Green Anionic polymerization of vinyl acetate using Maghnite-Na + (Algerian MMT): Synthesis characterization and reactional mechanism

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Research Article

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Abstract

In this work, the green polymerization of vinyl acetate is carried out by a new method which consists in the use of clay called Maghnite-Na⁺ as an ecological catalyst, non-toxic, inexpensive and recyclable by simple filtration. X-ray diffraction (XRD) showed that Maghnite-Na⁺ is successfully obtained after cationic treatment (sodium) on raw Maghnite. It is an effective alternative to replace toxic catalysts such as benzoyl peroxide (BPO) and Azobisisobutyronitrile (AIBN) which are mostly used during the synthesis of polyvinyl acetate (PVAc) making the polymerization reaction less problematic for the environment. The synthesis reaction is less energetic by the use of recycled polyurethane as container for the reaction mixture and which is considered as a renewable material and a good thermal insulator which maintains the temperature at 0°C for 6h. The reaction in bulk is also preferred to avoid the use of a solvent and therefore to stay in the context of green chemistry. In these conditions, the structure of obtained polymer is established by Nuclear Magnetic Resonance Spectroscopy ¹H NMR and ¹³C NMR. Infrared spectroscopy (FT-IR) is also used to confirm the structure of PVAc. Thermogravimetric analysis (TGA) showed that it is thermally stable and it starts to degrade from 330°C while Differential Scanning calorimetry (DSC) shows that this polymer has a glass transition temperature (T_g = 50°C). The composition in PVAc/maghnite-Na⁺ (7wt% of catalyst) is the most tensile resistant with a force of 182 N and a maximum stress of 73.16 MPa, the most flexible (E = 955 MPa) and the most ductile (ε_r = 768 %)

1. Introduction

The chemistry is now moving towards the use of less polluting polymers and synthetic methods that involve low energy and less toxic reagents and therefore environmentally friendly [1-2]. Polyvinyl acetate (PVAc) meets these requirements, it is non-toxic biodegradable under specific conditions and it received considerable attention in various fields such as the pharmaceutical industry as anti-tumor in cosmetics, in the food and food packaging as the cheese coating and gum base for chewing gum [3-4]. This polymer is also used in the building industry as paint and glue for wood, paper and in the textile industry [5-6]. PVAc can be polymerized using various catalysts. Indeed, polymerized vinyl acetate using Benzoyl Peroxide (PBO) as a catalyst at a temperature of 55°C, while another researchers used Azobisisobutyronitrile (AIBN) in benzene as a solvent at a temperature of 60°C [7-8]. Vinyl acetate was also polymerized in emulsion, the reaction was catalyzed by Potassium Persulfate (KPS) in the presence of surfactants at 70°C [9-10]. However, these catalysts are not recyclable and require treatment of waste, they are toxic and harmful to human health and can be dangerous if swallowed and cause skin and eye irritation [11-12]. In addition, the synthesis of this polymer in presence of these catalysts is carried out at high temperature requiring more energy which is less attractive economically and environmentally [13-14]. In this perspective, we developed in our laboratory a new method for the synthesis of PVAc with a sustainable way based on the principles of green chemistry [15-16]. For these reasons, we used rigid polyurethane as container for the reaction mixture that we recovered and recycled. Moreover, it is a material that remains undamaged for a very long time and considered today as the insulation that has the best thermal insulation

maintaining a low temperature (0°C) for several hours [17]. The advantage of this new method of polymerization is also the use of a catalyst that promotes the synthesis reaction called Maghnite- Na^+ [18-19]. It is easily available, green, non-toxic, inflammable and inexpensive because it can be separated from the system by simple filtration and reused in other reactions [20-21]. Unlike natural clay minerals maghnite- Na^+ has high crystallinity, controllable composition and fewer impurities [22]. For these reasons, the use of such modified maghnite as host materials is expected to be more advantageous than the use of raw natural clay minerals [23]. In the past years, this catalyst has been used successfully as initiator for the anionic polymerization of several acrylamide and vinyl monomers as well as the synthesis of nanocomposites/clay [24-25]. The polymerization of vinyl acetate catalyzed by Maghnite- Na^+ is carried out under mild conditions, in bulk without solvent and with the minimum of reagent, reducing the waste treatment [26-27].

