

Enhancing stability by trapping palladium inside N-heterocyclic carbene-functionalized hypercrosslinked polymers for heterogeneous C-C bond formations

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Abstract

Catalyst deactivation caused by the aggregation of active metal species in the reaction process poses great challenges for practical applications of supported metal catalysts in solid-liquid catalysis. Herein, we developed a hypercrosslinked polymer integrated with N-heterocyclic carbene (NHC) as bifunctional support to stabilize palladium in heterogeneous C-C bond formations. This polymer supported palladium catalyst exhibits excellent stability in the one-pot fluorocarbonylation of indoles to four kinds of valuable indole-derived carbonyl compounds in cascade or sequential manner, as well as the representative Suzuki-Miyaura coupling reaction. Stabilizing mechanism investigations disclose that this catalyst displays a molecular fence effect in which the coordination of NHC sites and confinement of polymer skeleton contribute together to stabilize the active palladium species in the reaction process. This work provides new insight into the development of supported metal catalysts with high stability and will also boost their efficient applications in advanced synthesis.

Background

Palladium (Pd)-catalyzed C-C bond formation reactions, such as Suzuki-Miyaura, Heck, Sonogashira and carbonylation reactions, have been ranked among the most useful and versatile tools for modern organic synthesis¹⁻⁵. Compared with the molecular Pd complex-based homogeneous catalysts, supported Pd catalysts could advance these reactions toward more sustainable conditions due to the easy separation and reuse of the catalyst in the solid-liquid catalysis^{6,7}. The “Cocktail” type nature, in terms of complicated catalyst evolution processes (such as dissolution, aggregation and redeposition) and various metal species (molecular complexes, clusters and nanoparticles), has been found to be involved in these supported Pd catalyzed C-C bond formation reactions^{8,9}. Unlike the common heterogeneous reactions, wherein the reactants chemically absorb on the catalyst surface and undergo the activation/transformation, these reactions experience a Pd dissolution and redeposition process (Fig. 1). Typically, the detachment of Pd from the catalysts by catalytic interaction¹⁰⁻¹², such as the oxidative addition of aryl halides (ArX) to Pd(0) species, initiates the reactions of C-C bond formation in solution^{13,14}, and the redeposition of the regenerated Pd(0) species from reductive elimination back onto the support after the completion of the reaction recovers the supported Pd¹⁵. However, the regenerated Pd(0) species are very unstable and tend to aggregate either in reaction solution or in the process of redeposition, yielding a catalyst quite different from the fresh one¹⁶⁻¹⁸, and then resulting in a decline in the catalytic activity and undesirable recyclability, which poses a major issue for the practical application of supported Pd catalysts in solid-liquid catalysis.

To suppress the Pd aggregation, many nanostructured inorganic and organic supports modified with N-, S- and P-containing groups have been developed to stabilize the active Pd species through enhancing the metal-support inter-actions^{19,20}. However, these approaches mainly focused on the control of the Pd aggregation over the catalyst surface, and the Pd aggregation in solution has not been effectively prohibited, which still inevitably led to a remarkable deactivation and frustrated recyclability although

these Pd aggregates could be redeposited onto the catalyst surface^{10,11,21-26}. Thus, the development of a novel support, which can not only efficiently recapture the dissolved Pd before their aggregation in solution but also stabilize the resupported Pd species during the redeposition process, is highly desired to improve the stability of the supported Pd catalysts.

The facilely accessible knitting aryl network polymer (KAP) integrated with strong electron-donating coordination sites are expected to meet the above requirements²⁷, since these hypercrosslinked KAPs have cage-like polymer skeletons which are supposed to work as molecular fences to efficiently isolate the active Pd species and prevent them from migration/aggregation^{28,29}, while a built-in monomer bearing electron-rich group can effectively recapture and stabilize the dissolved Pd species³⁰. Notably, N-heterocyclic carbenes (NHCs) have recently been recognized as outstanding ligands to coordinate with molecular metal complexes or electronically modify surface metal nanoparticles for efficient homogeneous and heterogeneous catalysis³¹⁻³³. Therefore, it can be speculated that the integration of confinement and coordination effect enabled by the NHC-based hypercrosslinked KAPs could lead to the development of a series of new supported Pd catalysts with enhanced stability for heterogeneous C-C bond formations in solid-liquid catalysis.

The carbonylation reactions represent an important class of C-C bond formation reactions which can introduce carbonyl group into molecules and produce a variety of functional compounds³⁴. While ArX are widely used in Pd-catalyzed carbonylation reactions, direct and heterogeneous carbonylation of unfunctionalized Ar-H has been rarely investigated due to the inertness of aryl C-H bond and poor catalyst stability under harsh and oxidative conditions³⁵. In this work, we have developed a series of KAPs supported Pd complex precatalysts with different crosslinking structures (KAP-Pd-PEPPSI-x, x = 1, 2, 3, Fig. 2a), which showed remarkable activities in carbonylation of indoles to corresponding valuable acyl fluoride, amides, aroyl azides and aroyl cyanides with significantly enhanced stability, surpassing other conventional supported Pd catalysts. Detailed stabilizing mechanism investigations revealed that the KAP-Pd-PEPPSI displayed a molecular fence effect in which the coordination of the NHC sites and confinement of the hypercrosslinked polymer skeleton worked together to trap the active Pd species inside the swollen polymer in the reaction process, which effectively decreased the Pd content in solution and therefore prevented them from aggregation. Likewise, KAP-Pd-PEPPSI also showed superior catalytic stability over traditional Pd@C in the representative Suzuki-Miyaura reaction. This work not only provides a facile entry to develop stable heterogeneous catalyst for solid-liquid catalysis, but also establishes an efficient route to synthesize various indole-derived carbonyl compounds which have wide applications in the preparation of bioactive compounds.

Results And Discussion

Catalyst synthesis and characterization. First, the NHC precursor-functionalized KAPs with different crosslinking structures were prepared by Friedel-Crafts alkylation of chloromethyl functionalized NHC precursor with tetramethylbenzene, triphenylmethane and tetraphenylmethane in 1,2-dichloroethane

under the catalysis of anhydrous FeCl_3 in high yields²⁸, and KAP-Pd-PEPPSI-x ($x = 1, 2, 3$) were then synthesized by treatment of the above NHC precursor-functionalized KAPs with PdCl_2 in 3-chloropyridine with over 90% yields (Fig. 2a and 2b)³⁶. They exhibited good swelling capabilities in organic solvents, and the swelling degree of KAP-Pd-PEPPSI-1, KAP-Pd-PEPPSI-2 and KAP-Pd-PEPPSI-3 in specific dimethyl formamide (DMF) measured by equilibrium swelling method were 239%, 171% and 127%, respectively, which are inversely correlated with their degree³⁷. The high crosslinking degree enabled by the polymeric skeletons of KAP-Pd-PEPPSI-2 and KAP-Pd-PEPPSI-3 are supposed to provide better molecular fences to stabilize the involved Pd species than KAP-Pd-PEPPSI-1.

The structural properties of KAP-Pd-PEPPSI were characterized using various spectroscopic measurements. The fourier transform infrared (FT-IR) spectra show bands at 1667, 1465 and 1447 cm^{-1} attributed to the benzene skeleton vibrations (Supplementary Fig. 1)²⁸. The peaks at 2860-2970 cm^{-1} are assignable to the alkyl C-H stretch vibration of NHC groups, indicating their incorporation into polymer skeleton³⁸. The chemical shifts of ^{13}C cross polarization magic angle spinning (CP/MAS) nuclear magnetic resonance (NMR) spectroscopy located at 25, 29, 49, 53 and 66 ppm are ascribed to aliphatic methyl, isopropyl, benzyl, imidazolynyl and the quaternary carbon core of tetraphenyl methane (Supplementary Fig. 2), respectively³⁹, and the broad peaks in the range of 120 to 160 ppm reveal the presence of aromatic subunits within the polymer networks⁴⁰. These results confirm the integrated structure of NHC-based KAPs. Moreover, the Pd contents determined by inductively coupled plasma mass spectrometry (ICP-MS) in KAP-Pd-PEPPSI-1, KAP-Pd-PEPPSI-2 and KAP-Pd-PEPPSI-3 are 0.24, 0.23 and 0.28 mmol g^{-1} , respectively. The Pd 3d X-ray photoelectron spectroscopy (XPS) spectra of KAP-Pd-PEPPSI show peaks at 337.56 (Pd $3d_{5/2}$) and 342.81 eV (Pd $3d_{3/2}$), indicating that the Pd in KAP-Pd-PEPPSI are present in +2 oxidation state and strong electron donation from NHC to the Pd(II) exists (Supplementary Fig. 3)⁴¹.

