

# The *Samuel F. B. Morse* Statue in Central Park: Toward a Sustainable Regime of Science-Based Monuments Conservation

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## Research article

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# Abstract

Many of Central Park's bronze statues were coated with lacquer as a protective measure between the late 1970s and early 1990s. In several cases, these coatings outlasted their performance life and were no longer adequately protecting the sculptures. In 2010, a cyclical campaign to remove and refurbish the coatings provided an opportunity to closely assess the condition of de-coated surfaces and to develop suitable, sustainable treatment approaches. The *Samuel F. B. Morse* statue, created in 1870 by Byron Pickett, commemorates the 19th -century painter and inventor known for the telegraphic code that bears his name. Initially, samples from this statue were collected on the field and analyzed with various techniques to characterize the degradation products that were heavily affecting its surface and the organic coatings used to protect the bronze. Various corrosion species, mainly sulfates and chlorides, were identified by X-ray powder diffraction, while pyrolysis – gas chromatography / mass spectrometry revealed that the organic coatings were based on mineral wax and acrylics. These results informed a decision to relocate the statue to a controlled workshop space to aid in the manipulation of the large figure and optimize cleaning by laser ablation. Treatment was monitored through three campaigns of analysis using portable X-ray fluorescence (pXRF) spectroscopy, performed on select locations, to characterize the extent of corrosion across the sculpture's surface and monitor changes on the target areas throughout the cleaning. Following the initial phase of treatment, pXRF showed the substantial removal of iron- and lead-rich compounds from the brownish-toned coating and atmospheric pollution deposition. In the subsequent treatment phases, variations in the relative amounts of sulfates and chlorides were noted and further investigated through the examination of a cross section by means of scanning electron microscopy coupled with energy-dispersive X-ray spectroscopy (SEM/EDS). Based on the findings of iron-containing pigments in the organic coating, a decision to apply a new, similar, light brown-toned coating containing such coloring materials was made to achieve the desired visual nuance for the finish and provide sacrificial anodic protection for the bronze. Scientific analysis on the Morse statue has helped characterize its materials, methods of fabrication, and current conditions on a microscopic scale; in addition, it has guided the means and methodologies of treatment and relays essential information useful for future preservation.

## Introduction

More than fifty bronze sculptures have been placed in Central Park since its founding in the mid-19th century. Bronze, meant as representative of a wide range of copper alloys, is the most widespread material used in sculpture in the United States [1]. It is more resistant than stone to outdoor environmental conditions, nonetheless bronze corrosion represents a severe issue that requires constant monitoring of an object's conservation state and related interventions. Exposed to a variety of atmospheric agents over the years, the condition of the park's sculptures was dramatically compromised after the industrial revolution and the resulting increase of pollution. This situation was further undermined by the fact that New York City is located on a coastal environment, where marine chlorides in the local atmosphere may significantly contribute to the physicochemical degradation of bronzes. Starting from the 1970s, further concerns were raised about acid rain effects on outdoor cultural heritage. In the case of outdoor bronzes, in particular, acid attack appears to be more effective on patinas rather than on the bare metal [2].

Beginning in the late 1970s, many of Central Park's bronze statues were coated with lacquer as a protective measure. In several cases, these coatings outlasted their performance life and were no longer adequately protecting the sculptures. In 2010, a cyclical campaign to remove and refurbish the coatings provided an opportunity to assess the condition of de-coated surfaces, to develop suitable treatment approaches, and to consider context and sustainability of such treatments through a contemporary lens. Originally installed on the Central Park Mall, the *Samuel F. B. Morse* memorial was moved to its current location near Fifth Avenue at East 72nd street in 1988. When it was dedicated in 1871, the stately bronze figure of the 19th -century painter and inventor became the last monument in Central Park to honor a living person. Morse (1791–1872), a founder and the first president of the National Academy, is portrayed with his left

hand resting on a single line telegraph machine (Fig. 1). Created by sculptor Byron M. Pickett (1833–1907), the statue was well received, and a bust version was reproduced and offered for sale to the public.

Over time, the statue's surface has suffered from severe degradation. Exposed areas are deeply pitted, and sculptural detail has been lost as a consequence of decades of unprotected exposure to corrosive atmospheric agents such as salts and acid rain. Furthermore, blisters of active corrosion and associated localized coating failure have long been noted in the deep recesses of the figure's long coat. These observations prompted conservation staff at the Central Park Conservancy (CPC) to undertake a research project aimed to investigate the ongoing issues and formulate an adequate treatment. The Morse statue was thus removed from its granite base and relocated to a controlled workshop space to aid in the manipulation of the large figure and in devising a cleaning plan. The need of a precise, quick, and safe approach for the object, conservators, and the environment, guided the choice of treatment towards laser cleaning. This treatment uses specific, tunable laser parameters such as wavelength, as well as pulse frequency and duration, to preferentially excite and safely remove undesired coatings or contaminants from a surface through photomechanical, photothermal, and photochemical reactions [3]. Among its successful applications, laser cleaning has proven highly effective for removing corrosion products from bronze sculptures [3–9] and archaeological objects [10–13]. One of the main advantages of this cleaning methodology is the possibility to select suitable settings for achieving the expected results, preventing the operator from damaging the artwork [14].

In the present work, various techniques of scientific analysis were used to shed light on the conservation state of the *Samuel F. B. Morse* statue and to offer guidance to conservators during the cleaning treatment. In detail, X-ray powder diffraction (XRD) revealed the exact nature of the corrosion products found on the sculpture's surface, mostly consisting of sulfates and chlorides, while optical microscopy and scanning electron microscopy with energy-dispersive X-ray spectroscopy (SEM/EDS) provided precious information on their distribution and thickness within a cross section's stratigraphy. In addition, pyrolysis - gas chromatography / mass spectrometry (Py-GC/MS) was employed for the characterization of the organic coatings covering the statue, revealing a combination of mineral wax and acrylics. Notably, three analytical campaigns using portable X-ray fluorescence (pXRF) spectroscopy were also carried out to monitor the substantial removal of the Morse figure's unstable bronze condition and its failing lacquer coating. The Conservancy's goal for conserved bronzes in the Central Park's collection - with input and oversight of New York City's Design Commission - is to impart minimal net change from the sculpture's previous intended appearance. Therefore, with these considerations in mind and based on the detection of iron-containing pigments in the sculpture's organic coatings, a decision was made, following the laser cleaning treatment, to apply a new, similar, light brown coating containing such coloring materials to achieve the desired visual nuance for the finish and provide sacrificial anodic protection for the bronze.

Scientific analysis on Central Park's *Samuel F. B. Morse* statue has provided significant insight into the materials, methods of fabrication, and current condition issues of this outdoor sculpture on a microscopic scale; in addition, it has offered invaluable guidance in the choice of the means and methodologies of treatment and relays essential information useful for future preservation.

## Experimental

A careful visual assessment of the conservation state of the Morse sculpture provided initial information on the distribution of the dark brown organic coating on the statue's surface and locations where such coating was missing, as well as on the presence of green and red corrosion products. To investigate the ongoing condition issues, thirteen microscopic samples were removed from corroded areas using a metal scalpel (Table 1). Most of these samples were analyzed by XRD in the form of scrapings, whereas one was embedded in EpoThin™ 2 epoxy resin and mounted as a cross section, abraded using Micromesh® cloths and then finely polished with diamond pastes (6, 3, 1, 0.25 micron) on

wool cloths in preparation for SEM/EDS analysis. One sample of organic coatings was also removed from the surface in an adequate amount for Py-GC/MS analysis. Portable XRF was then used non-invasively to monitor the laser cleaning treatment through point analysis of several select areas on critical locations of the sculpture.

Table 1

List of microscopic samples removed from the Morse sculpture with corresponding descriptions and sampling sites.

