Sigma-Aldrich), were utilized individually. The powders; Mg, Fe (or Co), and B with the formal composition of $\text{Mg}_{1-x}\text{Tm}_x\text{B}_2$ were homogenized by grinding in an agate mortar and then compacted into pellets with a diameter of 10 mm at a pressure of about 10 ton. The pellets were then sealed inside a steel tube and went under heat-treatment at $850 \text{ }^\circ\text{C}$ for 6 h. The post-annealing process is the same as bare MgB_2 ^{57,58} (refer to Figure 8).

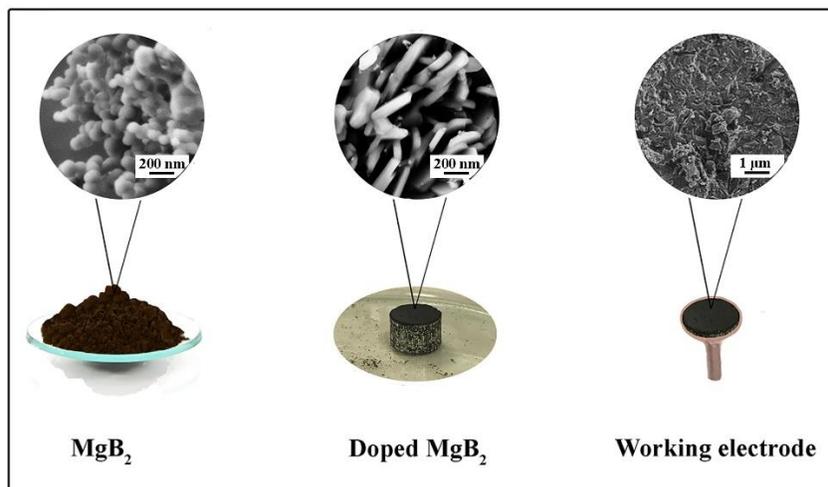


Figure 8. Illustration of final states of MgB_2 powder, doped MgB_2 pellets, and the working electrode.

Apparatus and Procedure

Apparatus

The crystal structure and the phase purity analysis of the samples were carried out by X-ray diffraction (XRD, Rigaku Mini Flex 600) with $\text{Cu K}\alpha$ ($\lambda = 1.5418 \text{ \AA}$) radiation (40 kV voltage and 15 mA). Diffraction patterns were collected between 20° and 85° of 2θ . A zero-background powder specimen holder was utilized for the XRD analysis of all samples. High-resolution X-ray photoelectron spectroscopy (XPS) was performed with a Thermo K-Alpha with an $\text{Al K}\alpha$ source. The fittings were accomplished with Avantage software. All peaks were corrected with respect to the C1s peak at 284.5 eV. For the analysis of the microstructure and surface morphology coupled with chemical composition, Field Emission Scanning Electron Microscopy (FESEM) equipped with Energy Dispersive X-ray spectroscopy (EDX) analyses were employed by Zeiss Ultra Plus Field Emission Scanning Electron Microscope using accelerating voltages of 5 and 12 kV, separately.

Electrochemical Measurements

Electrochemical measurements of the prepared electrocatalysts were conducted in a 1 M KOH solution using a three-electrode cell equipped with standard Reduced Hydrogen Electrode (RHE; HydroFlex[®]) as reference electrode and Pt wire as a counter electrode. All experiments were evaluated at room temperature using a VersaSTAT Potentiostat Galvanostat. The working electrodes were made on a copper foil with a surface area of 0.35 cm^2 via deposition of powders through a polymeric binder. In brief, 30 mg of powders were mixed with a drop of the polymeric binder (paraffin wax) at about 100°C and anchored on the copper foil supported by copper wire with a diameter of 1 mm. Then, the mixture was dried at ambient temperature. Finally, the designed electrodes were covered by a tubular plastic to prevent any contact with the solution. It is worth noting that the powder was dispersed uniformly on the foil and covered the whole area completely. For the electrochemical measurements, 0.5 g (powder) per mL (solution) was used for each electrode (see Figure 8).

