

Origin of Life: Aliphatic aldehydes in the Earth's crust – remains of prebiotic chemistry?

Yildiz Großmann

University of Duisburg-Essen

Ulrich Schreiber

University of Duisburg-Essen

Christian Mayer

University of Duisburg-Essen

Oliver Schmitz (✉ oliver.schmitz@uni-due.de)

University of Duisburg-Essen

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Abstract

The Origin of Life is a question that has not yet been solved in the natural sciences. Some promising interpretative approaches are related to hydrothermal activities. Hydrothermal environments contain all necessary elements for the development of precursor molecules. There are possibly catalytically active surfaces and wide ranges of pressure and temperature conditions. The chemical composition of hydrothermal fluids together with periodically fluctuating physical conditions should open up multiple pathways towards prebiotic molecules. Already in 2017, we detected prebiotic organic substances, including a homologous series of aldehydes in more than 3 billion years old Archean quartz crystals from Western Australia. In order to approach the question if the transformation of inorganic into organic substances is an ongoing process, we investigated a drill core from the geologically young Wehr caldera in Germany at a depth of 1000 m. Here we show the existence of a similar homologous series of aldehydes (C₈ to C₁₆) in the fluid inclusions of the drill core calcites, a finding that supports the thesis that hydrothermal environments could possibly be the material source for the origin of life.

Introduction

Naturally occurring aliphatic aldehydes may play an important role as prebiotic precursor molecules. With growing chain lengths, the so-called fatty aldehydes gain amphiphilic properties and could serve as starting points for a large collection of membrane-forming lipids (Brown et al. 2009, Brown et al. 2011). The reactivity of the aldehyde head group allows a wide variety of chemical modifications, including oxidation to fatty acids, reduction to alcohols, the formation of amines and amides and finally ester or ether formation with glycerol (Fritz et al. 2013). They still play a role in the recent biosynthesis of sphingolipids where a long-chain aldehyde is undergoing a condensation reaction with serine (Awasthi et al. 1980).

The biochemical formation of long aliphatic chains with a hydrophilic head group is quite complicated and involves a considerable amount of energy, it is not very likely to have occurred as initial development. On the other hand, early metabolic processes required functional membranes at an early point of time. Therefore, during the early steps of life, they were very likely depending on an external occurrence of such molecules. All theories on the origin of protocells which are based on initial membrane formation require an abiotic source of amphiphilic components (Deamer et al. 2002, Sakuma et al. 2015, Deamer 2017).

Recently, aliphatic aldehydes have been discovered in carbon-rich chondritic meteorites (Aponte et al. 2019). Aponte et al. analyzed ten different carbonaceous chondrites and, using a novel analytical approach, detected and quantified 16 different aldehydes. Among them were butanal, pentanal and hexanal in concentrations around 2-6 nmol/g. Most aldehydes in chondrites exhibit a clear enrichment of ¹³C isotopes compared to their biogenic counterparts, a finding which indicates their abiotic formation (Aponte et al. 2019). This result is highly interesting because it represents convincing proof for an abiotic (in this case, extraterrestrial) origin of such potentially amphiphilic compounds which are possible starting points for membrane-forming lipids. The analytical method used to analyze the aldehydes was

based on derivatization with (S,S)-(-)-1,4-dimethoxy-2,3-butanediol ((S,S)-DMB-diol). The analysis ended after 95 min and the derivatized hexanal was detected at a retention time of approximately 93 min; therefore, any longer-chain aldehydes present were not observable.

Moreover, there are also clear indications for a terrestrial abiotic formation of aldehydes. A complete family of homologous aliphatic aldehydes of even and uneven chain lengths was detected inside Archean quartz crystals (Schreiber et al. 2017). Inclusions in these Archean quartz crystals grown in a hydrothermal environment contained eleven different species from heptanal up to heptadecanal. This terrestrial evidence is even more intriguing concerning early membrane formation because of the wide and continuous distribution of chain lengths and the fact that the chain functionalization occurs almost exclusively on the methyl end groups of the chain. Due to the selected analytical conditions, the analysis of short-chain aldehydes was not possible here. Results from stable isotope analysis of the methane content of the inclusions are in clear accordance with an abiotic formation of the detected hydrocarbons (Schreiber et al. 2017).