Number	Description	Location
S1	Scraping of corrosion products	Inner portion of mantle at bottom, below proper right hand
S2	Scraping of corrosion products	Inner portion of coat at proper left
S3	Scraping of organic coatings	Area underneath telegraph, below proper left thumb
S4	Scraping of corrosion products	Inner portion of mantle at lower proper right
S5	Scraping of corrosion products	Inner portion of mantle in between legs
S6	Scraping of corrosion products	Inner portion of mantle at lower proper right
S8	Scraping of corrosion products	Inner portion of coat at proper right
S10	Scraping of corrosion products	Area above proper left knee at front
S11	Scraping of corrosion products	Area from proper right thigh at front
S12	Fragment of metal alloy for cross section	Area of internal crease of the mantle at top back
S13	Scraping of corrosion products	Surface of pedestal, proper left front corner
S14	Scraping of corrosion products	Area of proper left coat, above first button
S15	Scraping of corrosion products	Area of coat, over proper left shoulder blade at back

## X-ray powder diffraction

XRD analysis was performed using an open-architecture Philips PW1835 X-ray diffractometer and PW1710 controller on thirteen samples, selected on the basis of macroscopic observations. Samples were finely ground in an agate mortar and mounded with acetone on low-background Si (510) plate. Diffraction patterns were collected between 5° and 70° 2 $\theta$ , using Cu K $\alpha$  radiation, set at 40 kV voltage and 30 mA current.

## Pyrolysis – gas chromatography / mass spectrometry

Analysis was carried out on an Agilent 5973N gas chromatograph equipped with a Frontier PY-2020iD Double-Shot vertical furnace pyrolyzer fitted with an AS-1020E Auto-Shot autosampler. The GC was coupled to a 5973N single quadrupole mass selective detector (MSD). Two samples of 30–50  $\mu$ g were weighed out in deactivated pyrolysis sample cups (PY1-EC80F Disposable Eco-Cup LF) on a Mettler Toledo UMX2 Ultra microbalance. Samples were then either pyrolyzed without derivatization or derivatized with tetramethyl ammonium hydroxide (TMAH) before pyrolysis. Derivatization took place in the same cups as follows: 3–4  $\mu$ l of 25% TMAH in methanol (both from Fisher Scientific), depending on the sample size, were added directly to the sample in each cup with a 50- $\mu$ l syringe and, after 1 min, loaded onto the autosampler. The interface to the GC was held at 320 °C and purged with helium for 30 s before opening the valve to the GC column. The samples were then dropped into the furnace and pyrolyzed at 550 °C for 30 s. The pyrolysis products were transferred directly to a DB-5MS capillary column (30 m  $\times$  0.25 mm  $\times$  1  $\mu$ m) with the helium carrier gas set to a constant linear velocity of 1.5 mL/min. Injection with a 30:1 split was used, in accordance with the sample size. The GC oven temperature program was: 40 °C for 1 min; 10 °C/min to 320 °C; isothermal for 1 min. The

Agilent 5973N MSD conditions were set as follows: transfer line at 320 °C, MS Quad 150 °C, MS Source 230 °C, electron multiplier at approximately 1770 V; scan range 33–550 amu. For samples run with TMAH, the detector was turned off until 3 min to avoid saturation by excess of derivatizing agent and solvent. Data analysis was performed on an Agilent MSD ChemStation D.02.00.275 software by comparison with the NIST 2005 spectral libraries.

## **Portable X-ray fluorescence spectroscopy**

Analysis by pXRF was performed non-invasively using an Elio XGLab energy dispersive X-ray fluorescence analyzer, with a high resolution large area Silicon Drift Detector (SDD) with 130 eV at manganese (Mn) K $\alpha$  with 10 kcps input photon rate (high resolution mode), 170 eV at MnK $\alpha$  with 200 kcps input photon rate (fast mode). The system is equipped with changeable filters, and a rhodium (Rh) transmission target with 50 kV maximum voltage and 4W maximum power. The size of the analyzed spot is 1 mm, and analyses were collected for 60 seconds live time, setting the voltage at 40 kV and the current at 50 mA.

## **Scanning electron microscopy with energy-dispersive X-ray spectroscopy**

SEM/EDS analysis of a cross section of the metal alloy was performed with a FE-SEM Zeiss Sigma HD, equipped with an Oxford Instrument X-MaxN 80 SDD detector. Back-scattered electron (BSE) imaging, energy-dispersive spectroscopy (EDS), and X-ray mapping were performed in high vacuum with an accelerating voltage of 20 kV, on the cross section mounted on an aluminum stub and coated with a 14-nm layer of carbon, at a working distance of 8.5 mm.

## **Laser cleaning equipment**

Laser treatment was selected for its precision, relative speed and safety, environmental consciousness, and tuning capability that limits the potential for damage. A class 4 laser safety area (LSA) with ample warning signage was configured inside Central Park's conservation workshop. Two overhead hoists permitted the 1500 lbs+ (680 kg+) figure to be laid down and manipulated during treatment to provide access to recessed surfaces for optimal laser cleaning and outfitting with new bronze mounting hardware. The LSA consisted of a 12' H x 18' L x 12' D (about 3.66 m x 5.49 m x 3.66 m) opaque, curtained enclosure to prevent accidental viewing from outside during laser operation. Cleaning was accomplished using a GC-1 solid state, tunable 1064-nm wavelength near infrared laser. Overall laser cleaning of the surface employed a 16-cm focal length lens and consisted of multiple passes with settings ranging 240–270 ns in pulse duration, 70,000 pulse frequency at 90–100% laser power, and 9,000–12,000 RPM. Fluence (energy density) at the surface was delivered at a maximum of 5.66 J/cm<sup>2</sup>. Selective, increasingly aggressive cleaning of corrosion blisters associated with bronze disease employed a 10-cm focal length lens fired at 270-ns pulse duration at 70,000 pulse frequency, using 100% laser power and 5,000 RPM. Fluence was delivered at a maximum of 19.9 J/cm<sup>2</sup>. A Nederman modular gas and vapor carbon filtration system, outfitted with a #12603461 replacement filter and a 1.5 HP NIB 2001 model blower/fan run on 110V AC, was used to capture fumes generated by ablated materials.

## **Results And Discussion**

Preliminary visual examination of the Morse sculpture enabled conservators to identify the main ongoing conservation issues, which were mostly related to corrosion and to the presence of one or more degraded organic coatings. Over time, the statue's surface has suffered from severe corrosive attack. Pitting and loss of sculptural surface details (Fig. 2) were caused by the exposure to atmospheric pollutants such as acid rain and chlorides from deicing salts and the marine environment, in addition to dry deposition of biogenic and anthropogenic pollutants such as NO<sub>2</sub>, SO<sub>2</sub>, and carboxylic acid. The exposed surfaces have remained stable since being regularly treated with cleaning interventions and protective coatings beginning in the 1990s. However, localized patches of virulent, irregularly shaped pustules of active corrosion were contributing to progressive surface loss (Fig. 3a) and associated coating failure in the deep recesses of the figure's

long cloak (Fig. 3b). The surfaces within the recesses are so well protected that they do not benefit from washing during typical rain events or associated run off. Morse's sculptural geometry may explain the localization of corrosion phenomena in those recessed and protected locations where elevated chloride concentration accelerates surface loss [15]. The aged and degraded coatings had likely cross-linked and was otherwise admitting moisture and corrosive agents directly onto the bronze substrate, and was no longer providing adequate protection to the sculpture.

### Analysis of Organic Coatings

The organic coating was investigated to determine its constituent materials and the possible presence of pigments. In order to do so, a sample (S3) was removed from an area of the sculpture underneath the telegraph, below the figure's proper left thumb, and analyzed by Py-GC/MS. All chromatograms obtained with and without TMAH derivatization appeared to contain series of alkanes organized in a specific distribution that is characteristic of mineral waxes (Fig. 4). From a conservation perspective, mineral waxes are known for having a limited lifetime, ranging from 2 to 5 years, and for being reversible [16]. Besides, they are easily cleaned with water, which makes them suitable for a low cost and low maintenance application. However, the sample examined was found to contain also high relative amounts of acrylics, mostly methyl methacrylate and acrylic acid ethyl ester. The presence of these compounds is consistent with the use of Incralac® [17], which is, in fact, a product commonly used in the past and reported to have been applied to the Morse statue in the early 1990s according to the available conservation records. Acrylic coatings have a longer lifetime compared to mineral waxes, and they also show good reversibility properties [16]. Acrylic lacquers are coupled with a sacrificial wax topcoat that is renewed yearly in the case of Morse. Since 2003, two types of wax topcoats have been used by CPC's conservators: 1) Boston Polish™, also known as "butchers" paste wax, a carnauba-based wax, in use from 2003 to 2006; 2) Renaissance microcrystalline wax™, a mixture of Cosmoloid H80 and BASF A wax [18–19], used from 2007 to date for protecting outdoor bronze sculptures in Central Park.