For electrochemical measurements, each sample was undergone open circuit potential (OCP) measurements for 900 s to stabilize in the electrolyte. Then electrochemical impedance spectroscopy (EIS) was used to evaluate the charge transfer process as the critical factor determining the electrocatalytic activity on the electrodes. The OCP potential was recorded from a frequency of 100 kHz to 0.1 Hz with a 10 mV rms sinusoidal modulation. Afterward, the HER electrocatalytic activity was measured by applying Linear Sweep Voltammetry (LSV) method from -0.2 to -1.6 V at a sweep rate of $5 \text{ mV}\cdot\text{s}^{-1}$ towards HER catalysis. Also, overpotential calculation was performed at -10 mA cm^{-2} as an index of HER performance. For the OER, LSV was recorded at the same rate as HER from 0 to $+1.8 \text{ V}$ towards OER catalysis.

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Acknowledgements

The authors would like to appreciate Dr. Barış Yağcı and other researchers at Koç University Surface Science and Technology Center (KUYTAM) for their help in FESEM and XPS measurements.

Author contributions statement

E.S synthesized and characterized the electrocatalysts. NS.P prepared the electrodes and conducted the electrochemical experiments. M.K proposed the concept of the electrode and helped with its fabrication. U.U helped with the interpretation of the electrochemical measurements. U.A supervised the whole project from synthesis to interpretation of the results. All authors reviewed the manuscript.

Competing interests

The authors declare no competing interests.