Scanning electron microscopy (SEM) images show that all the KAP-Pd-PEPPSI are consisted of large irregular particles on the order of a few microns with a relatively smooth surface (Fig. 2c and Supplementary Fig. 4a-c), and no metal or salt aggregates are found on the catalyst surface. Besides, the scanning transmission electron microscopy (STEM) images and energy disperse spectroscopy mappings of KAP-Pd-PEPPSI reveal that the Pd and N components are uniformly distributed throughout the catalyst (Fig. 2e-g and Supplementary Fig. 5). It is interesting to find that the porosities of KAP-Pd-PEPPSI are highly dependent on their crosslinking degrees. As seen in Supplementary Fig. 6, the Brunauer-Emmett-Teller (BET) surface area in KAP-Pd-PEPPSI-1 is 227 $\text{cm}^3 \text{g}^{-1}$, which decreases to 10 $\text{cm}^3 \text{g}^{-1}$ and 7 $\text{cm}^3 \text{g}^{-1}$, respectively, in KAP-Pd-PEPPSI-2 and KAP-Pd-PEPPSI-3 with higher crosslinking degree. Although the hypercrosslinking structure results in the decrease in the number of intrinsic pores, the swelling properties of KAP-Pd-PEPPSI in solid-liquid condition benefit the access of reactants to active sites⁴², which can be illustrated in Fig. 2d and Supplementary Fig. 4d-f that a large number of small cracks are formed and the surface roughness of these particles increases markedly after swelling in DMF. In this case, the unique elimination of the large intrinsic pores enabled by hypercrosslinked polymer skeleton and the controlled

swelling properties of KAP-Pd-PEPPSI may produce an enhanced spatial confinement effect, which is believed to be crucial for stabilizing the active Pd species with synergetic coordination with NHC sites^{43,44}.

To systematically understand the confinement and coordination effects on the stability of KAP-Pd-PEPPSI catalysts, other supported Pd metal catalysts with or without N, P heteroatom doping, such as Pd@C, N-doped carbon supported Pd nanoparticles (Pd@NC)⁴⁵, P-doped carbon supported Pd nanoparticles (Pd@PC)⁴⁶, porous silica supported Pd nanoparticles (Pd@SiO₂)⁴⁷, as well as supported Pd complex precatalysts with phosphine and NHC as coordinative site, such as PPh₃-functionalized KAP supported Pd(II) complex (KAPs-Pd-PPh₃) and porous organic polymer with built-in Pd(II)-NHC complex (POPs-Pd-PEPPSI)^{28,48}, have been prepared according to the reported procedures, and the detailed synthetic processes are given in the attached methods and supplementary information (Supplementary Fig. 7 and Supplementary Table 1).

Fluorocarbonylation of Indole. Acid halides are an important class of carbonyl compounds which are widely used as acyl reagents or intermediates to produce numerous carbonyl compounds⁴⁹. Conventional synthesis of acid halides from transformations of C-H bonds usually undergo carboxylation or oxidation to generate the carboxylic acids, and then react with halogenating reagents, such as PCl₃, thionyl or oxalyl chloride, to produce the corresponding acyl chlorides, which generate a large amount of wastes⁵⁰. Note that Pd-catalyzed carbonylations of aryl halides have been developed to be facile alternatives to synthesize acid halides⁵¹⁻⁵³. Particularly, Andersen et al. achieved an unprecedented breakthrough in establishing carbonylation of aryl halides to produce acid chlorides via sterically encumbered ^tBu₃P and CO coordination enabled reductive elimination of (^tBu₃P)(CO)Pd(COAr)Cl⁵¹. On the other hand, direct C-H functionalization has greatly boosted the carbonylation methods in direct preparation of carboxylic compounds without using prefunctionalized substrates like aryl halides⁵⁴. However, to the best of our knowledge, carbonylation of C-H bonds to acid halides has not been reported. Compared with the acid chlorides (C-Cl bonds, ~78 kcal/mol), acid fluorides are more thermodynamically stable (C-F bonds, ~105 kcal/mol) and can be readily handled for further transformation under simple conditions⁵⁵. We herein select carbonylation of indole to corresponding acid fluoride as model reaction to investigate the catalytic activity and stability of KAP-Pd-PEPPSI, and this one-pot transformation is designed to proceed via oxidative iodination of C-H bond first and then carbonylation.

Initial screening of reaction conditions was conducted with Pd-catalyzed fluorocarbonylation of 1-methyl-1H-indole (**1a**) using KAP-Pd-PEPPSI-2 as the precatalyst (Supplementary Table 2). After a series of primary investigations, the yield of indole-3-carbonyl fluoride (**A**) reached 89% when using CsF as the fluorine source and DABCO as the base (Supplementary Table 2, entry 5). Under the optimized reaction conditions as presented in Table 1, all Pd-based heterogeneous catalysts provided **A** in high yields (82-93%, entries 2-10), indicating for the first time the availability of synthesizing acid halides by heterogeneous carbonylation. Surprisingly, a relatively lower yield (69%) was obtained when using

molecular Pd-PEPPSI as homogeneous control catalyst (entry 1), and the formation of Pd black in the solution was observed (Supplementary Fig. 8a), suggesting the significant challenge of this reaction condition to catalyst stability. Notably, the Pd contents in reaction solutions were tested by ICP-MS to be 5.5%-91% of the total Pd when treating different supported Pd catalysts with the initial reactants for 4 h in the absence of CsF (Washing Test) or keeping yield of **A** at low levels (Supplementary Table 3)¹⁴, however, only a very small portion of Pd remained in solution after the completion of the reactions (Table 1). These results clearly demonstrate the “Cocktail” type nature of supported Pd-catalyzed carbonylation reactions, in which a Pd dissolution and redeposition process is involved¹². The STEM morphologies of used catalysts revealed that severe Pd aggregation and uneven dispersion were detected on Pd@C, Pd@NC and Pd@SiO₂ (Supplementary Fig. 8b, 8c and 8e). Notably, the hierarchical porous Pd@PC with P-doping was found to be more stable than other supported Pd catalysts, but the formation of Pd aggregates was still not avoided (Supplementary Fig. 8d). Gratifyingly, the organic polymer supported Pd complex catalysts exhibited much better stabilities than the Pd catalysts over inorganic supports. As shown in Supplementary Fig. 9, no obvious changes compared to the fresh catalysts were observed in the STEM images of KAP-Pd-PPh₃, POP-Pd-PEPPSI and KAP-Pd-PEPPSI after reaction for 24 h, and the Pd components were homogeneously distributed throughout these catalysts without formation of detectable nanoparticles or aggregates. It can be primarily deduced that the polymer supports with built-in electron-donating P- or NHC-ligands can prevent the *in-situ* formed active Pd species from aggregation in this Pd dissolution and redeposition process.

Then, we selected KAP-Pd-PPh₃, POP-Pd-PEPPSI and KAP-Pd-PEPPSI-2 which have similar skeletal structures to further study the recyclability of the supported Pd complex catalysts. As shown in Fig. 3a, all these catalysts could be used at least four times without a significant decline in the activity. However, gradual activity decreases were observed both in KAP-Pd-PPh₃ and POP-Pd-PEPPSI when further increasing the reaction cycles, and obvious Pd particle size changes evidenced the deactivation of these two catalysts via metal aggregation (Fig. 3c and Supplementary Fig.10). To our delight, KAP-Pd-PEPPSI-2 could be recycled for nine times without significant loss of activity (Fig. 3a). STEM images of spent KAP-Pd-PEPPSI-2 demonstrated that the catalyst after the 5th use was very similar to the fresh one (Fig. 3c and Supplementary Fig.11). It retained a good dispersion of Pd and no large-sized nanoparticles or aggregates were observed. Moreover, the composition and chemical state of the surface Pd on the above catalysts after the 5th cycle were investigated by XPS, revealing that there exist three types of Pd species with Pd(0) and Pd(II) oxidation states on the spent catalysts generated from KAP-Pd-PPh₃, POP-Pd-PEPPSI and KAP-Pd-PEPPSI-2. The content of the oxidized Pd(II) species is 68.7% (yellow areas) on the supported catalyst originated from KAP-Pd-PPh₃, while only about 7.2% of Pd(0) clusters which are electronically coordinated by phosphine ligand, and 24.1% of naked Pd(0) nanoparticles stabilized by the polymer skeleton, are observed with their Pd 3d_{5/2} core level peaks located at 335.09 eV and 335.60 eV, respectively⁵⁶. Whereas, the supported Pd catalyst from KAP-Pd-PEPPSI-2 bearing similar polymer skeleton to KAP-Pd-PPh₃ has different Pd composition, revealing high content of Pd(0) clusters (72.6%, blue areas) coordinating with the NHC sites⁵⁷. These results suggest that the NHC site demonstrates

much better electron-donating effect, compared with phosphine, in stabilizing Pd species during their dissolution into reaction solution and redeposition onto the hypercrosslinked polymeric catalyst. Another POP-Pd-PEPPSI with the same built-in NHC sites but differently structured polymer skeleton, showed much lower content (17.5%, blue curves) of the NHC coordinated Pd(0) clusters, indicating that the polymeric skeleton also plays a very important role on stabilizing Pd species. It can be preliminarily inferred from above results that the hypercrosslinked KAP with built-in NHC sites acts as ideal molecular fence, wherein the polymer skeleton and the strong electron-donating NHC work synergistically to suppress the Pd aggregation in the Pd dissolution and redeposition process.