In terms of coloring, verbal accounts and visual observation suggested that the Morse sculpture might have been coated with pitch at some early stage, although there are no written records of this intervention. The color of organic coatings in bronze sculptures is also sometimes obtained through the addition of bitumen, which might lend the coating a dark or brown hue [20]. Accordingly, visual inspection indicated that, possibly, very low amounts of pitch or bitumen residues may have been present on the back of the bronze, although, in this case, no samples were removed to verify this hypothesis, given the small dimension of the areas in question and in order to preserve the sculptures' integrity and appearance. However, an unexpectedly high amount of iron (Fig. 5) was detected by pXRF analysis on several locations of the surface still covered by the organic coatings, suggesting the use of iron oxide/hydroxide pigments as coloring agents in the acrylic layer. Comparisons with spectra collected on areas of bare metal confirmed the exclusive presence of iron, and the absence of other elements, in such organic layer.

### Analysis of Corrosion Products

In areas of the Morse sculpture that were less exposed to the atmospheric agents and not covered by the protecting coatings, various shades of green corrosion products were observed and sampled for analysis. XRD detected several compounds mostly attributable to the following categories: sulfates, sulfides, chlorides, and oxides. The hydrous copper sulfate mineral antlerite ( $\text{Cu}_3(\text{SO}_4)(\text{OH})_4$ ), found in all samples examined, appears to be the main sample constituent among those present, sometimes in association with minor amounts of copper trihydroxichlorides such as atacamite or clinoatacamite ( $\text{Cu}_2(\text{OH})_3\text{Cl}$ ) (as in the case of samples S1, S2, S4, S5). In addition, brochantite ( $\text{Cu}_4\text{SO}_4(\text{OH})_6$ ) and cuprite ( $\text{Cu}_2\text{O}$ ) were identified in samples S5 and S7, the latter also containing copper sulfide and gypsum (Fig. 6). All these products are described in the literature as bronze disease typically related to copper alloys [21], either originating from the reaction of copper with water or as a result of its interaction with atmospheric pollutants (leading to the formation of chlorides and brochantite, respectively).

Antlerite is frequently found as a main constituent of patinas in outdoor sculptures [22], while less common might appear the finding of chlorides. Due to the decrease of rainwater pH since the mid-20th century, antlerite also became a widespread component of corrosion crusts, mainly in partially exposed or sheltered surfaces [21]. Robbiola et al. [16] observed that brochantite often transforms into antlerite in outdoor sculptures, hypothesizing that the low pH rainwater might be the enhancing factor of this reaction. However, it has been demonstrated that degradation of bronze outdoor sculptures was more affected by dry deposition rather than acid rain in the 1980s [1]. In New York City, levels of SO<sub>2</sub> have constantly decreased especially in the last decade [23], with the consequence that dry deposition should no longer have substantial effects, while the corrosion rates are expected to return to pre-industrial levels in the future [1].

Considering that Central Park is relatively close to the ocean and therefore subjected to its atmospheric corrosion environment, and, above all, taking into account the enormous amount of sodium chloride distributed on the roads during wintertime, the detection of copper chlorides on the Morse statue, should not surprise. The sculpture had originally been sited on the Mall in Central Park, a pedestrian promenade, where it remained for more than 100 years. In 1988, the statue was relocated to its current location at the East 72nd Street entrance to the Park on the South side of the Drive roadway and in proximity to Fifth Avenue, where the exposure to ice melting salts is substantial during the winter months. In fact, chloride ions that separate from sodium as the ice melts are extremely corrosive. It seems plausible that the water enriched with these ions could become aerosolized by the constant traffic whisking by and deposited onto the bronze. Thickett and coauthors [18–19] have demonstrated that water may be retained at the interface between the superficial layer of Renaissance wax and the coating underneath, further enhancing the corrosive attack. On ferrous metals, like steel, the prolonged exposure to this trapped moisture after the protective period of the coating might increase the rate of corrosion. A similar process cannot be ruled out for copper-based alloys.

### Monitoring of Laser Cleaning

One of the main goals of the present study consisted in monitoring the laser cleaning intervention on the Morse statue by qualitatively analyzing the surface elemental composition before (pXRF campaign 1) and during (pXRF campaigns 2 and 3) the process. To achieve this goal, analysis was repeatedly carried out on eighteen select locations prior to cleaning and throughout a few subsequent phases of treatment to monitor changes on the target areas. Following a comparison between measurements recorded on the sculpture's external surface and reference ones collected on areas of bare metal on the pedestal, sulfur, chlorine, and iron were selected as the ideal candidate elements to monitor the progress of the laser cleaning. These elements, in fact, are considered respectively indicative of the presence of sulfates and sulfides (e.g. antlerite, brochantite, copper sulfides), chlorides (e.g. atacamite, clinoatacamite, paratacamite), which are the main corrosion products constituting the patina, as well as the coating, whose brown color, as explained above, is likely due to iron-containing pigments.

Following the initial phase of the cleaning, pXRF measurements showed relatively high levels of iron in all the spots analyzed, with the exception of a reference reading collected from an area of bare metal. Iron tends to decrease from measurements acquired during pXRF campaign 1 to those recorded in campaign 2, while it is found in trace amounts on the uncoated bronze (Fig. 5). This likely indicates that the occurrence of this element might be limited to the superficial organic coating, and, therefore, associated with the presence of iron oxide/hydroxide pigments in such coating, as mentioned earlier. Minor amounts of lead were also detected on locations yet to be cleaned, with a noticeable decrease in spectra collected upon treatment during campaign 2. This observation, along with the absence of lead in the reference measurements acquired from the bare metal, suggests that this element, whose presence was limited to the uppermost portion of the sculpture's stratigraphy and thus attributable to external causes, may be interpreted as a residue of the atmospheric pollution deposition. Interestingly, the amount of greenish corrosion products, namely chlorine-based and sulfur-based compounds, was found to decrease upon initial lasering, while an increase followed by a small decrease was noted following the second and third phases of treatment. The increased concentrations of chlorine and sulfur after

cleaning phase 2 indicate that the laser cleaning performed up until that point might have exposed a deeper portion of the corroded layer (Fig. 5). Another possible explanation for the observed increase of these elements' peak intensity might lie in the practice of misting the sculpture with distilled water after each lasering session. This process activates entrained chlorides and precipitates them onto the surface, from which they are ablated in subsequent cleaning sequences. As initial pXRF analyses were carried out after misting but before the following cleaning, it is plausible that re-precipitated chlorides were measured. A final phase of overall laser cleaning was applied to add variance of the laser's angle of attack. This was especially important for eradicating the deeply pitted areas where minute undercuts partially obscured direct access to those surfaces. The goal of this additional cleaning step was to further reduce the presence of powdery corrosion products and tenaciously adhered coating residue that may negatively affect proper adhesion of the new coating.

To evaluate the distribution and thickness of the various corrosion layers within the sculpture's stratigraphy, and to determine the exact composition of the metal alloy, a small sample removed from the rear of the Morse statue was mounted as a cross section and analyzed by SEM/EDS (Fig. 7). Results on the composition of the sculpture's constituting copper alloy (layer 1), calculated by means of EDS, are reported in Table 2. In detail, the relative amounts of copper (Cu) and tin (Sn) detected, alongside the observed zinc (Zn) content, are consistent with the composition of a ternary alloy and are very similar to gunmetal bronze (also known in the United States as red brass). This alloy was first identified in a 13th -century English memorial sculpture and later used for casting cannon. Its utilitarian characteristics include ease of casting in detail, toughness, resistance to corrosion, and machinability. The frequent use of this alloy to cast statuary in the United States in the mid-19th century may be attributed to the prevalence of the English emigrant group and its craft traditions, or even possibly to an expedient coincidence. For instance, the Ames Manufacturing Company of Chicopee, Massachusetts, is a prominent example of an American foundry concurrently producing military ordnance and "bronze" statuary during that period.

Table 2

Compositions of the Morse sculpture's corrosion layers and zinc sulfides in the main copper alloy obtained from SEM/EDS analysis of a cross section and expressed as normalized wt%. Numbers in brackets refer to layers in Fig. 7.