Additional information

Supplementary information is available for this paper at

Correspondence and requests for materials should be addressed to U.A

Figures

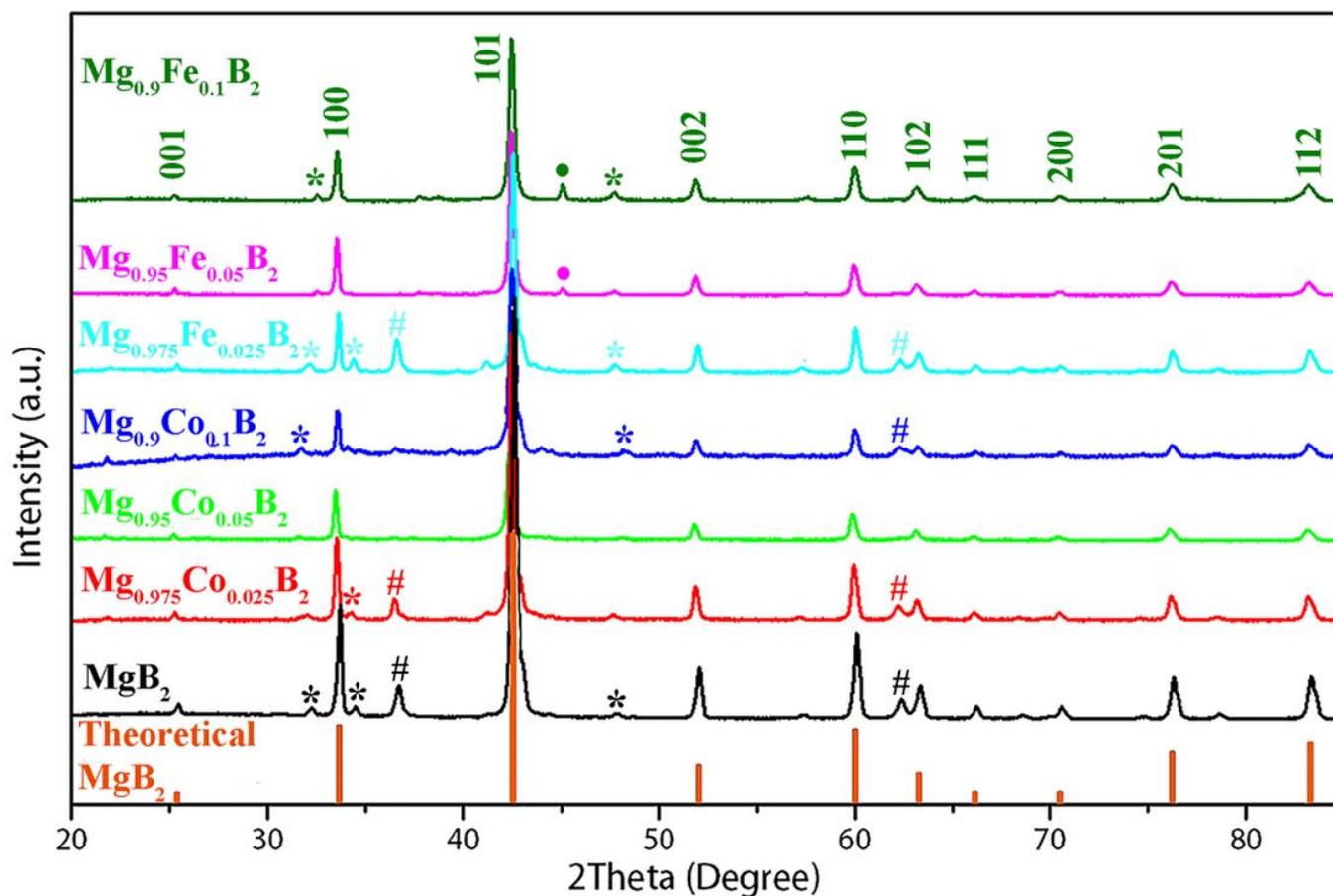


Figure 1

XRD patterns of MgB_2 and $\text{Mg}_{1-x}\text{Tm}_x\text{B}_2$ ($x = 0.025, 0.05, \text{ and } 0.1$; $\text{Tm} = \text{Fe and Co}$): impurity peaks of Mg, MgO, and FeB/Fe₂B are marked by *, #, and •, respectively.

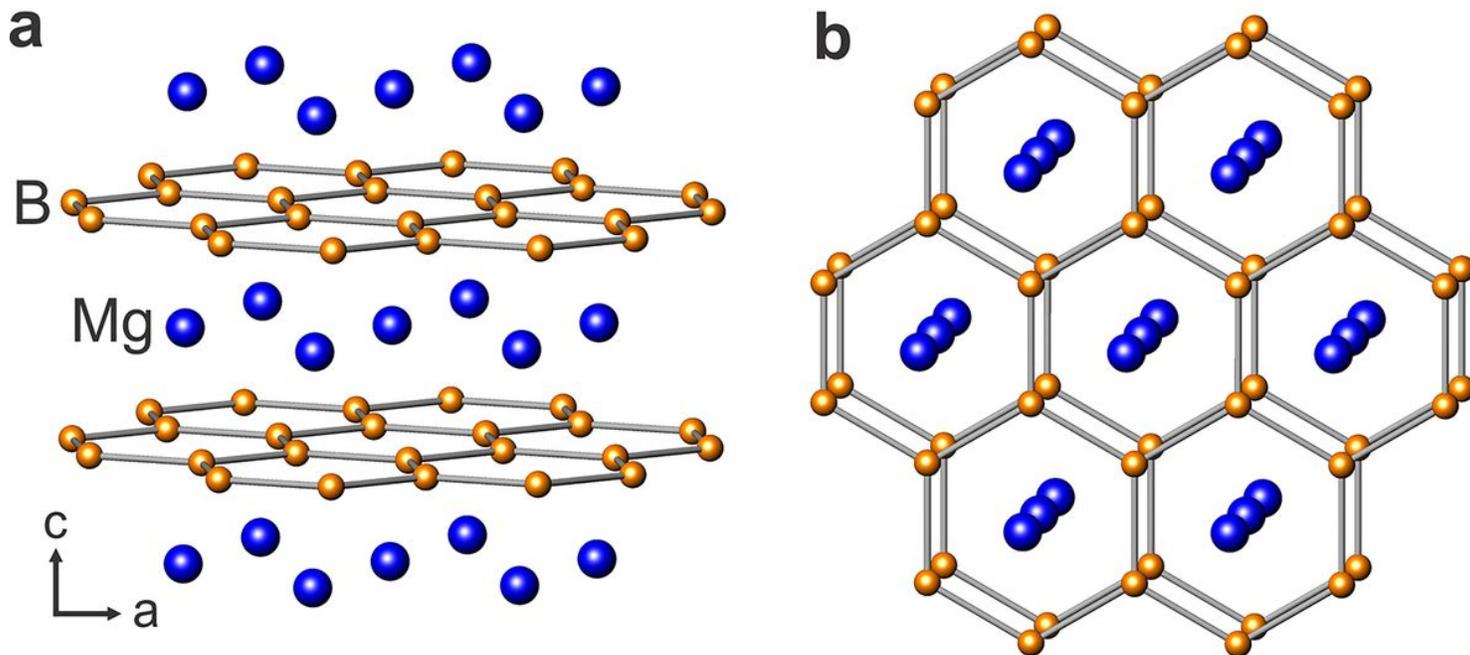


Figure 2

Schematic design of layered MgB₂ structure. a) along the a-axis, b) along the c-axis.