Investigation on Stabilizing Effect. Following the above basic understanding, the coordination influence of NHC-site was investigated by using two structurally similar KAP supports, KAP-4 and KAP-2 with or without incorporated NHC precursor as shown in Fig. 4a, to interact with the *in-situ* formed Pd species from PdCl₂ in the model fluorocarbonylation of indole. It was found that, although the PdCl₂/KAP-4 catalytic system gave a 83% yield of acid fluoride **A** in the fresh reaction, only trace amount of the product was obtained in its second use, which could be explained by its STEM images showing the formation of large aggregates with size up to hundreds of nanometers on the catalyst surface (Fig. 4c and Supplementary Fig. 12) and the Pd 3d XPS spectra which support the formation of Pd(0) aggregates with Pd 3d_{5/2} core level peaks at 335.60 eV (Fig. 4b). These results suggest a poor metal-support interaction between KAP-4 and dissolved Pd species, and it also can be experimentally deduced from its second use that this types of electron-deficient Pd(0) aggregates show very low activity toward this one-pot carbonylation of indole. On the contrary, the PdCl₂/KAP-2 catalytic system provided a 73% yield of **A** in its first use, and a higher **A** yield (89%) was obtained in the second reaction cycle. The corresponding STEM images revealed a good dispersion of Pd components and no obvious aggregates were detected (Fig. 4c and Supplementary Fig. 12). Besides, the deconvoluted Pd 3d peaks at 335.09 eV and 340.34 eV confirmed the formation of NHC-coordinated Pd(0) clusters (Fig. 4b). These findings indicate that the electron-rich NHC sites generated via deprotonation of the NHC precursor in KAP-2 can electronically donate to and then effectively stabilize the dissolved Pd species via coordination interaction in the reaction process.

Having figured out the electronic coordination effect of NHC sites, we next investigated the influence of the catalyst structure on the stability of active Pd species. For this purpose, three KAP-Pd-PEPPSI-x (x = 1, 2, 3) precatalysts with different crosslinking degrees (Fig. 1a) were used to investigate their recyclabilities in carbonylation of indole at about 50% yield of acid fluoride **A**. As shown in Fig. 4d, a gradual loss of activity was observed in low-crosslinking KAP-Pd-PEPPSI-1 with a chain structure, and the STEM images of the spent catalyst after the 6th cycle showed large aggregates up to 100 nm (Fig. 4e), indicating a severe particle aggregation on the catalyst surface. KAP-Pd-PEPPSI-2 with its NHC moiety linked by triphenylmethane group, which has a higher crosslinking degree than that of KAP-Pd-PEPPSI-1, delivered an enhanced catalytic stability. The STEM images of the spent KAP-Pd-PEPPSI-2 showed small nanoparticles without the formation of large aggregates (Figure 4e). KAP-Pd-PEPPSI-3 which has the highest crosslinking degree among these three catalysts, showed even higher stability than KAP-Pd-

PEPPSI-2. The catalyst could be used at least nine times without loss of activity and no any observable nanoparticles were detected on the STEM images of used catalyst (Figure 4e). The above results indicate that the highly hypercrosslinking moieties of the KAP skeleton work like a molecular fence to confine the *in-situ* formed active Pd species inside each polymeric domain, which benefits the recapture of the dissolved Pd species by the support via electronic interaction/coordination and prevents the supported small Pd clusters from aggregation on the support surface^{28,58,59}. Therefore, it can be concluded from these two parts of studies that the synergetic cooperation of the polymeric confinement and the electronic NHC coordination inside the hypercrosslinked KAP skeleton contributes to the efficient Pd stabilization during their dissolution into the reaction mixture and redeposition back onto the KAP catalyst.

Investigation on Carbonylation Process. To get more insights into this KAPs-supported Pd catalyzed carbonylative C-C bond formation reaction under solid-liquid condition, the time-course analyses of Pd evolution and distribution as well as corresponding reaction activity in fluorocarbonylation of indole were carried out using typical Pd@C and the optimized KAP-Pd-PEPPSI-3 as the catalyst, respectively. As shown in Fig. 5a, the KAP-Pd-PEPPSI-3 experienced a long activation time to reduce the polymeric Pd(II)-NHC precatalyst to the active NHC-coordinated Pd(0) clusters, the formation of which could be testified by the complete prohibition of the reaction after adding mercury (mercury poisoning experiment)^{9,60}, and **A** started to be generated with its yield rapidly increasing to 87% in the next six hours (blue curve with dots). On the other hand, the concentration of Pd species in solution slowly increased in the first four hours, and it reached a maximum of 5.5% of total Pd when **A** started to be generated and then rapidly decreased (blue curve with squares). Such an obvious mismatch between the Pd content in solution and the product yield indicates a fast redeposition of the regenerated Pd(0) onto the polymeric support. Therefore, the Pd species involved in the carbonylation were kept recycling via following the process of initial detachment from the polymer catalyst into solution, catalyzing carbonylation and redeposition back onto the polymer without deactivation, and thus the very active Pd species avoided accumulating in the solution and were maintained at a dilute concentration, which is significantly helpful to suppress Pd aggregation⁶¹. Conversely, the typical Pd/C displayed a different Pd distribution, the Pd content in solution increased rapidly to about 37.3% of total Pd in about two hours after the reaction started (purple curve with squares), and then remained at a high concentration level (28.8-40.7%) for over six hours with the yield of **A** increasing to about 80% (purple curve with dots in Fig. 5a and Supplementary Table 4). Although most of the dissolved Pd (only 2.2% of total Pd remained in solution as present in Table 1) redeposited back onto the carbon support, many of them have grown into bigger sized Pd aggregates with ununiform dispersion as evidenced by STEM (Supplementary Fig. 8b). Compared to the conventional supported catalysts, this polymeric catalyst with swelling property could efficiently trap the active Pd species inside the skeletal domain by the confinement of the polymeric skeleton and coordination of the NHC sites, keeping the dissolved Pd species at a low concentration level in solution. Moreover, the hot filtration test of the KAP-Pd-PEPPSI-3 catalyst at its peak Pd concentration in solution only gave a 30% yield of **A** after the filtrate reacted for 24 h (red line in Fig. 5b), and Pd black was clearly

observed, indicating a fast deactivation of the dissolved Pd without the stabilization of polymer skeleton and NHC.

A stabilizing mechanism based on the above experimental results is proposed (Fig. 5c). First, the built-in polymeric Pd(II) complex is *in-situ* reduced to NHC coordinated Pd(0) clusters, wherein the NHC group acts as an electron-donating ligand that increases the electron density of Pd(0), thereby stabilizing the Pd(0) clusters from aggregation⁶². Next, the oxidative addition of Pd(0) species into the C-I bond of 3-iodo-1-methylindole, generated by iodination of C3-H bond of indole, leads to the detachment of Pd from the clusters, affording the representative Pd(II) intermediate [Indole]PdI confined inside each of the hypercrosslinked polymer domains¹³. Subsequently, CO insertion, fluoride exchange and reductive elimination give product **A** with regeneration of electron-deficient Pd(0) species, which are then recaptured onto the polymeric skeleton mainly by electronic interaction with NHC sites to recover the supported Pd(0) clusters for next reaction cycle. As a result, the hypercrosslinked KAPs with built-in NHC sites function as molecular fences to effectively trap the active Pd species inside the polymer during the reaction process and reload them after completion of one catalytic cycle, which prevent the Pd species from accumulating in solution and the formation of large inactive aggregates.

Applications in Other C-C Bond Formations. Besides above carbonylative synthesis of acid fluoride, carbonylations of aryl halides with hydrogen, alcohol, amine, organometallic compound, have been extensively used to produce aldehyde, ester, amide and ketone³⁴. However, some unconventional nucleophiles, such as weakly nucleophilic aromatic amines (especially NH-heteroarenes)⁶³, strongly coordinating CN⁻⁶⁴ and intolerable N₃⁻⁶⁵ have been scarcely used for carbonylative transformation due to their inherent limitations, and corresponding heterogeneous carbonylation procedures have not been reported yet. But nevertheless, the corresponding carbonyl compounds of above nucleophiles are important building blocks in the syntheses of fine chemicals and bioactive compounds⁶⁶⁻⁶⁸. Based on the establishment of KAP-Pd-PEPPSI catalyzed direct fluorocarbonylation to access acid fluoride, we subsequently investigated carbonylations of indole with above three challenging nucleophiles to synthesize indole-derived acylamides, aroyl azides and aroyl cyanides in one-pot cascade or sequential manner (Eq. 1), respectively.

We first tried to develop a direct aminocarbonylation of indole with weakly nucleophilic N-H of pyrazole via the combination of fluorocarbonylation and subsequent acylation in a cascade reaction manner without isolation of acid fluoride intermediate in a single reactor. It is necessary to note that the conventional direct aminocarbonylation of C-H bond with N-H moiety took selectivity challenge, because it needs an oxidative condition which inevitably leads to the oxidative carbonylation of N-H bonds, giving undesirable urea. Herein, initial screenings indicated that the Arndtsen's chlorocarbonylation methodology with tetrabutyl ammonia chloride as Cl⁻ resource by specific Pd(PtBu₃)₂ catalyst, and Manabe's fluorocarbonylation procedure with KF as F⁻ reagent by Pd(OAc)₂/xantophos catalyst, only provided 14% and 11% yield of product **3aa**, respectively (Supplementary Table 5), suggesting that these two catalytic procedures afforded low yields of corresponding acid chloride and acid fluoride from carbonylation of *in-*

situ formed iodindole. Whereas, a 74% yield of **3aa** was obtained upon using our KAP-Pd-PEPPSI-3/CsF catalytic system (Table 2 and Supplementary Table 5), and no urea byproduct was observed. The corresponding control experiments also revealed that the *in-situ* generated A could readily react with pyrazole (**2a**) to give the corresponding indole-3-amide, while no fluorocarbonylation or aminocarbonylation occurred in the absence of fluoride or replacing fluoride with chloride, bromide or iodide (Supplementary Table 6).