	S	Fe	Cu	Sn	Zn	SiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	K <sub>2</sub> O	P <sub>2</sub> O <sub>5</sub>	SO <sub>3</sub>	Cl	FeO	CuO	SnO <sub>2</sub>
Copper alloy (1)			91.0	6.4	2.6									
Red corrosion layer (2)										0.8	1.0		85.5	12.5
Green corrosion layer (3)						0.2			0.3	18.1	0.1	0.5	55.8	25.1
Very thin green corrosion layer (4)						1.6	0.9	0.2	0.3	23.5	2.2	0.5	70.3	0.6
Copper sulfide	21.4	0.2	78.4											
Zinc sulfide	33.1	0.3	4.4		62.2									
Zinc sulfide	32.9	0.6	5.7	0.3	60.5									
Zinc sulfide	33.1	0.4	4.2	0.2	62.1									
Zinc sulfide	34.1	0.5	4.9		60.6									
Zinc sulfide	33.6	0.4	4.0		62.0									

EDS analysis showed the additional presence, in the sculpture's ternary alloy, of small particles of less than 10 µm in size, with irregular shapes, scattered throughout the metal matrix, which were identified as zinc sulfides. These compounds appear to be sometimes associated with copper sulfides. Exact compositions of both zinc and copper sulfides are reported in Table 2. Regarding the pitting phenomenon also observed in the Morse sculpture, one must consider that the presence of pits in copper-based alloys is often related to the dezincification of the alloy [21]. Here, however, this hypothesis is unlikely, as red brass displays less of a proclivity for this type of deterioration than other alloys, in particular when protected by a thin layer of copper oxides, such as cuprite [24]. In the present case, analysis of the corrosion layers shows a clear transition from an underlying cuprite-based layer to a more superficial region that is mostly composed of copper sulfates, as indicated in Fig. 7. The first corrosion layer observed (layer 2, Fig. 7), showing red in the optical microscopy image, was found to be mainly composed of cuprite and enriched in tin. A second corrosion layer (layer 3, Fig. 7) above that, appearing green under polarized light, is characterized by the presence of copper sulfates and is richer in tin than layer 2. Traces of an additional, very thin, green corrosion layer (layer 4, Fig. 7), which contains a mixture of copper chlorides and dirt, are still visible on the surface. These data suggest that the corrosion process might have induced a preferential dissolution of copper and zinc linked to the presence of oxygen within the whole thickness of the corroded layer, as shown by the X-ray elemental maps of copper, chlorine, tin, and sulfur within the corrosion layer (Fig. 7). In outdoor conditions, the initial formation of a reddish-brown layer of cuprite is typically followed by the formation of different additional compounds, as also confirmed in the present study: copper basic sulfates, such as antlerite and brochantite, can be found on bronzes exposed in urban environments, while copper

hydroxychlorides, such as atacamite and paratacamite, are usually formed in chloride-rich environments [9 and references therein].

### Additional Conservation Treatment

As some of the Morse statue's surfaces were not easily accessible for laser cleaning, an additional treatment using propane-fueled torches to heat the surface and volatilize the remnant coating and corrosion products was performed in the deepest pits and recesses of the sculpture. Following the technical guidelines of the U.S. General Services Administration [25], two applications of benzotriazole (BTA), an effective corrosion inhibitor, were selectively applied on the undercuts and recesses where pustules had formed in an attempt to stabilize the surface from further loss by arresting or eradicating propagation of the condition. BTA reacts with corrosion products such as cuprite and copper trihydroxychlorides to form Cu(I)BTA and Cu(II)BTA copolymers, resulting in an optimal and stable coating that ensures durability of the artwork [26]. The double application of BTA, guaranteed thorough coverage in recessed areas, ensured complete reaction of this chemical with all the remaining available unstable compounds that were not complexed during the first application. After application of BTA, the treated area was cleaned with soft bronze bristle brushes and allowed to dry. The resulting precipitate matter was then flushed clear and the surfaces thoroughly rinsed with distilled water. After that, the sculpture was carefully steam pressure washed to remove any remaining residue. On the undersides and recesses of the long coat, i.e. the location most affected by bronze disease, the patina had been removed completely by laser ablation, exposing the bare bronze. Apart from this area, complete removal of the corrosion layers was outside the scope of the present project.

In terms of coloration, a selective application of cupric nitrate and cupric chloride patina solutions to achieve a range of green, and of ammonium sulfate for darker brown in the recesses of the long coat, were used to blend and unify tone and to enhance legibility of the sculptural forms. Finally, a translucent, warm brown coating that established a contiguous protective film of appropriate dry film thickness and unveiled nuances of the underlying greenish patina was applied as the new finish (Fig. 8). The identification of iron-containing pigments to color the lacquer finish guided the decision of applying a new coating containing the same type of coloring materials. Such lacquer, selected from Nikolas Coatings, is a low volatile organic compounds (VOC) lacquer (#14135 Clear NY Coat RFU) that was specifically formulated and re-named for this project (#15794 Morse brown toner RFU). The new toned lacquer achieved the desired visual nuance for the finish and the iron contained in it may provide sacrificial anodic protection to the bronze. In fact, it has been demonstrated that lamellar micaceous iron oxide, used as inert, electrochemically non-reactive pigment for anticorrosive coatings, reduces the reactivity of the material it is in contact with, acting as a protective barrier [27–28]. However, the substance used in this case is in fact an iron-based pigment composed of spherical particles. The solvent-borne coating was thinned and brush applied twice overall to assure penetration into the pitted texture of the Morse sculpture and to obtain complete coverage in deep recesses and undercuts. Three subsequent high volume - low pressure spray coats completed the coating application.

## Conclusions

The present study aimed to shed light on the materials, methods of fabrication, and ongoing condition issues of the *Samuel F. B. Morse* sculpture, located in Central Park, New York City, with special attention being paid to investigating the exact composition of the artwork's constituting metal alloy, the corrosion products affecting its surface, and the organic coatings that protect the bronze; to monitor the progress and assess the effectiveness of the laser cleaning treatment through scientific analyses; and to inform present and future conservation strategies. Results of scientific analysis identified the main metal alloy of the sculpture as a copper ternary alloy very similar to gunmetal bronze. The primary corrosion products found on the Morse statue consisted of sulfates (e.g. antlerite, brochantite) and chlorides (e.g. atacamite, clinoatacamite, paratacamite), along with lower amounts of sulfides and oxides (e.g. cuprite). On the other

hand, characterization of the organic coatings covering the surface revealed a combination of mineral wax and acrylics. Three analytical campaigns of portable XRF were performed on target locations of the work to monitor the laser cleaning and demonstrated the treatment's effectiveness in removing selected residues and compounds. Results of scientific analysis, in particular, were crucial in guiding the formulation and optimization of a suitable conservation protocol tailored to the removal of atmospheric pollutant deposition, of the outdated organic coatings, and of the uppermost portion of the corrosion patina. Apart from the bronze disease condition in the recesses of the figure's coat, however, complete ablation of the corrosion layers was outside the scope of this project. An additional consideration is that antlerite and brochantite are the most stable corrosion products among those identified, and as such they can play a beneficial role in preserving the current state of the artwork. Finally, the detection and identification of iron oxide/hydroxide pigments in the sculpture's organic coating guided a decision to replace it with a new light brown-toned lacquer containing the same type of coloring materials in order to achieve the desired visual nuance for the finish and provide sacrificial anodic protection to the bronze.

## Abbreviations

pXRF: portable X-ray fluorescence spectroscopy; XRD: X-ray powder diffraction; SEM/EDS: scanning electron microscopy with energy-dispersive X-ray spectroscopy; Py-GC/MS: pyrolysis - gas chromatography / mass spectrometry.

## Declarations

### Availability of data and materials

All data generated during this study are included in this published article.

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### Competing Interests

The authors declare that they have no competing interests.

## Authors' Contributions

EB performed XRD, SEM/EDS, pXRF, interpreted these data, and prepared the manuscript; FP performed Py-GC/MS, interpreted results, wrote the corresponding section of the article, reviewed the whole manuscript, and helped with pXRF analysis; MR designed and performed all conservation treatments, and contributed to drafting the conservation-related section of the paper. All authors read and approved the final manuscript.