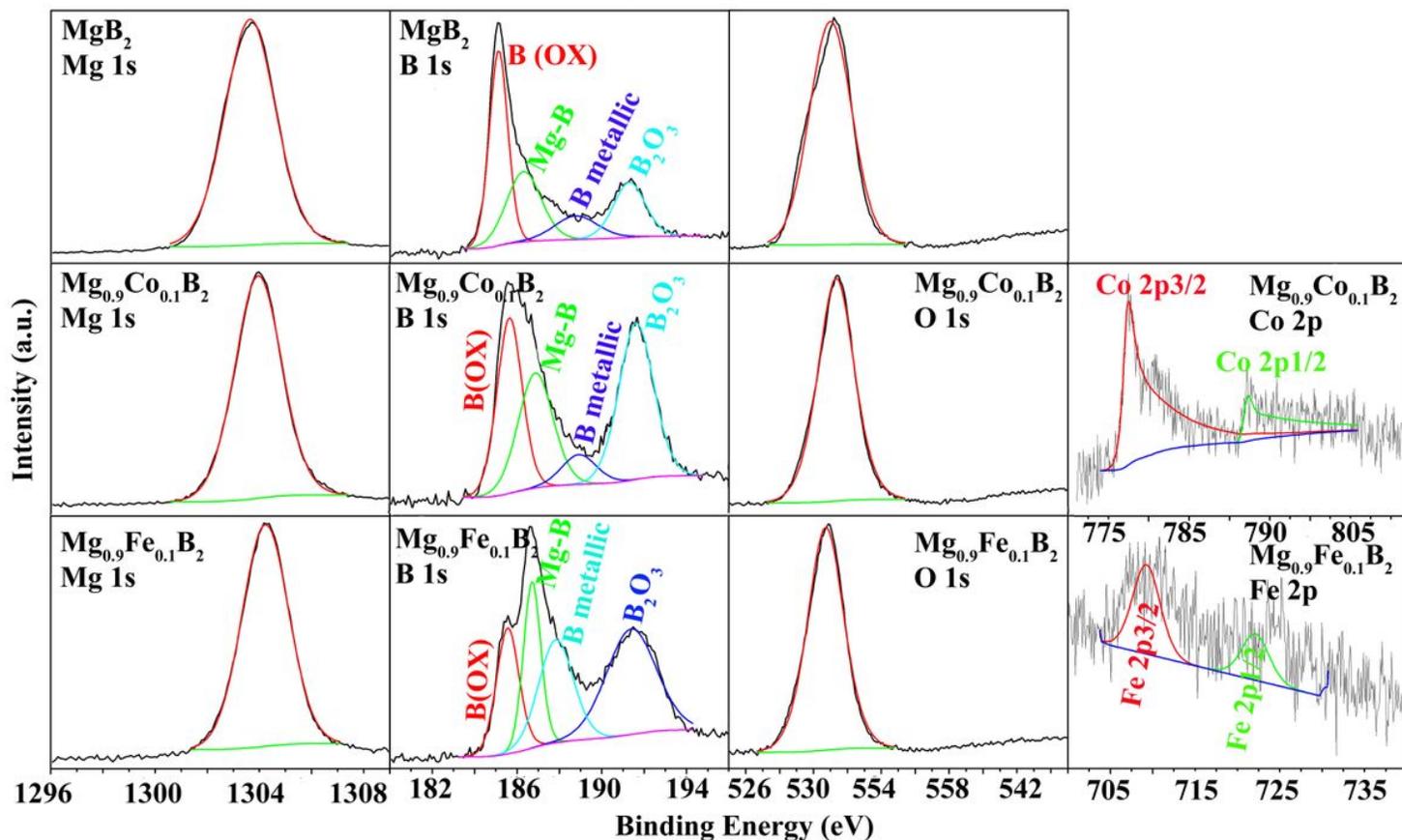


Figure 3

XPS spectra of Mg 1s, B 1s, O 1s, Co 2p, and Fe 2p in MgB₂, Mg_{0.9}Co_{0.1}B₂, and Mg_{0.9}Fe_{0.1}B₂.