It is likely that analyses using both analytical methods would detect both short-chain and long-chain aldehydes in meteorites and quartz samples. This assumption is supported by the results of the recent paper by Mißbach et al. (Mißbach et al. 2021). The authors used chromatographic methods to investigate the constituents of fluid inclusions in 3.5-billion-year-old barites, which - just like the quartz samples already studied in 2017 by Schreiber et al. - originated from Western Australia. It is particularly noteworthy that the authors were only able to detect short-chain aldehydes here, namely ethanal, propanal, pentanal and heptanal, based on the analytical parameters.

Thus, it could be shown that a prebiotic presence of short- and long-chain aldehydes is not exclusively due to meteorite impacts on the early Earth but also due to hydrothermal formations, probably based on Fischer-Tropsch type reactions.

Altogether, the abiotic formation of aldehydes may be a common feature of planetary bodies in - at least - our solar system. At this point, one has to ask the question if it is an ongoing process that occurs up to the present day. If it is, one should be able to find evidence for abiotically formed aliphatic aldehydes in recent hydrothermal environments. To minimize contamination by metabolites from microbial life in a corresponding search, sampling should occur outside of densely populated environments. On the other hand, it should focus on materials that are in close contact with hydrothermal fluids and that could be able to collect possible organic products over an extended period of time.

For that purpose, a drilling project was started in the phonolite/trachyte complex of the Wehr caldera (Eifel mountains) southwest of Cologne, near the lake "Laacher See" (50°25'35.139" N, 7°13'10.4322" E). In this volcanic environment, we expect a rich flow of hydrothermal fluids consisting of water and carbon dioxide as the bulk solvents. Under conditions of larger depths in the Earth's crust, we assume a large variety of organic compounds to form (Schreiber et al. 2017). Among others, Fischer-Tropsch type chemistry should lead to aliphatic chains, which are expected to undergo partial oxidation on their methyl end groups, eventually leading to a homologous series of aldehydes (Durham et al. 2010, Xiang et al.

2016). At approximately 1 km of depth, carbon dioxide is expected to undergo a phase transition from the supercritical to the sub-critical gaseous phase. During that transition, the carbon dioxide essentially loses its capability to act as a hydrophobic solvent (Schreiber et al. 2012). As a consequence, this leads to the precipitation of mostly hydrophobic products at this point, forming an accumulation zone of corresponding organic compounds in this depth range (Mayer et al. 2017). In the following, we want to report on analytical data obtained from fluid inclusions in the solid core material from 1 km depth.

Results And Discussion

The focus of the analyses was on the detection of aldehydes in the calcite samples of the drill core. Figure 1 shows the result of the calcite sample in comparison to those of the aldehyde standards, the hexane and procedure blanks.

The organic content in calcite samples show at least minor concentrations of aldehydes in the fluid inclusions of the calcite in the core sample. All measurements are triple measurements and the aldehydes were identified with standards. The quantification was made by external calibration (Supplementary Information D and E). Table 1 shows the results of the measurements. Further information such as mass spectra are shown in the Supplementary Information E.

Table 1
Identified aldehydes in fluid inclusions of calcite taken from Wehrer Kessel in Volcanic Eifel, Germany.

| Aldehyde | Chemical Formula | Mass [g/mol] | Rt _{sample} [min] | RSD _{Rt} (n=3) [%] | Rt _{ref} [min] | NIST [%] |
|--------------|-----------------------------------|--------------|----------------------------|-----------------------------|-------------------------|----------|
| Octanal | C ₈ H ₁₆ O | 128.2144 | 7.12 | 0.02 | 7.12 | 90.8 |
| Nonanal | C ₉ H ₁₈ O | 142.2413 | 8.72 | 0.02 | 8.72 | 89.2 |
| Decanal | C ₁₀ H ₂₀ O | 156.2682 | 10.26 | 0.01 | 10.27 | 88.8 |
| Undecanal | C ₁₁ H ₂₂ O | 170.2951 | 11.71 | 0.02 | 11.71 | 90.1 |
| Dodecanal | C ₁₂ H ₂₄ O | 184.3220 | 13.08 | 0.01 | 13.08 | 81.7 |
| Tridecanal | C ₁₃ H ₂₆ O | 198.3449 | 14.37 | 0.01 | 14.37 | 91.2 |
| Tetradecanal | C ₁₄ H ₂₈ O | 212.3715 | 15.59 | 0.01 | 15.59 | 80.7 |
| Pentadecanal | C ₁₅ H ₃₀ O | 226.3981 | 16.74 | 0.07 | 16.74 | 81.7 |
| Hexadecanal | C ₁₆ H ₃₂ O | 240.4247 | 17.83 | 0.08 | 17.84 | 86.0 |