## References

- [1] Livingston RA. Acid rain attack on outdoor sculpture in perspective. *Atmos Environ.* 2016;146:332-345.
- [2] Kratschmer A, Odnevall Wallinder I, Leygraf C. The evolution of outdoor copper patina. *Corros Sci.* 2002;44:425-450.
- [3] Dajnowski A. and Dajnowski B. Using the new G.C. Laser Cleaning System for Cleaning and Surface Preparation For Re-Gilding of a Large Outdoor Bronze Monument of Alexander Hamilton. In: Piotr Targowski et al., editors. *Proceedings of LACONA XI.* Toruń: NCU Press; 2017. p. 217-228.
- [4] Matteini M, Lalli C, Tosini I, Giusti A, Siano S. Laser and chemical cleaning tests for the conservation of the *Porta del Paradiso* by Lorenzo Ghiberti. *J Cult Herit.* 2003;4:147s-151s.
- [5] Siano S and Salimbeni R. Advances in laser cleaning of artwork and objects of historical interest: the optimized pulse duration approach. *Acc Chem Res.* 2010;43(6):739-750.
- [6] Buccolieri G, Nassisi V, Buccolieri A, Vona F, Castellano A. Laser cleaning of a bronze bell. *Appl Surf Sci.* 2013;272:55-58
- [7] Lee H, Cho N, Lee J. Study on surface properties of gilt-bronze artifacts, after Nd:YAG laser cleaning. *Appl Surf Sci.* 2013;284:235-241
- [8] Sansonetti A, Colella M, Letardi P, Salvadori B, Striova J. Laser Cleaning of a nineteenth-century bronze sculpture: in situ multi-analytical evaluation. *Stud Conserv.* 2015; 60(1): S28-S33.
- [9] Petiti C, Toniolo L, Gulotta D, Mariani B, Goidanich S. Effects of cleaning procedures on the long-term corrosion behavior of bronze artifacts of the cultural heritage in outdoor environment. *Environ Sci Pollut R.* 2020.

- [10] Pini R, Siano S, Salimbeni R, Pasquinucci M, Miccio M. Tests of laser cleaning on archaeological metal artefacts. *J Cult Herit.* 2000;1:S129-S137.
- [11] Colao F, Fantoni R, Lazic V, Caneve L, Giardini A, Spizzichino V. LIBS as a diagnostic tool during the laser cleaning of copper based alloys: experimental results. *J Anal At Spectrom.* 2004;19:502-504.
- [12] Garbacz H, Fortuna-Zalesna E, Marczak J, Koss A, Zatorska A, Zukowska GZ, Onyszczuk T, Kurzydowski KJ. Effect of laser treatment on the surface of copper alloys. *Appl Surf Sci.* 2011;257:7369-7374.
- [13] Cacciari I, Ciofini D, Mascalchi M, Mencaglia A, Siano S. Novel approach to the microscopic inspection during laser cleaning treatments of artworks. *Anal Bioanal Chem.* 2012;402:1585-1591.
- [14] Cooper M. *Laser cleaning in conservation.* Oxford: Butterworth Heinemann; 1998.
- [15] King GA and O'Brien DJ. The influence of marine environments on metals and fabricated coated metal products, freely exposed and partially sheltered. In Kirk WW and Lawson H (Eds): *ASTM STP 1239 Atmospheric corrosion.* American Society of Materials and Testing, Philadelphia. 1995, p. 167–192.
- [16] Robbiola L, Fiaud C, Pennec S. New model of outdoor bronze corrosion and its implications for conservation. In: *ICOM Committee for Conservation tenth triennial meeting.* Paris: International Council of Museums Committee for Conservation. 1993;2:796-802.
- [17] Shedlosky TJ, Stanek KM, Bierwagen G. On-line survey results of techniques used for outdoor bronze conservation. In: *AIC Objects Specialty Group Postprints.* 2002;9:3-13.
- [18] Thickett D and Stanley B. The use and mis-use of Renaissance wax. In: Chemello C, Brambilla L and Joseph E. *Proceedings of METAL 2019.* Neuchatel, Switzerland: ICOM-CC; 2019. p. 232-241.
- [19] Thickett D and Pretzel B. FTIR surface analysis for conservation. *Herit Sci.* 2020;8(5).  
<https://doi.org/10.1186/s40494-020-0349-8>

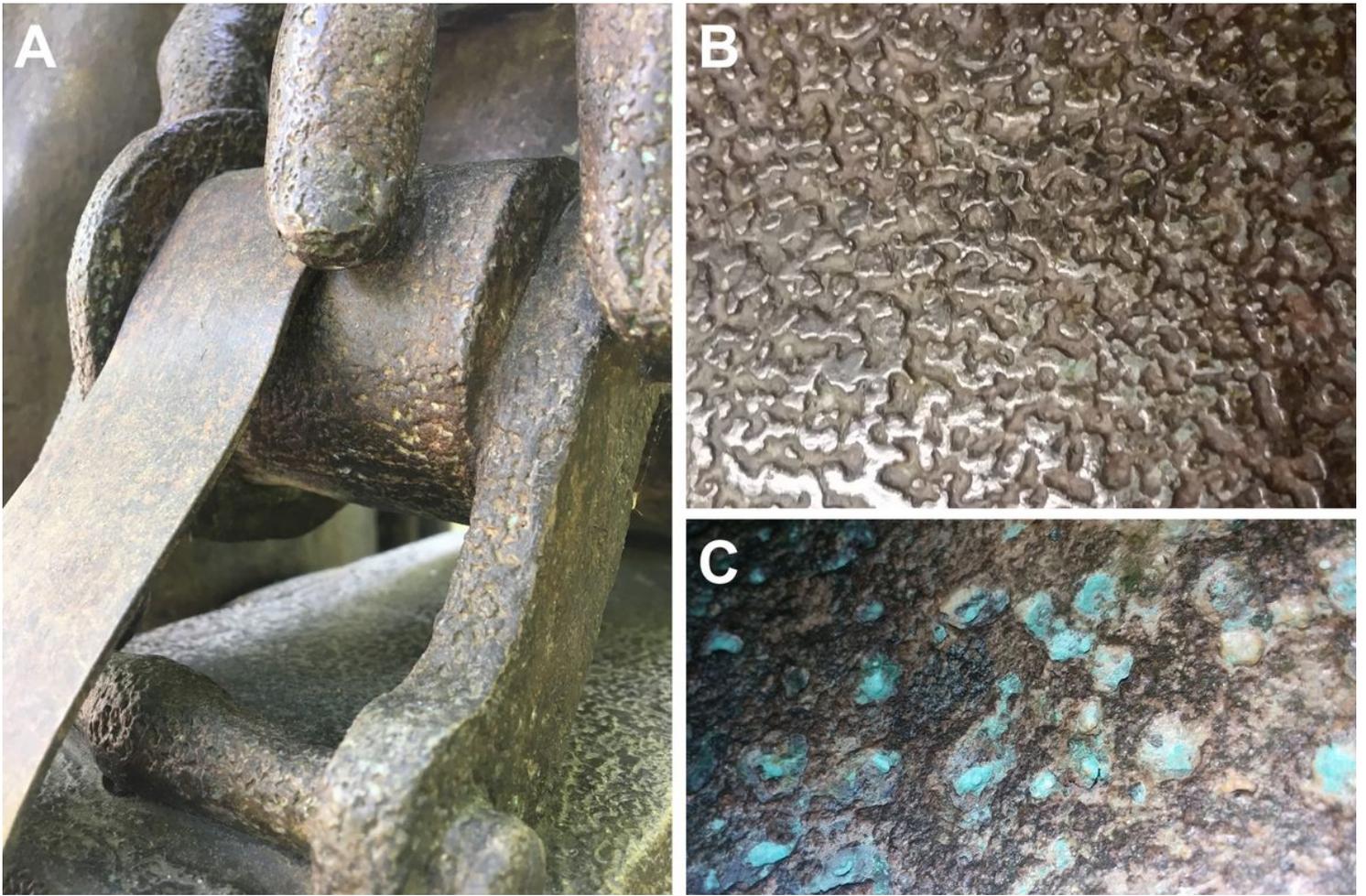
- [20] Pitthard V, Stone R, Stanek S, Griesser M, Gersch CK, Hanzer H. Organic patinas on Renaissance and Baroque bronzes-Interpretation of compositions of the original patination by using a set of simulated varnished bronze coupons. *J Cult Herit.* 2011;12:44-53.
- [21] Scott D. *Copper and Bronze in Art: Corrosion, Colorants, Conservation.* Los Angeles: Getty Publications; 2002.
- [22] Stambolov T. *The Corrosion and Conservation of Metallic Antiquities and Works of Art.* Amsterdam: CL Publication; 1985.
- [23] NYDEC 2018. New York State ambient air quality report for 2018. New York State Department of Environmental Conservation. [http://www.dec.ny.gov/docs/air\\_pdf/2018airqualreport.pdf](http://www.dec.ny.gov/docs/air_pdf/2018airqualreport.pdf) (accessed 04/14/2020)
- [24] Finnegan JE, Hummel RE, Verink, Jr ED. Optical studies of dezincification in alpha-brass. *Corrosion.*1981;37(5):256-261.
- [25] Outdoor Sculpture Manual - Center For Public Buildings. Applying Benzotriazole (BTA) To Bronze. [https://www.gsa.gov/technical-procedures/applying-benzotriazole-bta-bronze?Form\\_Load=88089](https://www.gsa.gov/technical-procedures/applying-benzotriazole-bta-bronze?Form_Load=88089). Accessed 27 Apr 2020.
- [26] Mezzi A, Angelini E, De Caro T, Grassini S, Faraldi F, Riccucci C, Ingo GM. Investigation of the benzotriazole inhibition mechanism of bronze disease. *Surf Interface Anal.* 2012;44:968-971.
- [27] Sørensen PA, Kiil S, Dam-Johansen K, Weinell CE. Anticorrosive coatings: a review. *J Coat Technol Res.* 2009;6(2):135–176.
- [28] Kakaei MN, Danaee I, Zaarei D. Investigation of corrosion protection afforded by inorganic anticorrosive coatings comprising micaceous iron oxide and zinc dust. *Corros Eng Sci Techn.*2013;48(3):194-198.