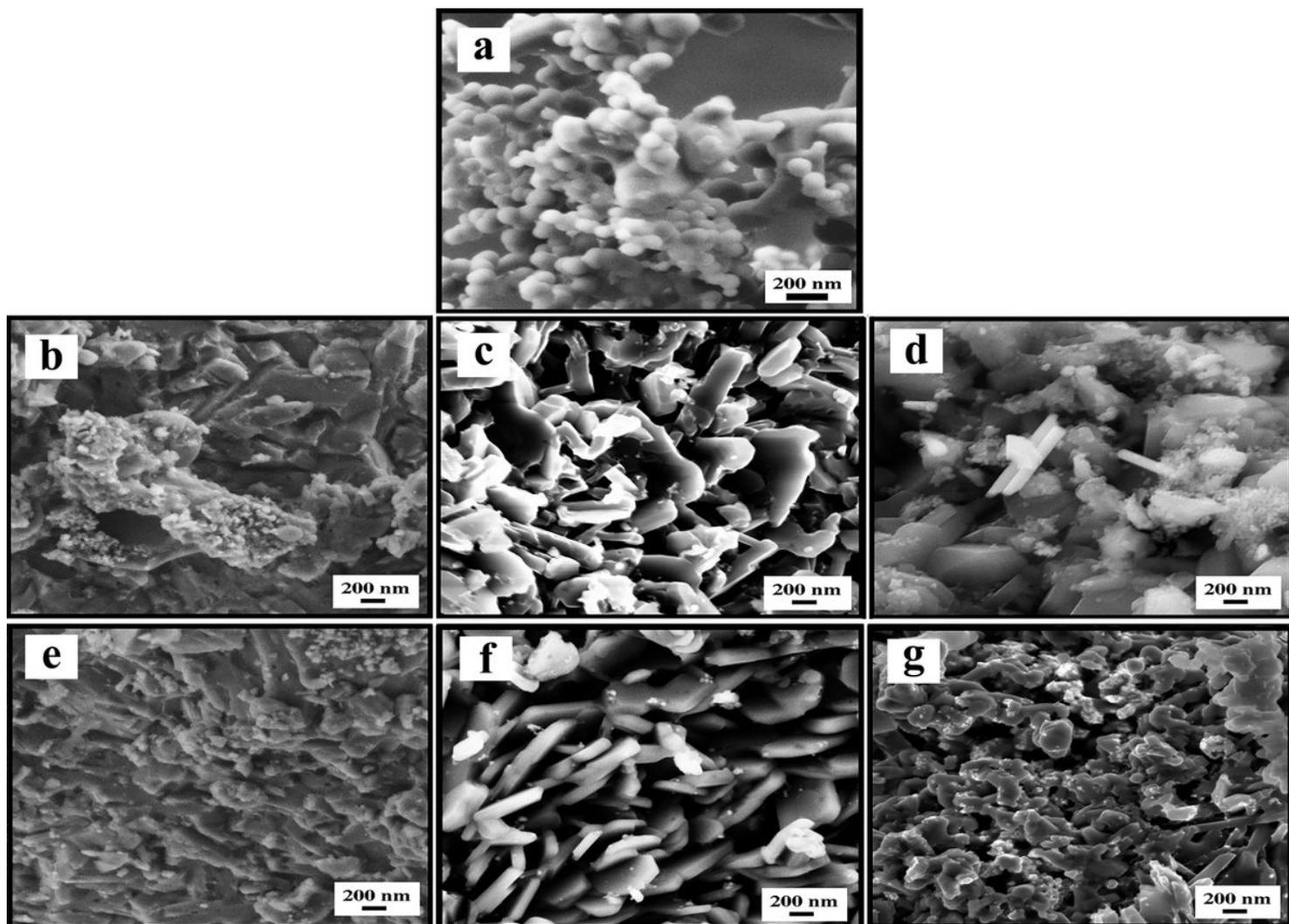


Figure 4

SEM images of a) pure MgB₂, b) Mg_{0.975}Co_{0.025}B₂, c) Mg_{0.95}Co_{0.05}B₂, d) Mg_{0.9}Co_{0.1}B₂, e) Mg_{0.975}Fe_{0.025}B₂, f) Mg_{0.95}Fe_{0.05}B₂, and g) Mg_{0.9}Fe_{0.1}B₂.