Under the optimal reaction conditions, the substrate scope was explored. The aminocarbonylations of N-methyl indoles, bearing a substituent such as -Me, -Cl, -Br and -OMe on the benzene ring, all proceeded smoothly to produce the corresponding indole-3-amides (**3ba-3ia**) in 57-91% yields. The reaction of pharmaceutically interesting substrate, 1-methyl-1H-pyrrolo[2,3-b]pyridine⁶⁹, gave **3ja** in 70% yield. Indoles with different N-protective groups were also well-tolerated, both N-allyl indole and N-benzyl indole showed good reactivity in the one-pot synthesis of **3ka** and **3ma**. The cascade reaction of N-methyl pyrrole instead of indole also took place efficiently, selectively giving monocarbonylated **3la** in 62% yield. Other substituted pyrazoles, imidazole, benzimidazole and aniline derivatives were also good nucleophiles for this carbonylation, providing **3ab-3aj** in 59-85% yield. Importantly, KAP-Pd-PEPPSI-3 demonstrated outstanding recyclability in this cascade process under more complicated conditions than that of fluorocarbonylation, and could be used at least five runs without any loss of activity in the synthesis of **3aa** (77%).

Carbonylation of the non-prefunctionalized substrate with N_3^- as the nucleophile (azidocarbonylation) provides a direct access to versatile aroyl azide, wherein two challenges exist: 1) the aroyl azide product is prone to react with commonly used tertiary phosphine ligand in carbonylation reactions to afford the corresponding benzoyl azaylide (known as the Staudinger reaction)⁷⁰; 2) the N_3^- coordinated to Pd can be converted into NCO^- in the presence of CO even under mild reaction conditions⁷¹. Taking advantage of the present phosphine-free KAP-Pd-PEPPSI catalytic system, we tried a one-pot and two-step procedure to establish this azidocarbonylation transformation, first fluorocarbonylation of indole to afford acid fluoride and then discharging excess CO for azidation with NaN_3 . Following this sequential procedure (Table 3), azidocarbonylation of N-methyl indoles bearing different substituents provided **4a-4i** in 49-92% yield. The reaction of 1-methyl-7-azaindole with a pyridine moiety also proceeded smoothly, giving a 69% yield of **4j**. The investigations on the substituent effect revealed that the electron-donating groups are more favorable for this carbonylation (**4k-4n**). When 1-benzyl-1H-indole was employed, **4o** was obtained in 81% yield. Besides, N-H indole could also furnish the corresponding product in 64% yield (**4p**) by prolonging the reaction time to thirty hours. Last but not the least, KAP-Pd-PEPPSI-3 could be used at least five runs without loss of its catalytic activity (83% yield), and the easily formed isocyanate or other byproducts were not detected by this one-pot sequential procedure.

Carbonylation with cyanide anion as the nucleophile (cyanocarbonylation) could simultaneously introduce two functional groups (C=O and $C\equiv N$) for diverse transformations, and the challenges therein are the poisoning of the catalyst by strong coordination of CN^- to Pd and byproduct formation via

coupling between ArX with CN⁻ without CO incorporation. Following the above one-pot sequential procedure, we have established a cyanocarbonylation of indole to prepare corresponding indole-3-carbonyl nitrile. As shown in Table 4, this protocol has good tolerance to diverse functional groups at different positions of N-methyl indole, giving the corresponding indole-3-carbonyl nitriles (**5a-5i**) in good to excellent yields (60-85%). 1-Methyl-7-azaindole was also a good substrate for this carbonylation, giving **5j** in 55% yield. The reaction of various substituted N-allyl indole also proceeded smoothly to afford the desired **5k-5m** in 57-75% yields. Using pharmaceutically interesting N-benzyl indole as the substrate⁷², **5n** was obtained in 77% yield. Besides, N-H indole also gave the corresponding acylcyanide **5o** in good yield. Moreover, KAP-Pd-PEPPSI-3 displayed remarkable reusability and provided a high yield of **5a** in its 5th use.

Besides efficient and stable applications in above carbonylative C-C bond formations, the optimized KAP-Pd-PEPPSI-3 also displayed excellent recyclability with up to 21 cycles and >95% product yield in another representative Suzuki-Miyaura coupling reaction at low catalyst loading (0.4 mol%) (Supplementary Fig. 13). STEM images of KAP-Pd-PEPPSI-3 after the 21st cycle showed uniformly distributed nanoparticles without formation of large aggregates (Supplementary Fig. 14c and 14d). XPS results revealed the presence of a strong metal-ligand interaction between Pd nanoparticles and NHC groups (Supplementary Fig. 15). On the contrary, typical Pd@C displayed a significant decrease in activity after the 6th cycle and the formation of large particle aggregates (up to 500 nm in size) suggests the severe Pd aggregation during the reaction (Supplementary Fig. 14a and 14b). These results further confirm the super stability of hypercrosslinked KAP-Pd-PEPPSI-3 and highlight its potential applications in other C-C bond formation reactions.

In summary, we have developed a series of efficient and highly stable supported Pd catalysts for different C-C bond formations under solid-liquid catalysis. For the first time, the synthesis of acid halide by heterogeneous carbonylation was established via KAP-Pd-PEPPSI catalyzed fluorocarbonylation of indole with enhanced catalytic stability surpassing other conventional supported Pd catalysts. The stabilizing mechanism studies revealed that the hypercrosslinked polymer skeleton of the catalyst worked like a molecular fence to effectively trap the dissolved Pd species inside the swollen polymer by the electronic NHC coordination and polymeric confinement in the reaction process, which facilitated the reloading of active Pd species back onto the skeletal support and avoids their accumulation in reaction solution, thus suppressing the catalyst deactivation caused by Pd aggregation. Moreover, the optimized KAP-Pd-PEPPSI-3 also showed excellent activities and durabilities in other three types of one-pot fluorocarbonylation/acylation reactions, providing effective procedures to access unconventional carbonyl derivatives (e.g. heterocyclic acylamide, aroyl azides and aroyl cyanides) that are valuable compounds for functional applications and diverse transformations. Likewise, KAP-Pd-PEPPSI-3 also displayed excellent recyclability in representative Suzuki-Miyaura coupling reaction at low catalyst loading. This work provides an effective strategy for the development of metal supported catalysts with high stability, which will also boost their efficient applications in advanced synthesis under solid-liquid catalysis

Methods

Synthesis of KAP-Pd-PEPPSI-1. 1,3-Bis(2,6-diisopropyl-4-(chloromethyl)phenyl)-4,5-dihydro-1H-imidazol-3-iumchloride (524 mg, 1 mmol), anhydrous FeCl_3 (1.95 g, 12 mmol) and 1,2,4,5-tetramethylbenzene (134 mg, 1 mmol) were mixed together in anhydrous 1,2-dichloroethane. The mixture was refluxed at 80 °C under N_2 for 24 h, and the resultant precipitates were collected by filtration and washed continually with methanol until the filtrate was clear and then treated by Soxhlet extraction in methanol for 24 h. The obtained reddish brown powder was dried under vacuum at 100 °C for 12 h to afford the crosslinked polymer with built-in NHC precursor (KAP-1, 68% yield). Then, a mixture of KAP-1 (384 mg), PdCl_2 (25 mg, 0.14 mmol) and anhydrous K_2CO_3 (345 mg, 2.5 mmol) in 5 mL of 3-chloropyridine was refluxed for 16 h, after which the precipitates were collected by filtration and washed successively with DMF, water and methanol. Further drying under vacuum at 60 °C for 24 h gave the KAP-Pd-PEPPSI-1 in 94% yield.

Synthesis of KAP-Pd-PEPPSI-2. The synthetic procedure was similar to that of KAP-Pd-PEPPSI-1. Briefly, a mixture of 1,3-Bis(2,6-diisopropyl-4-(chloromethyl)phenyl)-4,5-dihydro-1H-imidazol-3-iumchloride (524 mg, 1 mmol), anhydrous FeCl_3 (1.95 g, 12 mmol) and triphenylmethane (163 mg, 0.67 mmol) in anhydrous 1,2-dichloroethane was refluxed at 80 °C under N_2 for 24 h to afford the KAP-2 in 72% yield, then the coordination of PdCl_2 with the NHC sites in KAP-2 delivered the KAP-Pd-PEPPSI-2 in 91% yield.