## Figures



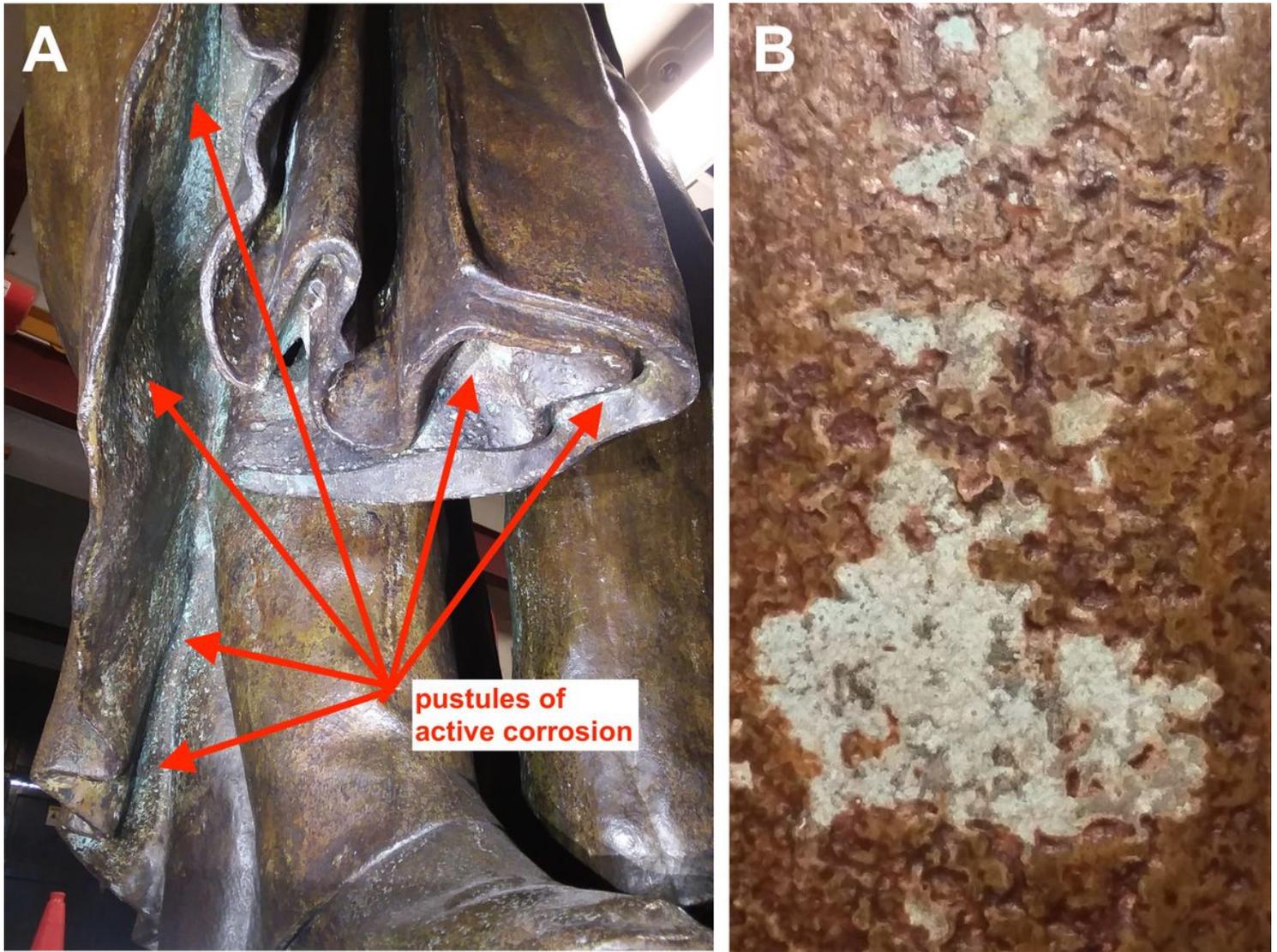
**Figure 1**

Statue of Samuel F. B. Morse. The statue of Samuel F. B. Morse in its current Central Park location at East 72nd Street, where it was moved in 1988.



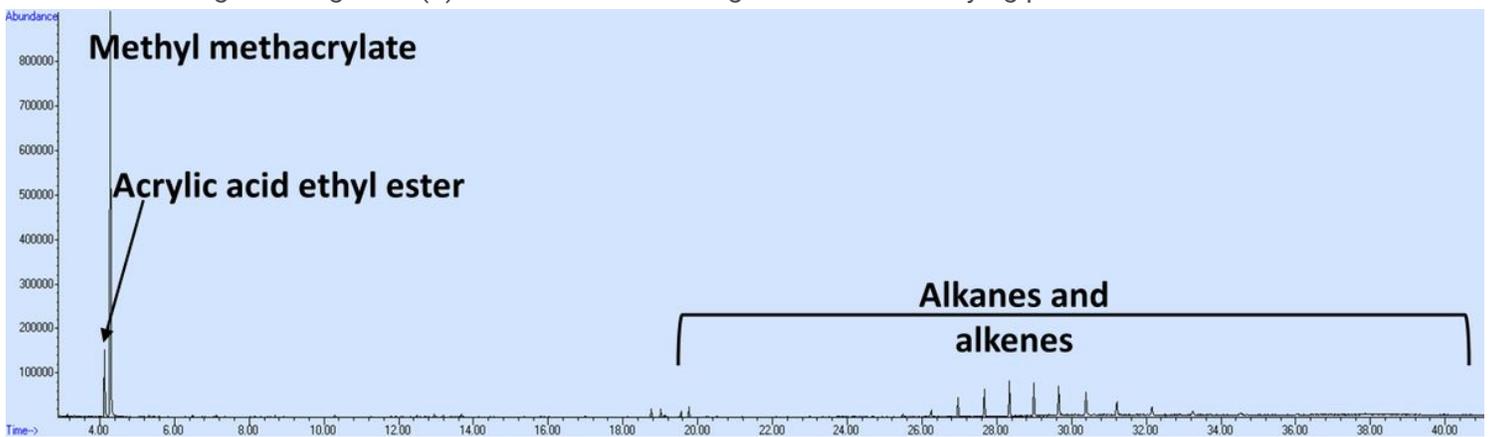
**Figure 2**

Condition issues observed on the Morse sculpture (I). (A) Detail of sculptural surface loss. (B) Typical pitting and loss pattern on exposed surfaces. (C) Detail of active corrosion present in the recessed areas.



**Figure 3**

Condition issues observed on the Morse sculpture (II). (A) Extent of active corrosion, or bronze disease, in the protected recesses of the figure's long coat. (B) Detail of toned coating failure and underlying patina.



**Figure 4**

Py-GC/MS chromatogram of sample S3 (organic coatings). Py-GC/MS chromatogram obtained upon derivatization with TMAH of sample S3, i.e. a scraping of the organic coatings. Compounds identified in the chromatogram indicate the presence of mineral wax and acrylics.