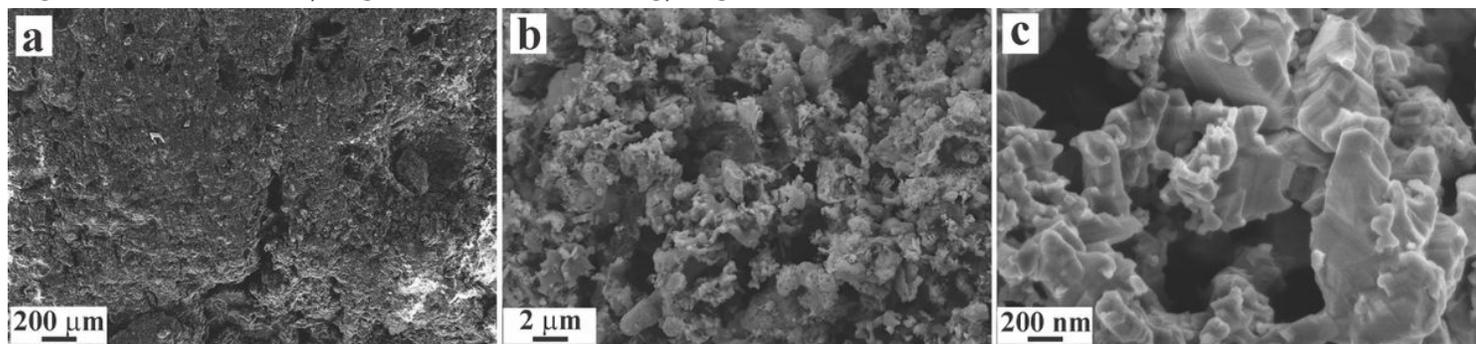


Figure 5

SEM images of the working electrode surface.