2. Experimental

2.1 Material

The chemicals and reagents used for the synthesis were obtained from commercial sources and were used as received without further purification. Vinyl acetate monomer ($\geq 99\%$) and diethyl ether ($\geq 99\%$) were obtained from Sigma Aldrich. NaCl was purchased from Sigma Aldrich. Raw-Maghnite (Algerian montmorillonite clay) was supplied by BENTAL (Algerian Society of Bentonite). Deionized water was used in the preparation of Maghnite- Na^+ .

2.2. Methods

Nuclear Magnetic Resonance (^1H NMR and ^{13}C NMR) spectra were recorded on a Bruker-Advance 300 MHz apparatus in deuterated chloroform (CDCl_3). FT-IR (Fourier transform infrared spectroscopy) absorption spectra are recorded on an Alpha Bruker FT-IR spectrometer based on Bruker Optics patented Rock Solid™ design, flexible sampling and Transmission, Attenuated Total Reflection (ATR), external and diffuse reflection FT-IR sampling accessories.

Thermogravimetric analyzes (TGA) are recorded on a PerkinElmer STA 6000 under nitrogen in the temperature range 30–690°C with a speed of 20°C/min. X-Ray Diffraction (XRD) studies of the clay are done using Bruker AXS D8 diffractometer with the VÅNTEC-500 detector using Cu-K α radiation ($\lambda = 1.5418 \text{ \AA}$). Differential Scanning Calorimetry (DSC) measurements are carried out on a TA instrument Q500, the heating rate is 10°C/min from -20°C to 200°C under N_2 and the sample weighed about 20 mg. Tensile tests were conducted using a LLOYD LR/10KN Universal Machine at room temperature and crosshead speed of 50mm min $^{-1}$ for the determination of tensile modulus and yield strength, according to the standard ASTM D638.

2.3.Preparation of Maghnite- Na^+

2.4. Synthesis of Poly (vinyl acetate)

$$\begin{array}{ccc}
 \begin{array}{c} \text{H}_2\text{C}=\text{CH} \\ | \\ \text{O}-\text{C}(=\text{O})-\text{CH}_3 \\ \text{Vinyl acetate} \end{array} & \xrightarrow[0^\circ\text{C}, 6\text{H}]{\text{Magnite-Na}^+} & \begin{array}{c} \text{H}_2\text{C}-\text{CH} \\ | \quad | \\ \text{O}-\text{C}(=\text{O})-\text{CH}_3 \\ \text{Poly (vinyl acetate)} \end{array}
 \end{array}$$

3. Results And Discussion

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Scheme 2. Experimental equipment for synthesis of anionic poly (vinyl acetate) catalyzed by maghnite-Na⁺

3.1. Catalyst characterization

Maghnite-Na⁺ is analyzed by FT-IR spectroscopy which confirmed the structure of montmorillonite. A broad band between 3200 cm⁻¹ and 3500 cm⁻¹ characteristic of the OH group bonded to octahedral aluminum is observed in (Figure 1). The band at 1629 cm⁻¹ corresponds to the stretching vibration of H₂O. The FT-IR spectrum also shows the presence of an intense band at 984 cm⁻¹ attributed to the Si-O-Si stretching vibration in the tetrahedral layer. The characteristic bands of Si-O-Al and Si-O-Mg groups appear respectively at 516.46 cm⁻¹ and 439.20 cm⁻¹ [28].

The X-ray powder diffraction profiles (Figure 2) show that there is an increase in basal spacing (d001) from 9.08 Å in the raw-maghnite after 12.70 Å this increase in the interlayer distance is explained by the adsorption of a water molecule on the surface of the sheets of the montmorillonite reflecting the changes in interlayer cation as a result of the basic treatment [29]. The other weak peaks are related to the structure of aluminum-oxygen octahedron and silicon- oxygen tetrahedron in the montmorillonite [30].