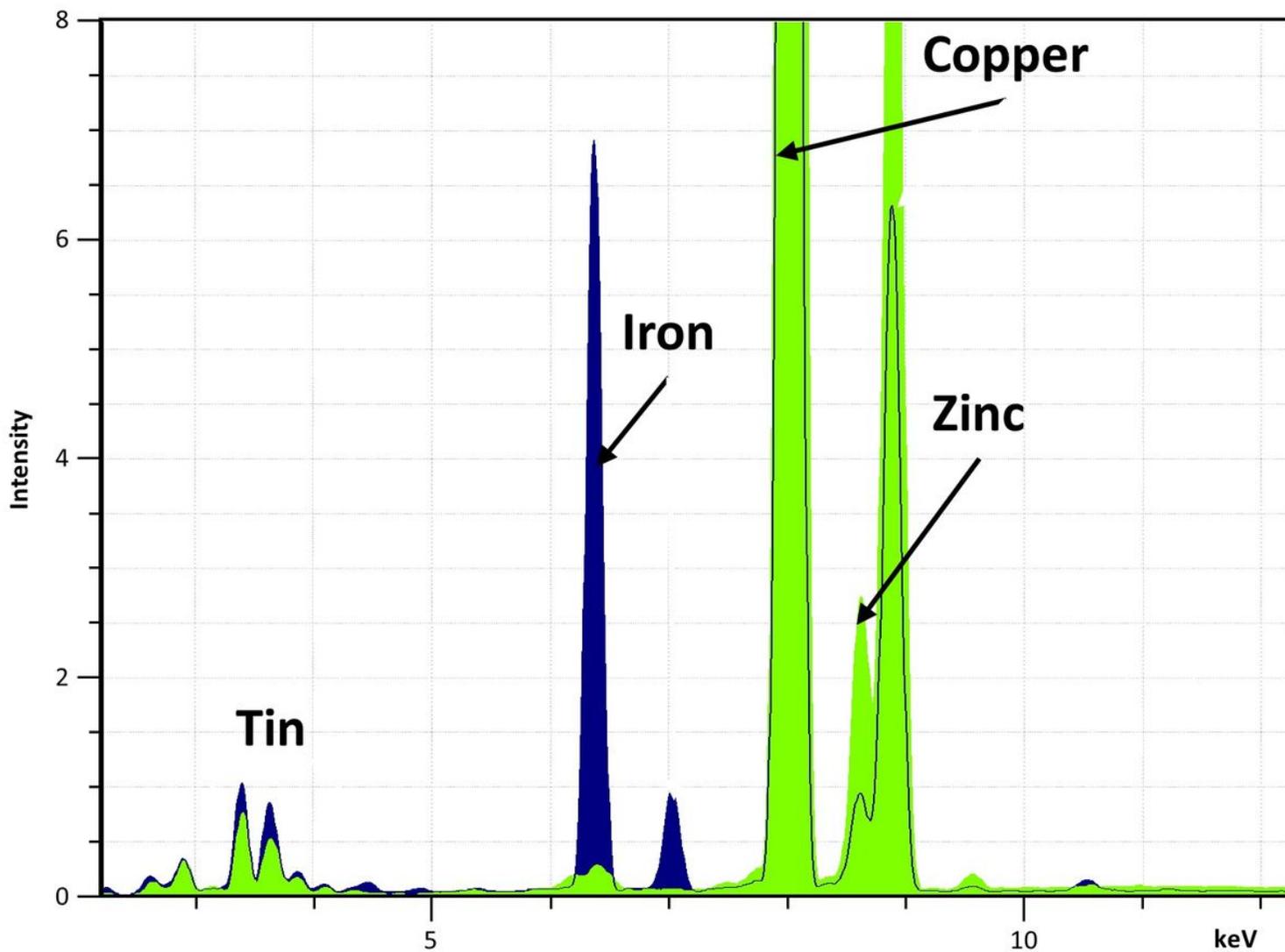


Figure 5

XRF spectra acquired prior to laser cleaning. Example of XRF spectrum acquired from the Morse statue before the laser cleaning treatment (pXRF campaign 1, blue pattern), showing the presence of higher amounts of iron than in the bare metal (green pattern).

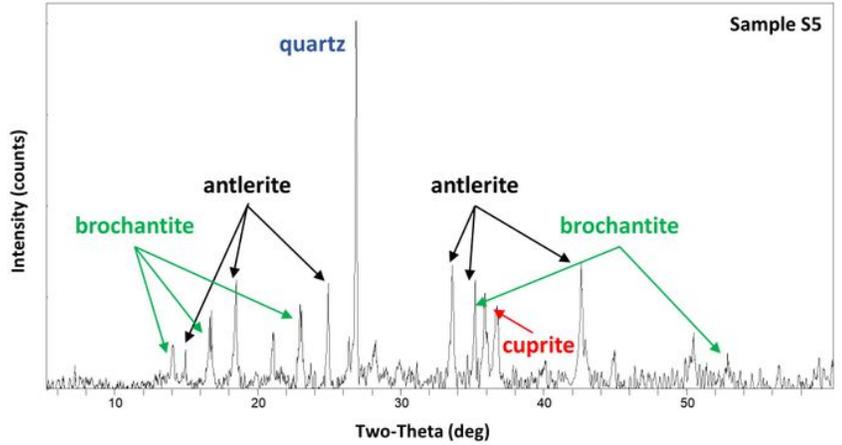
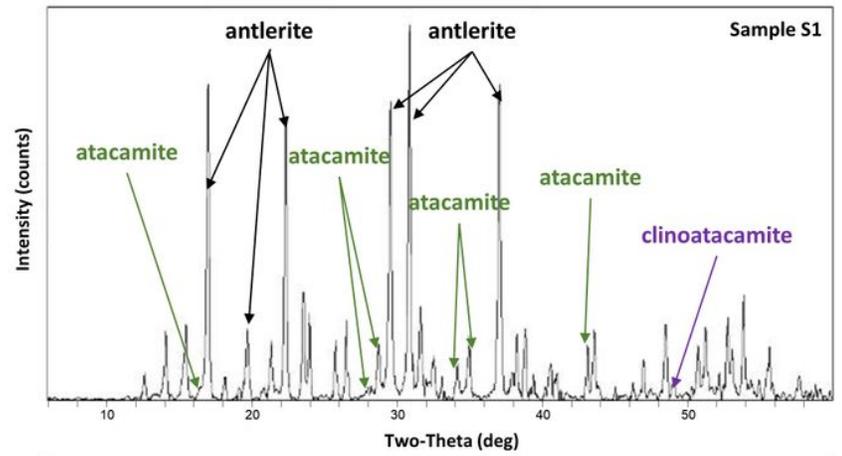
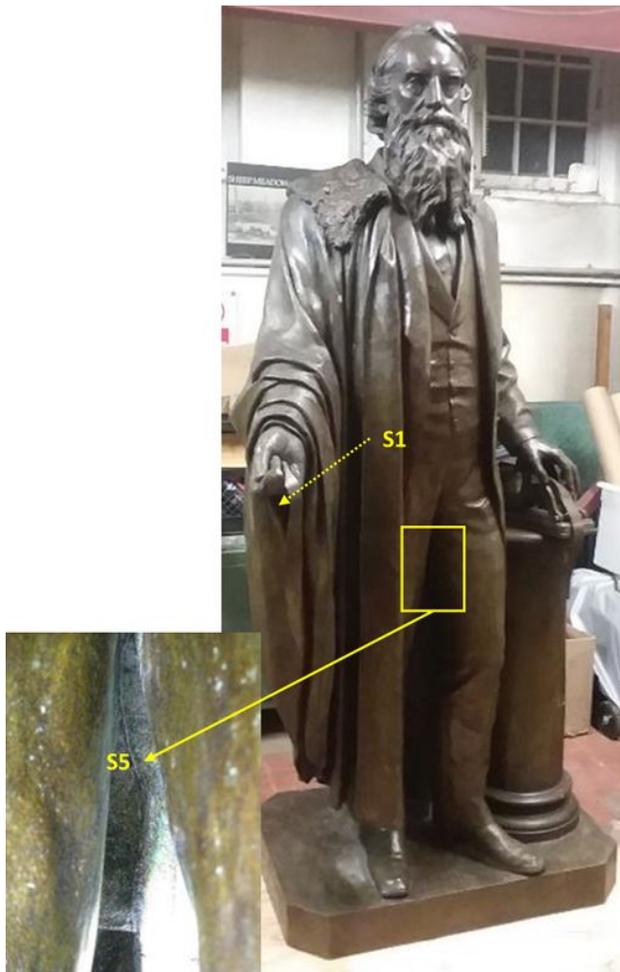
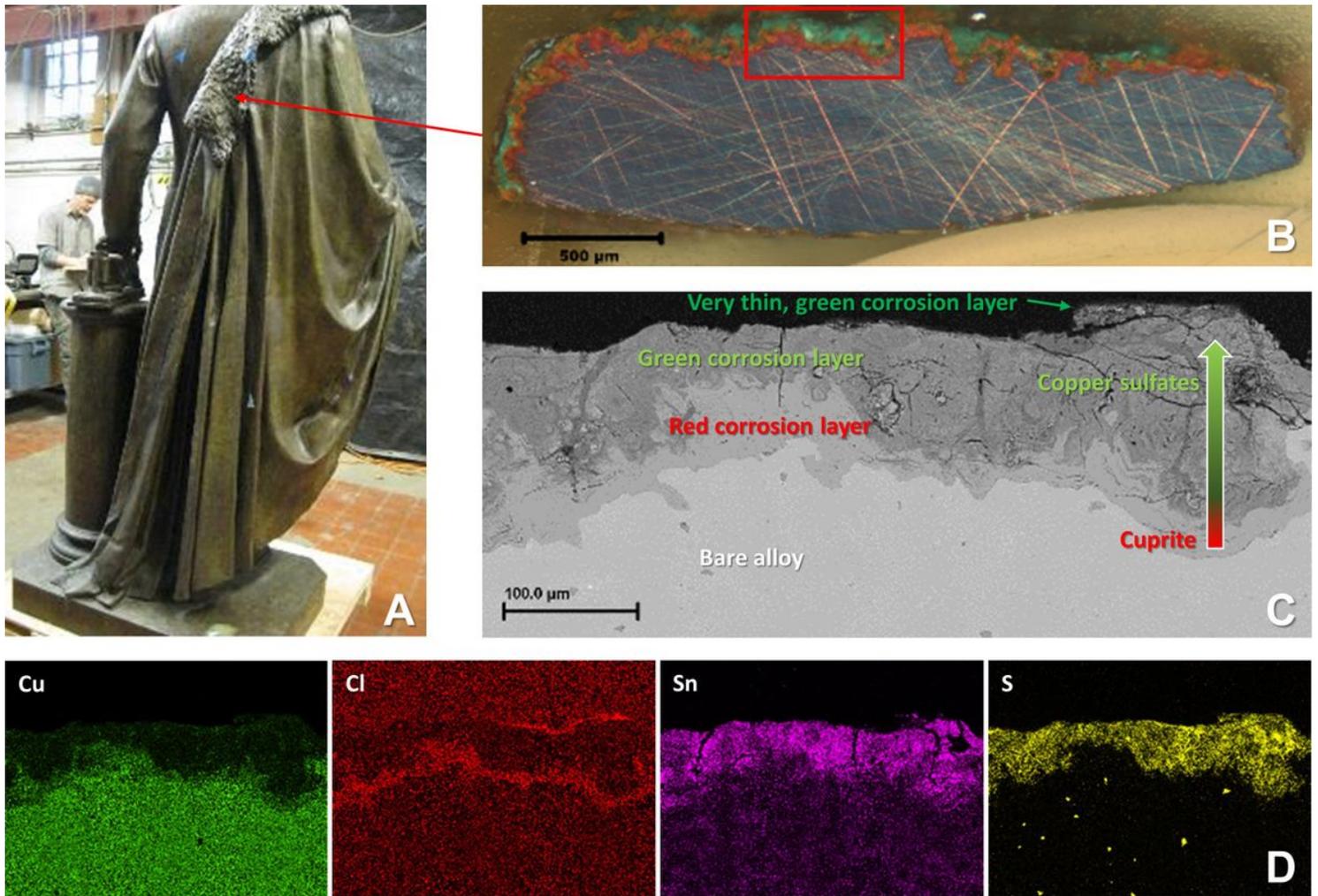