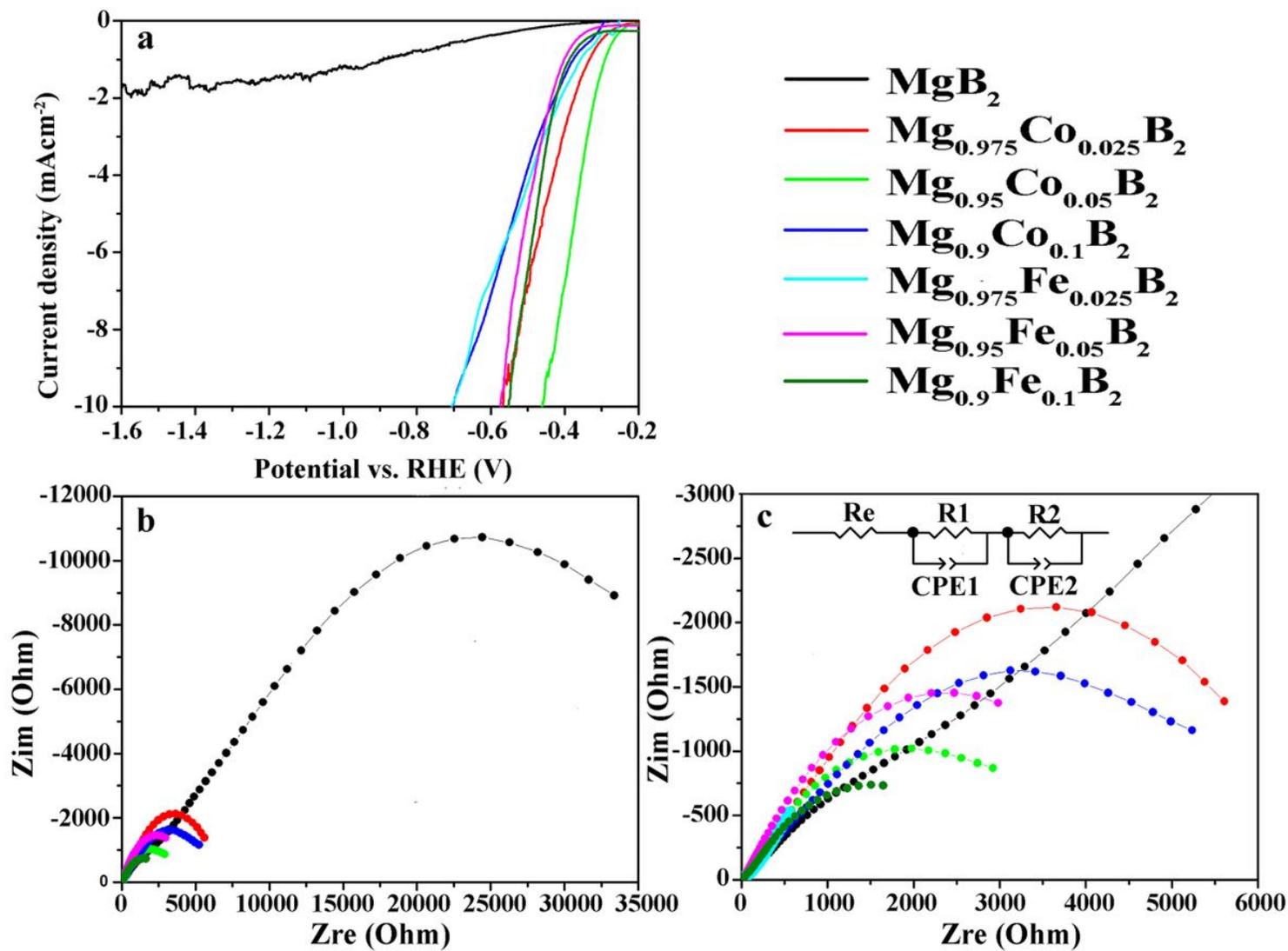


Figure 6

Polarization curves of a) the HER reaction of pure and doped MgB₂, b) and c) Nyquist plots measured at OCP in supporting electrolyte of 1 M KOH.

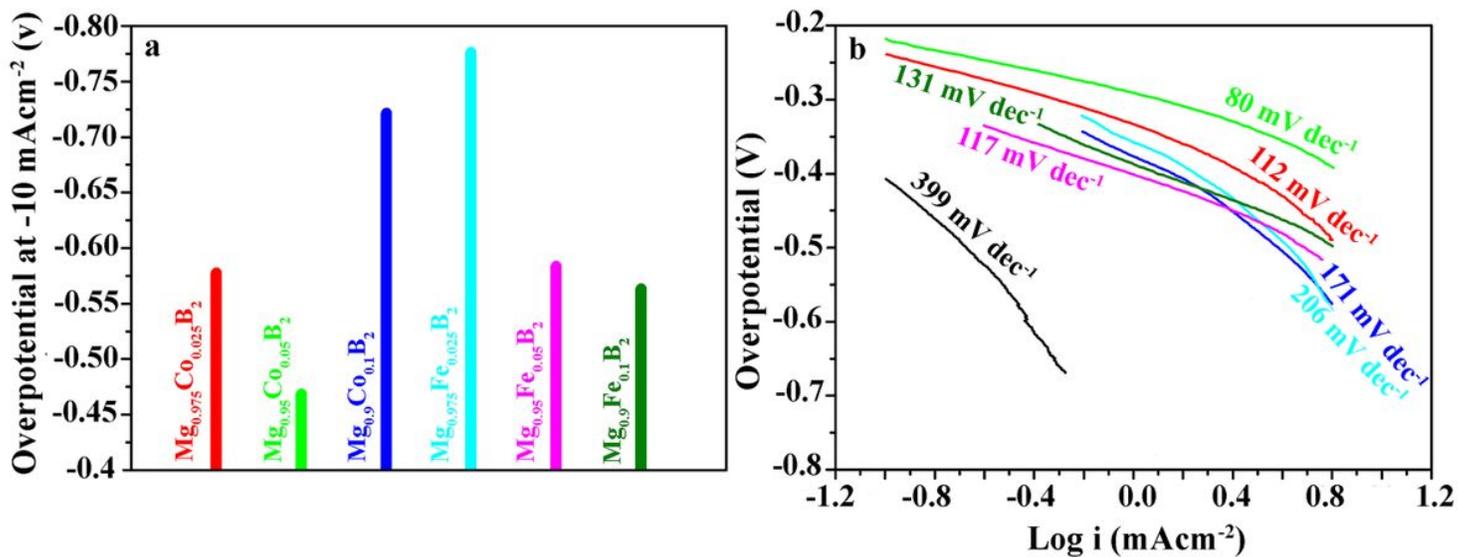


Figure 7

Fundamental electrochemical properties for HER. a) Overpotential at -10 mA cm^{-2} of metal-doped MgB_2 , b) The corresponding Tafel plots of undoped and doped MgB_2 . The colors indicating the same samples pictured in Figure 6.