3.2. Polymer characterization

The polymerization of PVAc is also established by the Infrared spectroscopy (FT-IR). Indeed, the characteristic bands of the vinyl group at 1646 cm⁻¹, stretching band at 949 cm⁻¹ at 873 cm⁻¹ corresponding to C=C and the stretching vibration at 1132.99 cm⁻¹ corresponding to the group =C-O-C observed in FT-IR spectrum of vinyl acetate (Figure 3) that disappear in FT-IR spectrum of PVAc (Figure 4) confirming that polymerization of the monomer catalyzed by Maghnite-Na⁺ (7% wt) has been successfully completed [31]. In addition, the stretching vibration C=O and the band corresponding to the ester group of vinyl acetate which appear, respectively at 1728.97 cm⁻¹ and at 1206.74 cm⁻¹ shift to 1755.89 cm⁻¹ and 1224.83 cm⁻¹ for PVAc [32]. The stretching vibrations of CH (2923.82 cm⁻¹), CH₂ (2923.82 cm⁻¹), CH₃ (3017.66 cm⁻¹) and COCH₃ (1018.77 cm⁻¹) are observed in the spectra of PVAc [33].

The structure of PVAc is analyzed and confirmed by ¹H NMR and ¹³C NMR spectroscopy. The peaks at 21.04 ppm, 39.16 ppm, 66.98 ppm and 170.38 ppm observed in the ¹³C NMR spectrum of (Figure 6) are

assigned, respectively, to the groups CH₃, CH₂, CH and C=O of the polymer [34]. The ¹H NMR spectrum (Figure 5) also confirms the PVAc structure by the presence of two peaks centered at 1.52 and 1.78 ppm corresponding to the methylene group (CH₂) corroborated by the disappearance of the peaks (two doublets) at 4.56 ppm and at 4.87 ppm corresponding to the double bond H₂C=CH₂ [35]. The peak situated at 2.01-2.05 ppm is assigned to the methyl group (CH₃) and the broad peak located at 4.88 ppm for the CH₂ group. These results are similar to the literature [36].

3.3 .Thermal properties of polymer

The Thermogravimetric analysis is used in order to study the thermal stability of PVAc. The obtained thermogravimetric curve which is the representation of the relative weight loss depending on the heating temperature is given in (Figure 7) and which shows that there are two distinct weight loss stages. The first and intense weight loss between 300 and 400°C (69% wt) corresponds to the deacetylation step which consists in smaller fragments corresponding to the acetic acid mother molecule [37]. During the deacetylation process, (CH)_n fragments evaporate from the polymeric material indicating scission of the polymer main chain at the end [38]. At the temperature of 400°C, PVAc decomposes into a highly regular unsaturated material or polyene [39]. The second weight loss observed at high temperatures, in the range of 400 to 500°C (19% wt) can be attributed to the complete degradation of the formed polyene through chain scission reactions. These results are in agreement with literature [40]

Analysis by Differential scanning calorimetry (DSC) determines the thermal properties of PVAc and in particular its glass transition temperature, T_g, which corresponds to the passage of the polymer from the vitreous state to the rubbery state [41]. The hermetic capsule containing a mass of the polymer is introduced into the oven at room temperature and then the temperature is lowered to -20°C by a flow of (N₂ gas). After that, a temperature gradient of -20°C to 200°C is achieved with a rate of 10°C/min [42]. At 200°C, the temperature is again lowered to -20°C, again with 10°C/min through the N₂ gas flows, before a second temperature rise in the same conditions. The DSC thermogram thus obtained allows the determination of T_g. The first passage in temperature gives an enthalpy relaxation peak which depends on the internal tensions, resulting from the synthesis process and the thermal history of the polymer cannot be taken into consideration [43]. To eliminate the «thermal history» of PVAc, T_g is determined from the second pass. The obtained curve (Figure 8) shows that T_g is about 50°C which is in agreement with the data from the literature [44].