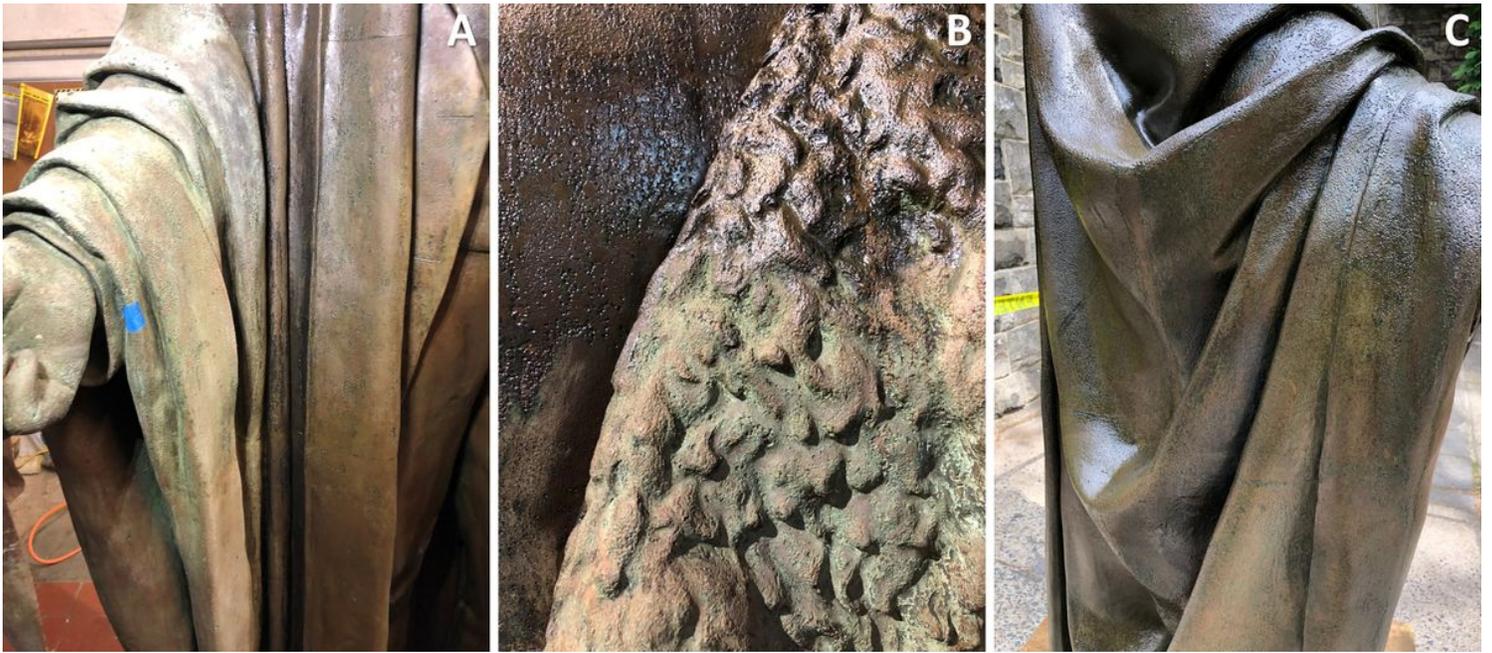
Figure 6

XRD patterns of corrosion products. XRD patterns of two samples of corrosion products, with their sampling locations.



**Figure 7**

Optical microscopy, BSE images, and X-ray elemental maps of cross section S12 (metal alloy and corrosion products). (A) Image of the Morse sculpture in Central Park's conservation laboratory. The red arrow indicates a location on the statue from which a multi-layered sample was removed for analysis of the metal alloy and corrosion products (sample S12). (B) Polarized light micrograph of sample S12 mounted as a cross section. (C) BSE image of the portion of sample S12 indicated by a red square in image B. The red and green arrow indicates changes in composition within the stratigraphy when moving from the main metal alloy up to the surface. (D) X-ray elemental maps of copper (Cu), chlorine (Cl), tin (Sn), and sulfur (S) collected from a portion of cross section S12.



**Figure 8**

Application of new coating. (A-B) Details of the Morse statue collected during the application of the new lacquer coating, including coated and uncoated areas. It is worth noting the contrast between the portion already brushed and the bare surface. (C) Final visual effect upon application of the new coating, after drying, over the green patina.