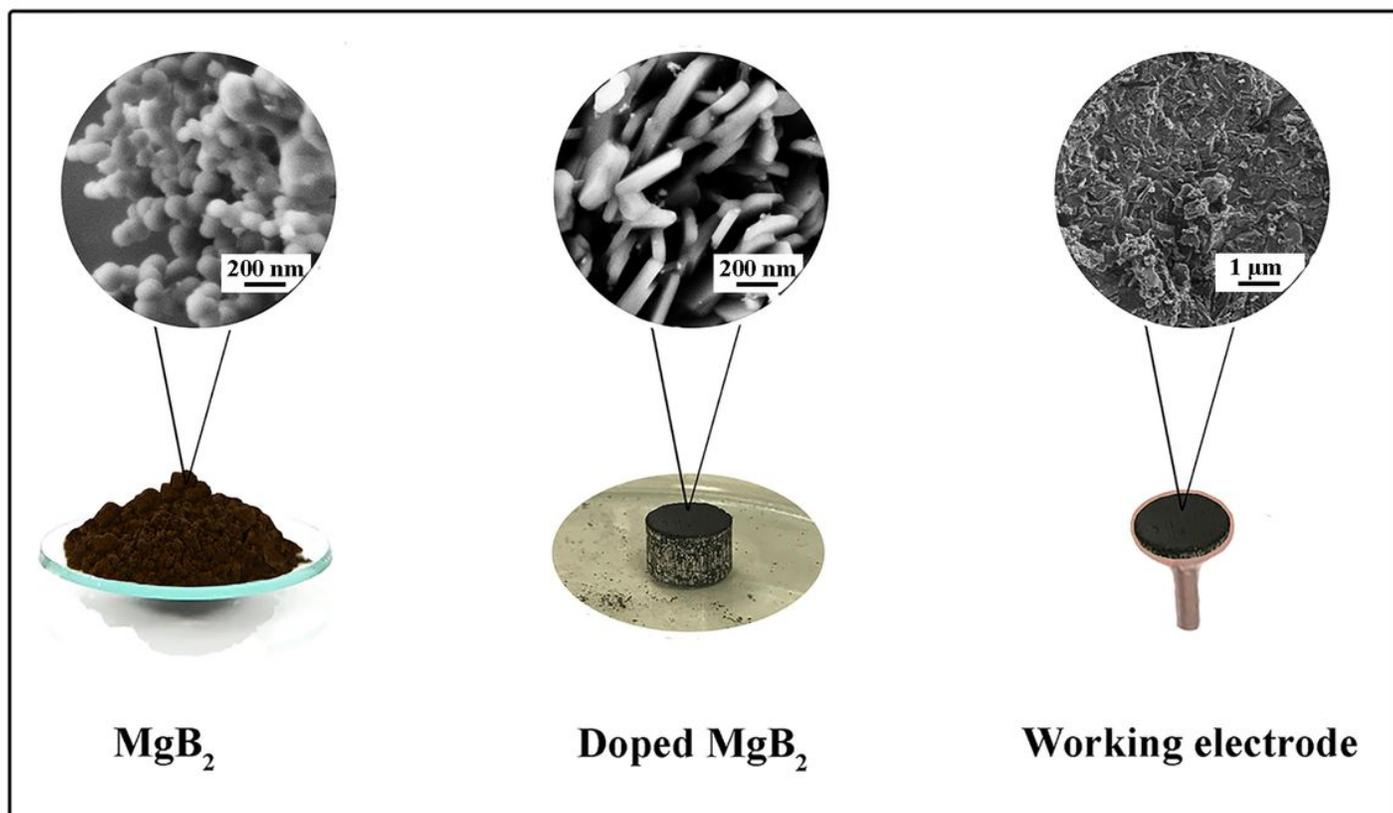


Figure 8

Illustration of final states of MgB₂ powder, doped MgB₂ pellets, and the working electrode.

Supplementary Files

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- [Supportinginformation.docx](#)