Synthesis of KAP-Pd-PEPPSI-3. The synthetic procedure was similar to that of KAP-Pd-PEPPSI-1. Briefly, a mixture of 1,3-Bis(2,6-diisopropyl-4-(chloromethyl)phenyl)-4,5-dihydro-1H-imidazol-3-iumchloride (524 mg, 1 mmol), anhydrous FeCl_3 (1.95 g, 12 mmol) and tetraphenylmethane (160 mg, 0.5 mmol) in anhydrous 1,2-dichloroethane was refluxed at 80 °C under N_2 protection for 24 h to afford the KAP-3 in 82% yield. Subsequently, the complexation of PdCl_2 with the NHC sites in KAP-2 produced the KAP-Pd-PEPPSI-3 in 91% yield.

Synthesis of KAP-Pd-PPh₃. Triphenylphosphine hydrobromide (229.9 mg, 0.67 mmol), anhydrous FeCl_3 (1.95 g, 12 mmol) and 1,4-Bis(chloromethyl)benzene (175.1 mg, 1 mmol) were mixed together in anhydrous 1,2-dichloroethane. The mixture was refluxed at 80 °C under N_2 for 24 h, and the resultant precipitates were collected by filtration and washed continually with methanol until the filtrate was clear and then treated by Soxhlet extraction in methanol for 24 h. The obtained reddish brown powder was redispersed in 10 mL MeOH and KOH (56 mg, 1 mmol) in 5 mL MeOH were slowly added. After stirring this mixture at room temperature for 2 h, the solid was collected by filtration and dried under vacuum to afford the crosslinked polymer with built-in phosphine (KAP-PPh₃, 89% yield). Then, a mixture of KAP-PPh₃ (384 mg) and bis(acetonitrile)palladium(II) chloride (36 mg, 0.14 mmol) in 10 mL CH_2Cl_2 was stirred at room temperature for 10 h, after which the precipitates were collected by filtration and washed with MeCN and CH_2Cl_2 . Further drying under vacuum at 60 °C for 24 h gave the KAP-Pd-PPh₃ catalyst in 92% yield.

Synthesis of POP-Pd-PEPPSI. 1,3-bis(4-iodo-2,6-diisopropylphenyl)-4,5-dihydro-1H-imidazol-3-ium chloride (488 mg, 0.72 mmol) and 1,3,5-triethynylbenzene (72 mg, 0.48 mmol) were dissolved in a mixture of toluene (150 mL) and diisopropylamine (6 mL). After adding bis(triphenylphosphine)palladium(II) chloride (30 mg, 0.042 mmol) and CuI (30 mg, 0.159 mmol), the reaction mixture was stirred at 90 °C for 24 h, after which the solid was retrieved by filtration and washed with excess acetone and methanol. The resultant polymer was dried under vacuum for 24 h at 60 °C (66% yield). Then, the complexation of PdCl₂ with the NHC sites in above polymer delivered the POP-Pd-PEPPSI in 96% yield.

Synthesis of Pd@PC. A mixture of sucrose (12.5 g), phytic acid solution (5 g) and silica colloid (20 g) in 50 g ultrapure water was stirred at room temperature for 12 h to give a homogeneous slurry. After being dried at 100 °C for 8 h, the obtained solid was further heated at 160 °C for 8 h. The obtained black lumps were grounded into fine powder and then transferred into a tube furnace for carbonization under nitrogen. The temperature was ramped linearly from r.t. to 600 °C at a heating rate of 2 °C min⁻¹, then raised to 900 °C at a heating rate of 5 °C min⁻¹ and maintained at 900 °C for another 3 h. Upon cooling to room temperature, the obtained solid (denoted as carbon@silica) was treated with hydrofluoric acid (10 wt%) to remove the silica nanoparticles and the resultant powder was washed thoroughly with ultrapure water. After being dried under vacuum at 100 °C for 12 h, P-doped mesoporous carbon (PC) was obtained. Then, 1.0 g of the PC was added into an impregnation solution of PdCl₂ (34 mg, 0.28 mmol) in HCl solution (20 mL, 0.5 wt %), and the mixture was stirred at 80 °C for 12 h. The resultant solid was collected and dried at 100 °C for 24 h, and further reduced under a flow of H₂ at 200 °C for 2 h (with a heating rate of 2 °C min⁻¹) to afford the Pd@PC.

Synthesis of Pd@NC. A mixture of sucrose (12.5 g), cyanamide (5 g) and silica colloid (20 g) in 50 g ultrapure water was stirred at room temperature for 12 h to give a homogeneous slurry. The procedures of drying, carbonization, template removing, impregnation and reduction were similar to that of PC.

Synthesis of Pd@SiO₂. PdCl₂ (34 mg, 0.28 mmol) was dissolved in HCl solution (20 mL, 0.5 wt %), then 1.0 g of the large pore silica gel (380-550 m²g⁻¹) were added into this impregnation solution. After stirring at 80 °C for 12 h, the resulting solids were dried at 100 °C for 24 h and calcined at 400 °C in air for 4 h. The resultant catalysts were later reduced under a flow of H₂ at 200 °C for 2 h (with a heating rate of 2 °C min⁻¹).

Materials characterization. ¹³C CP/MAS NMR measurements were performed using a Bruker 400 M MAS system with Adamantane as internal reference for C. Liquid NMR spectra were recorded using a Bruker Avance TM spectrometer operating at 400 MHz for ¹H and 100 MHz for ¹³C. XPS were carried out on a ESCALAB 250Xi X-ray photoelectron spectroscopy. BET surface area was performed on Micromeritics ASAP 2020 Physisorption analyzer. Before analysis, the catalysts were degassed at 120 °C for 6 h. The morphologies of catalyst were observed by SEM (Tecnai G2 F20 S-TWIN). STEM patterns were obtained using a FEI F20 with an acceleration voltage of 200 kV. FT-IR were performed on a Thermo Fisher

Scientific Nicolet 6700 spectrometer. The contents of Pd in catalysts and solution were measured with Thermo Fisher ICAP RQ instrument.

Calculation of swelling degree. Swelling degrees of KAP-Pd-PEPPSI in DMF are calculated based on equilibrium swelling method.⁵

$$Q = \left(\frac{W1}{\rho1} + \frac{W2}{\rho2} \right) / \frac{W1}{\rho1}$$

Here, $W1$ and $W2$ are specimen weights before and after swelling in DMF at 80 °C for 72 h, and ρ represents the densities of polymer ($\rho1$) and DMF ($\rho2$) at room temperature. The swelling degree of KAP-Pd-PEPPSI-1, KAP-Pd-PEPPSI-2 and KAP-Pd-PEPPSI-3 calculated here are 239%, 171% and 127%, respectively.

General procedure for fluorocarbonylation of indole. A mixture of **1a** (0.25 mmol), I_2 (0.5 mmol), DABCO (0.5 mmol) in 1 mL DMF was added into a 100 mL stainless steel autoclave with a magnetic bar. The mixture was stirred at room temperature for 60 minutes under N_2 protection. Then CsF (1 mmol), catalyst (5 mol%), DMF (1 mL) were added successively. Then the closed autoclave was vacuumed and purged with CO three times before it was finally pressurized with 2.0 MPa CO. Then the reactor was immersed in an oil bath preheated at 80 °C. After reaction for 24 h, the mixture was cooled down before excess CO was discharged. Finally, the resultant reaction mixture was analyzed by HPLC (Agilent 1260 Infinity with an Agilent ZORBAX C18 column, methanol/ H_2O = 90/10, 0.8 mL/min, λ = 268 nm) using biphenyl as an inner standard.

For the recycling study in Fig. 2, the fluorocarbonylation of indole was performed with the same reaction condition as described above except using recycled catalyst. After each completion of the reaction, the solid catalyst was collected and washed with 20 mL DMF, 20 mL water and 20 mL methanol to remove the organic and inorganic residues. Finally, the resultant catalyst was dried under vacuum at room temperature for 24 h, weighted, and reused in the next run.

For the recycling study in Fig. 3, the amount of **1a** and I_2 were increased to 0.5 mmol and 0.6 mmol while keeping other parameters unchanged. After reaction for 16 h, the solid catalyst was collected and washed with DMF, water and methanol to remove the organic and inorganic residues. Finally, the collected catalyst was dried under vacuum at room temperature for 24 h, weighted, and reused in the next run.

General procedure for aminocarbonylation of indoles. A mixture of indole (0.25 mmol), I_2 (0.5 mmol), DABCO (0.5 mmol) in 1 mL DMF was added into a 100 mL stainless steel autoclave with a magnetic bar. The mixture was stirred at room temperature for 60 minutes. Then CsF (1 mmol), catalyst (5 mol%), amine (1.25 mmol), DMF (1 mL) were added successively. Then the closed autoclave was vacuumed and purged with CO three times before it was finally pressurized with 2.0 MPa CO. Then the reactor was immersed in an oil bath preheated at 80 °C. After 24 h, the mixture was cooled down before excess CO

was discharged. Finally, the resultant reaction mixture was purified by chromatography using silica gel (petroleum ether and ethyl acetate) to afford the corresponding product. The obtained products were qualitatively analyzed by liquid NMR (Zhongke-Niujiu 400) and high resolution mass spectra (HRMS, Thermo Fisher Scientific LTQ FT Ultra with DART positive mode).