3.4. Mechanical properties of polymer

In general, mechanical properties of materials are strongly influenced by the microstructure. The mechanical properties of the native PVAc were characterized by tensile testing. To evaluate the effect of Maghnite-Na⁺ catalyst (Algerian MMT) in PVAc, we prepared five samples with the same procedure and in different ratios of VAc/Maghnite-Na⁺ (99 :1, 98:2 , 97:3, 95:5, 93:7). The young's modulus and yield strength are greatly enhanced as show in (Table1) . It shows that the mecanicals properties depend on amount of catalyst.

Table 1. Mechanical properties of PVAc /Maghnite-Na⁺ samples

Sample	Catalyst (w %)	Young's modulus (MPa)	Elongation at break (%)	Yield strength (MPa)
Maghnite-Na ⁺	-	845	69.45	-
PVAc (pure)	-	312	45.82	-
PVAc-1/Mag	1	1094	70.02	41.09
PVAc -2/Mag	2	1126	72.54	55.13
PVAc -3/Mag	3	1199	73.95	60.22
PVAc -4/Mag	5	1213	75.36	70.23
PVAc -5Mag	7	1289	86.65	82.17

The tensile test was carried out to evaluate the tensile properties of the various samples compositions in order to determine the influence of the addition of the clay on the tensile properties of the virgin matrix. Young modulus, tensile strength and elongation at break were evaluated as a function of the mass fraction of clay in all series of samples. Thus Young's modulus increased in compositions with the highest clay contents, (1-7%). The composition in PVAc/maghnite-Na⁺ (7w% of catalyst) has the highest tensile values [45]. A Significant increase in the elongation at the break to PVAc/maghnite-Na⁺ was observed. This was attributed to the aggregate formation of maghnite-Na⁺ within the polymer matrix and the matrix with 7wt% maghnite-Na⁺ exhibited an maxima elongation at break of 86.65%. These dramatic improvements were attributed to the fine dispersion of maghnite-Na⁺ in the polymer matrix [46]. This is attributed to the interactions between the polymer chains and the nanometric layers of the clay with a decrease in the value of the young's modulus [47]. This composition is the most tensile resistant with a force of 182 N and a maximum stress of 73.16 MPa, the most flexible (E = 955 MPa) and the most ductile ($\epsilon_r = 768\%$). From these results, it can be deduced that the incorporation of the clay into the PVAc matrix, with different percentages, has significantly improved all of its tensile properties [48].

3.5. Morphological properties of polymer

Scanning electron microscopy SEM was used to see the distribution of the maghnite clay in the polymer synthesized [49]. The morphology of the surfaces was observed in (Fig 9A, B, C and D) and the representative scanning electron micrographs of raw maghnite, maghnite-Na⁺ and PVAc/Maghnite-Na⁺ from the anionic polymerization, respectively. The particle diameter of raw maghnite (Fig 9A) and maghnite-Na⁺ (Fig 9B) ranged from 50 to 100 nm, while PVAc/Maghnite-Na⁺ (Fig 9C) and pure PVAc (Fig 9D) ranged from 150 to 200 nm, the smaller MMT particles were adsorbed around the larger PVAc particles to form the tube structure [50]. SEM images of PVAc/MMT (Fig. 9C) and pure PVAc (Fig. 9D) showed that when the maghnite was incorporated in the PVAc, the average particles size of the composite was increased to 200 nm. The obtained results indicate that the polymerizing chains were aggregated into the clay interlayer regions and consequently the exfoliation of maghnite-Na⁺ (MMT-Na⁺) was almost completed [51].