General procedure for azidocarbonylation of indoles. A mixture of indole (0.25 mmol), I_2 (0.5 mmol), DABCO (0.5 mmol) in 1 mL DMF was added into a 100 mL stainless steel autoclave with a magnetic bar. The mixture was stirred at room temperature for 60 minutes. Then CsF (1 mmol), catalyst (5 mol%), DMF (1 mL) were added successively. Then the closed autoclave was vacuumed and purged with CO three times before it was finally pressurized with 2.0 MPa CO. Then the reactor was immersed in an oil bath preheated at 80 °C. After 24 h, the mixture was cooled down before excess CO was discharged. Then NaN_3 (1 mmol) was added directly into the resulted mixture with no need for purification. The mixture was stirred at room temperature for 7 h. Finally, the resultant reaction mixture was purified by chromatography using silica gel (petroleum ether and ethyl acetate) to afford the corresponding product. The obtained products were qualitatively analyzed by liquid NMR and HRMS.

General procedure for cyanocarbonylation of indoles. A mixture of indole (0.25 mmol), I_2 (0.5 mmol), DABCO (0.5 mmol) in 1 mL DMF was added into a 100 mL stainless steel autoclave with a magnetic bar. The mixture was stirred at room temperature for 60 minutes. Then CsF (1 mmol), catalyst (5 mol%), DMF (1 mL) were added successively. Then the closed autoclave was vacuumed and purged with CO three times before it was finally pressurized with 2.0 MPa CO. Then the reactor was immersed in an oil bath preheated at 80 °C. After 24 h, the mixture was cooled down before excess CO was discharged. Then TMSCN (1 mmol) and CsF (1 mmol) were added directly into the resulted mixture with no need for purification. The mixture was stirred at room temperature for 7 h. Finally, the resultant reaction mixture was purified by chromatography using silica gel (petroleum ether and ethyl acetate) to afford the corresponding product. The obtained products were qualitatively analyzed by liquid NMR and gas chromatography in combination with mass spectrometry (GC-MS, Agilent 5975C/7890A equipped with a HP-5MS column).

Data availability. The data supporting the findings of this study are available within the article and its Supplementary Information files.

Declarations

Author contributions

C. Y. and Q. X. conducted the experiments, analyzed the results and wrote the manuscript. F. L. designed the research, supervised the project, and edited the manuscript. P. Sun and H. L. participated in writing the manuscript. Z. Z. performed the SEM experiments and data analysis.

Competing interests

The authors declare no competing financial interests.

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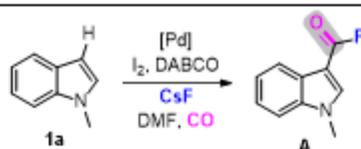
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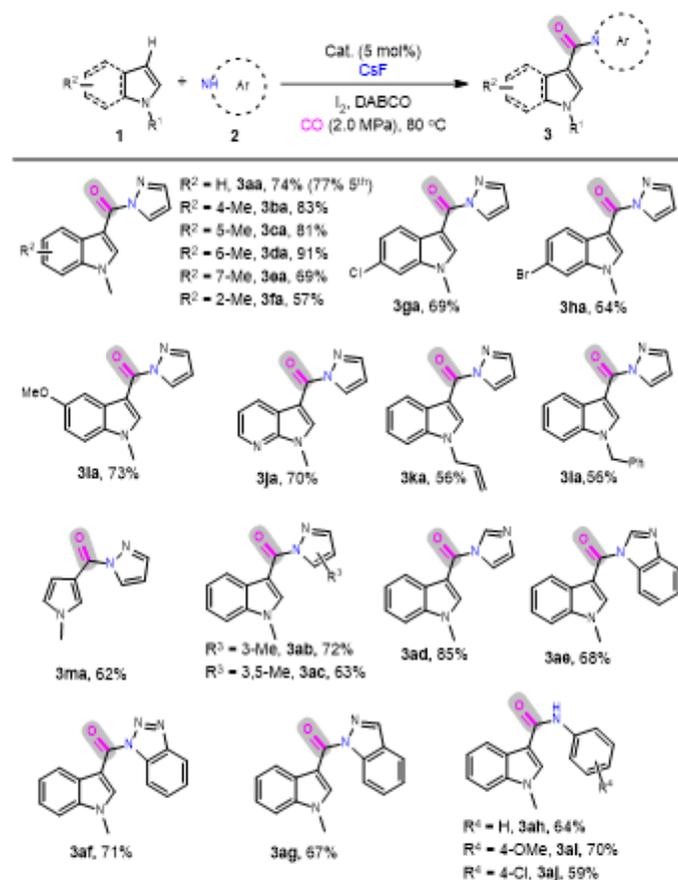
Tables

Table 1. Catalytic performances of supported Pd catalysts^a

Entry	Catalyst	S _{BET} (m ² g ⁻¹)	Yield ^b (%)	Pd leaching ^c (%)
1	Pd-PEPPSI		69	
2	Pd@C	619	93	2.2
3	Pd@NC	478	91	1.6
4	Pd@PC	557	87	1.3
5	Pd@SiO ₂	482	82	1.0
6	KAP-Pd-PPh ₃	473	91	0.8
7	POP-Pd-PEPPSI	458	90	0.5
8	KAP-Pd-PEPPSI-1	227	89	0.5
9	KAP-Pd-PEPPSI-2	10	89	0.4
10	KAP-Pd-PEPPSI-3	7	87	0.2

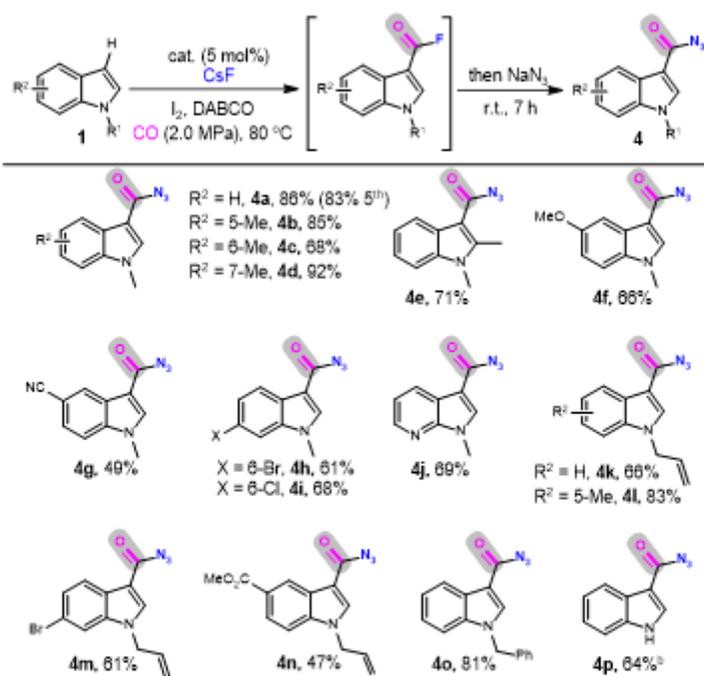
^aReaction conditions: **1a** (0.25 mmol), CsF (1 mmol), Pd catalyst (5 mol%), I₂ (0.5 mmol), base (0.5 mmol), DMF (2 mL), 80 °C, 24 h, 2.0 MPa. ^bHPLC yield. ^cPd content in the solution after completion of the reaction.

Table 2. Aminocarbonylation of different indoles with aromatic amines^a.



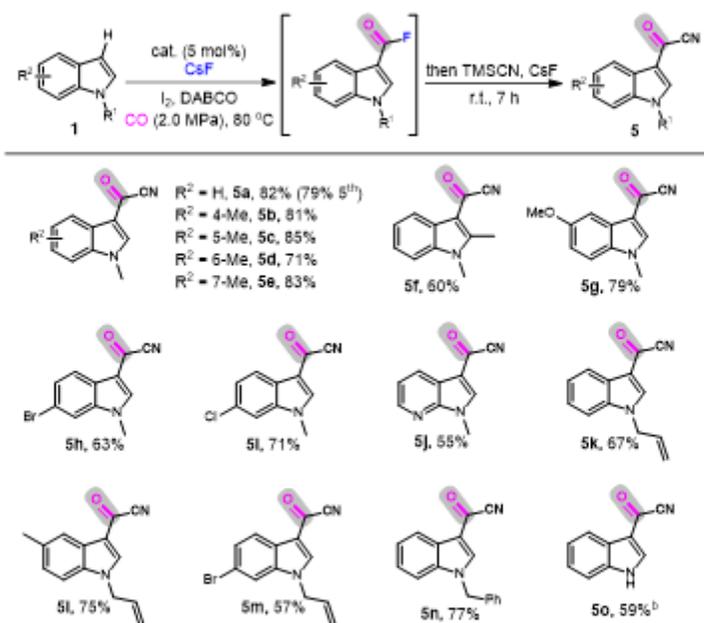
^aReaction conditions: KAP-Pd-PEPPSI-3 (5 mol%), indole (0.25 mmol), amine (1.25 mmol), I₂ (0.5 mmol), DABCO (0.5 mmol), CsF (1.0 mmol), CO (2.0 MPa), DMF (2.0 mL), 80 °C, 24 h. Isolated yield.

Table 3. Azidocarbonylation of indoles to synthesize indole-3-carbonyl azides^a.



^aReaction conditions: KAP-Pd-PEPPSI-3 (5 mol%), indole (0.25 mmol), I₂ (0.5 mmol), DABCO (0.5 mmol), CsF (1.0 mmol), DMF (2.0 mL), CO (2.0 MPa), 80 °C, 24 h. Then NaN₃ (1.0 mmol) was added, room temperature, 7 h. ^b30 h. Isolated yield.

Table 4. Cyanocarbonylation of indoles to synthesize indole-3-carbonyl cyanides^a.



^aReaction conditions: KAP-Pd-PEPPSI-3 (5 mol%), indole (0.25 mmol), I₂ (0.5 mmol), DABCO (0.5 mmol), CsF (1.0 mmol), DMF (2.0 mL), CO (2.0 MPa), 80 °C, 24 h. Then TMSCN (1.0 mmol) and CsF (1.0 mmol) were added, room temperature, 7 h. ^b30 h. Isolated yield.

Figures

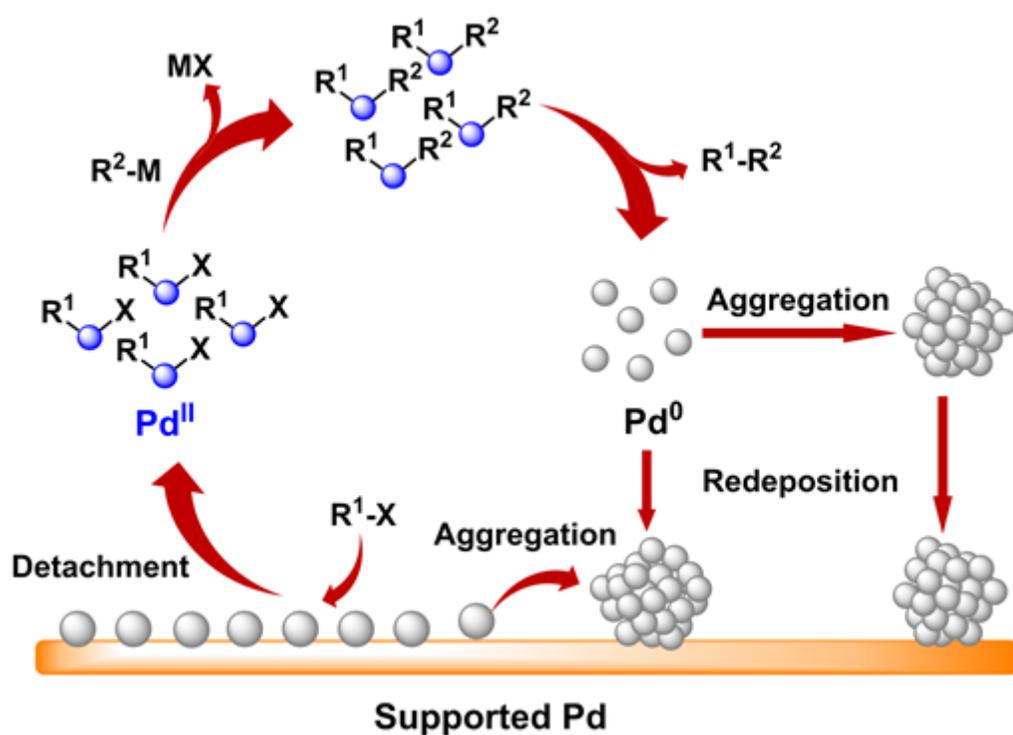


Figure 1

Schematic illustration of the “Cocktail” type nature and evolution of the catalytic species in C-C bond formation reactions catalyzed by supported Pd catalysts.

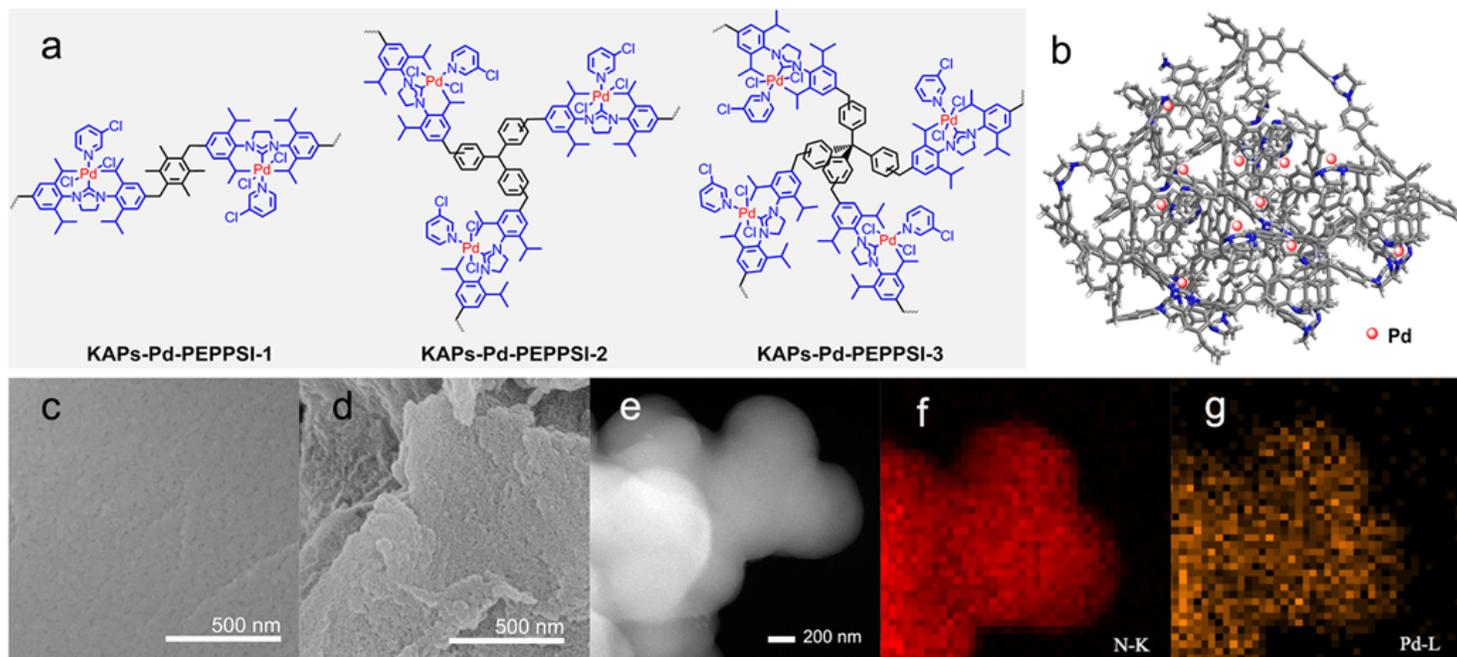


Figure 2

Structure and characteristics of KAP-Pd-PEPPSI. (a) Structures of KAP-Pd-PEPPSI with different degrees of crosslinking. (b) Simulation model for hypercrosslinked KAP-Pd-PEPPSI-x (x = 2, 3). (c) SEM image of KAP-Pd-PEPPSI-2. (d) SEM image of KAP-Pd-PEPPSI-2 after swelling in DMF for 24 h at 80 °C. (e) STEM image of KAP-Pd-PEPPSI-2. (f-g) EDS mapping images for N and Pd in KAP-Pd-PEPPSI-2.