4. Kinetic'S Studies

The effect of the amount of maghnite on the yield of polymerization expressed by using various weight ratios Mag-Na⁺/monomer, on the polymerization yield of vinyl acetate is showed in (Figure10). The polymerization is carried out in bulk at 0°C for (6h) using various amounts of Mag-Na⁺ (3%, 4%, 6%, 7%, 10% and 12 % wt/wt). It can be noted that the yield of PVAc increases with increasing the amount of the catalyst in which the effect of Mag-Na⁺ as a catalyst is clearly shown [52]. However, the maximum yield is obtained at 7% by weight of Mag-Na⁺ and from this rate it remains unchanged [53]. The phenomenon (increasing of the yield) can be attributed to the active sites available in the catalyst which are responsible for the initiation and acceleration of the polymerization reaction until the saturation of these sites (stabilization of the yield) [54].

(Figure 11) shows the effect of time on the yield of PVAc. The polymerization of vinyl acetate is carried out at different times in bulk at 0°C using (7%) by weight of Maghnite-Na⁺. The obtained results show that the yield increases slightly during four hours of reaction and this can be considered as an induction period [55]. From that time, the yield begins to increase rapidly until it stabilizes and reaching a maximum value of 72.85 % after 6 hours of reaction [56].

5. Proposed Mechanism

This new anionic polymerization consists of three main reactions: (a) initiation, (b) propagation, and (c) termination, as described in (Scheme 3). However, the initiation reaction is generally fast and is not reflected in the overall rate of the polymerization [57]. The interaction of propagating ion pairs with functional groups of the monomer or the polymer chain can affect the propagation rate and in some cases induces side reactions that can cease the polymerization [58]. The termination is brought about intentionally using a suitable electrophile, which can be useful for end group modification [59].

Initiation

Propagation

termination

Scheme 3. Proposed mechanism of poly vinyl acetate (PVAc) catalyzed by maghnite- Na^+ (Algerian MMT)

Conclusion

In this work, poly (vinyl acetate) is synthesized and structurally characterized. The synthesis of this polymer is carried out successfully by highlighting a completely ecological process by integrating the principles of green chemistry. This new approach allows developing an energy-efficient process by using a recycled polyurethane-based as container for the reaction mixture which is an excellent thermal insulator and thus let it possible to work at low temperature for a long time. In addition, PVAc is synthesized for the first time in bulk using an efficient green catalyst called Maghnite- Na^+ under

heterogeneous conditions. The polymerization proceeds via an anionic mechanism due to the presence of intercalated sodium ion (Na⁺) in the lamellar structure of the montmorillonite. FT-IR spectroscopy as well as ¹H NMR and ¹³C NMR spectroscopy confirm the structure of PVAc. TGA shows that the obtained polymer is thermally stable with a degradation temperature higher than 300°C and its glass transition temperature is 50°C which is obtained by DSC analysis. Consequently, Maghnite-Na⁺ showed that it is an attractive eco-catalyst with many advantages such as being cheap, safe and reusable. Indeed, it can be easily separated from the polymer products and regenerated by heating to a temperature above 100 °C.

Declarations

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Conflict of interest statement

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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Figures

Figure 1

FT-IR spectra of (A) Raw-Maghnite and (B) Maghnite-Na⁺

Figure 2

X-Ray Diffraction of: (A) Raw-Maghnite and (B) Maghnite-Na+

Figure 3

FT-IR spectra of vinyl acetate (VAc) monomer.

Figure 4

FT-IR spectra of Poly (vinyl acetate) synthesized.

Figure 6

^{13}C NMR spectrum of PVAc in (CDCl_3)

Figure 7

Thermogravimetric analysis TGA of PVAc catalyzed by Maghnite-Na⁺ (under N₂).

Figure 8

Differential scanning calorimetry (DSC) of PVAc catalyzed by Maghnite-Na⁺ (under N₂)

Figure 9

Scanning electron microscopic (SEM) images of : (A) raw maghnite (B) maghnite-Na⁺ (C) PVAc/maghnite-Na⁺ and (D) pure PVAc.

Figure 10

Effect of the amount of Maghnite-Na+ on the yield of PVAc : Reaction time (6h) , Reaction temperature (0°C), in bulk.

Figure 11

Effect of time on the yield of PVAc (Reaction temperature = 0°C, Mag-Na+ = 7%wt, in bulk)