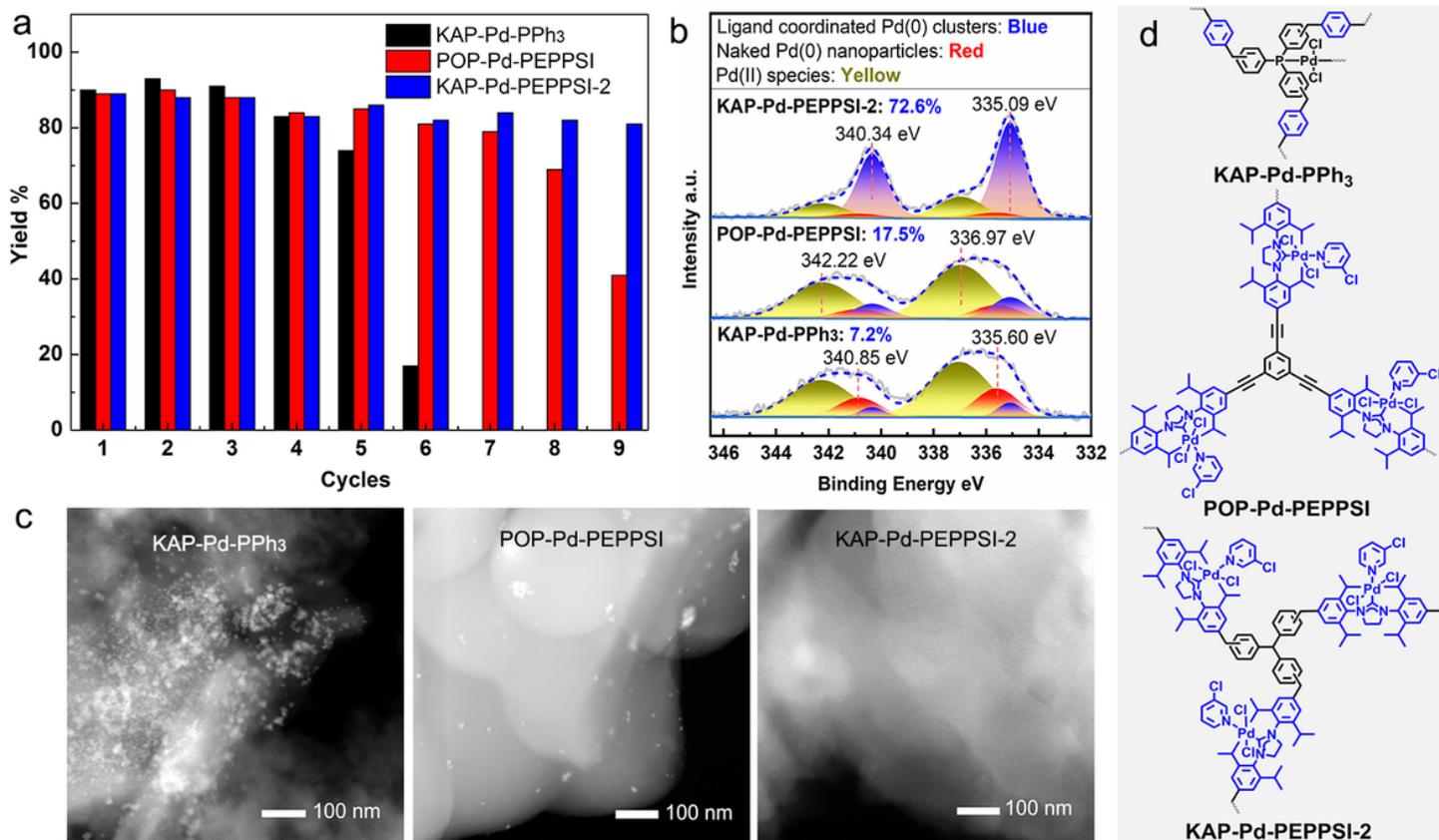


Figure 3

Recyclability of the selected catalyst. (a) Recyclability of the supported Pd complex precatalysts in fluorocarbonylation of indole. Reaction conditions: 1a (0.25 mmol), CsF (1 mmol), catalyst (5 mol%), I₂ (0.5 mmol), base (0.5 mmol), DMF (2 mL), 80 °C, 24 h, 2.0 MPa. (b) Pd XPS spectra and (c) STEM images of different catalysts after being used for five times. (d) Structures of the catalysts used here.

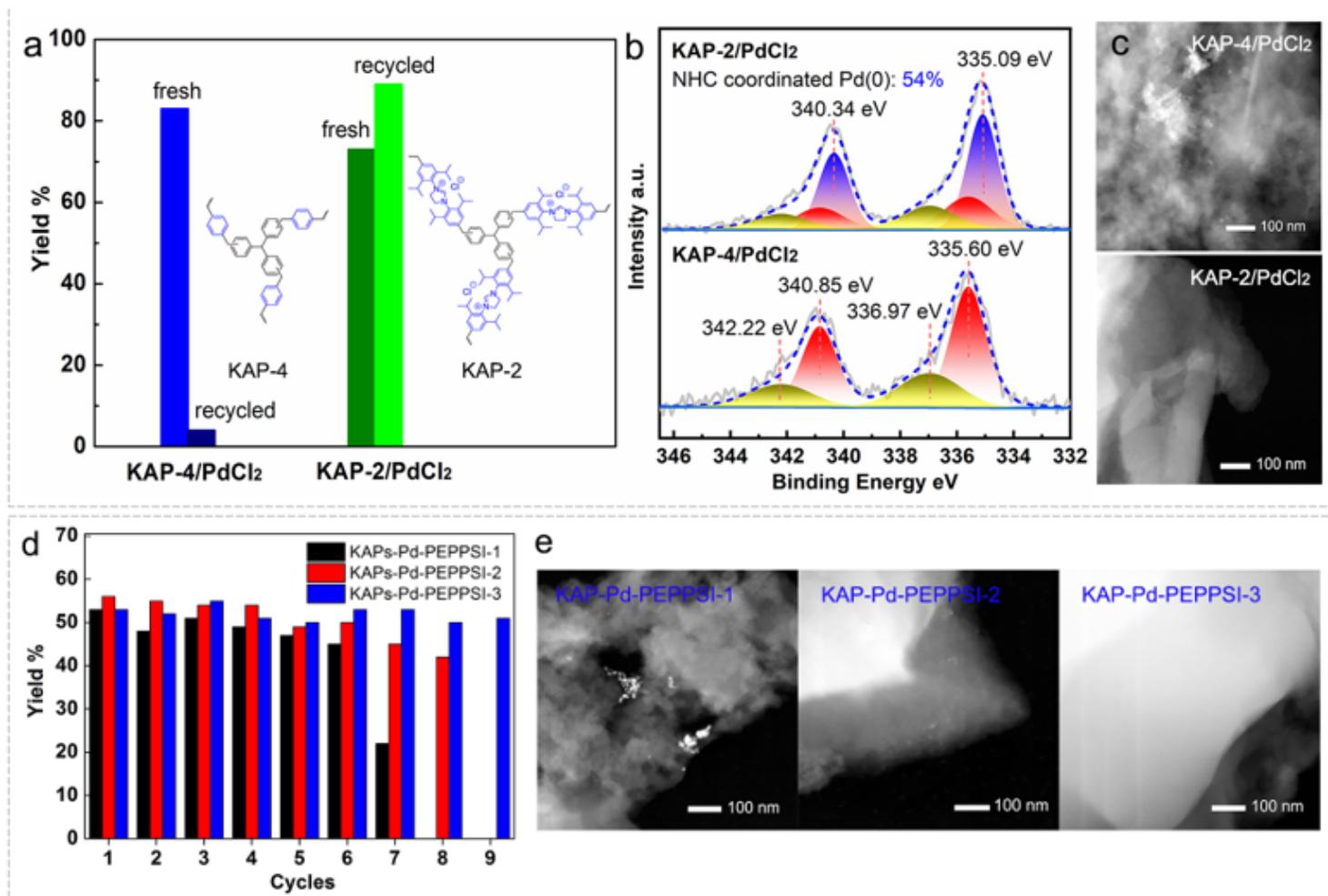


Figure 4

Stabilizing effect of KAP. (a) Control experiments to determine the stabilizing effect of NHC under the optimized reaction conditions. (b) Pd XPS spectra and (c) STEM images of recycled KAP-2/PdCl₂ and KAP-4/PdCl₂. (d) Control experiments to determine the stabilizing effect of catalyst structure. Reaction conditions: 1a (0.25 mmol), CsF (1 mmol), Pd catalyst (2.5 mol%), I₂ (0.3 mmol), base (0.5 mmol), DMF (2 mL), 80 °C, 16 h, 2.0 MPa. (e) STEM images of different KAP-Pd-PEPPSI-x after the 6th cycle.

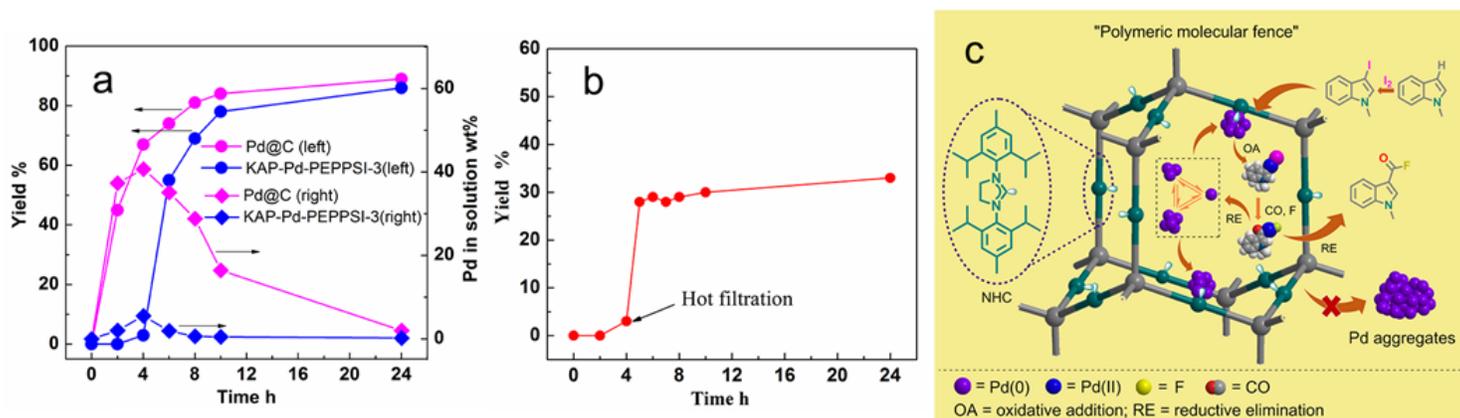


Figure 5

Reaction process investigation. (a) Kinetic investigation. (b) Hot filtration test. Under the optimized reaction conditions, the reaction mixture was quickly divided into two parts by centrifugation after reaction for 4 hours, and the upper clear liquid proceeded to react for another 20 h in the presence of CsF (c) Proposed stabilizing mechanism for KAP-Pd-PEPPSI-3 in the fluorocarbonylation of indole.

Supplementary Files

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