By means of a six-point external calibration procedure, it was possible to determine the concentration of the detectable aldehydes in the calcite sample. Certainly, it has to be taken into account that this is not a representative sample, since the drill core cannot represent a homogeneous environment. One calcite sample was analyzed in three replicates and the concentration of each aldehyde in the sample was calculated between approximately 10 and 600 $\mu\text{g}/\text{kg}$ (details in Table 2). The concentrations of aldehydes in the calcite sample are in the same concentration range as those of the more than 3-billion-years old quartz sample with 433 $\mu\text{g}/\text{kg}$ (Schreiber et al. 2017) and of the aldehydes present in various meteorites (Aponte et al. 2019).

Table 2
Concentration of aldehydes in the fluid inclusions of the calcite sample (more details in Supplementary Table S3).

| Aldehyde | Area | c [$\mu\text{g}/\text{kg}$] | RSD [%] |
|--------------|--------|-------------------------------|---------|
| Octanal | 15220 | 109 | 12.0 |
| Nonanal | 203139 | 582 | 2.1 |
| Decanal | 35050 | 142 | 2.6 |
| Undecanal | 5851 | 23 | 11.2 |
| Dodecanal | 111912 | 362 | 14.2 |
| Tridecanal | 6554 | 31 | 16.7 |
| Tetradecanal | 3420 | 18 | 1.9 |
| Pentadecanal | 6169 | 53 | 3.1 |
| Hexadecanal | 5813 | 36 | 13.3 |

Durham et al. described a supercritical Fischer-Tropsch synthesis using a potassium-promoted iron-based catalyst and supercritical hexane, which allows the production of large amounts of long chain aldehydes with more than ten carbon units (Durham et al. 2014). Whether this reaction is also possible in supercritical CO_2 has not yet been clarified, but in our opinion, it is likely. While alcohols and carboxylic acids were detected in the Fischer-Tropsch type experiments in a hydrothermal system by McCollom et al. (McCollom et al. 1999) with montmorillonite as a common mineral (sodium aluminum silicate), we have detected the intermediate oxidation state in the form of aldehydes (heptanal to heptadecanal) in fluid inclusions of 3-billion-year-old quartz samples (Schreiber et al. 2017) and calcites (octanal to hexadecanal) from a 1000 m deep drill core. These findings are supported by the work of Mißbach et al. (Mißbach et al. 2021), who detected short-chain aldehydes in fluid inclusions of 3.5-billion-year-old barites with a method) which does not allow the analysis of long-chain aldehydes. In addition, aldehydes were also detected in various meteorites in a similar concentration range. Aponte et al. (Aponte et al.

2019), who also used a method, which does not allow the analysis of long-chain aldehydes, were able to detect short-chain aldehydes in various carbonaceous chondrites. However, these results indicate that the aldehydes found in meteorites could possibly also be of extraterrestrial hydrothermal origin.

Just like in the Fischer-Tropsch type experiments of McCollom et al., we detected a lack of an even/odd carbon number predominance and a decreasing abundance of compounds with increasing number of carbon atoms. The experiments conducted by McCollom et al. with and without montmorillonite and in glass as well as new steel vessels led to the conclusion that montmorillonite acts less as a classical catalyst in the Fischer-Tropsch type reaction. Instead, the reaction products were removed from the reaction mixture by adsorption to the montmorillonite, which drives the reaction to the product side according to the Le Châtelier principle (McCollom et al. 1999). This assumption is supported by a work from Iuga and Vivier-Bunge, who investigated the physisorption of small aliphatic aldehydes on a model silicate *Brønsted* site by using quantum chemical methods (Iuga and Vivier-Bunge, 2008). The *Brønsted* site was formed by a large silicate cluster model of 15 Si atoms formed by four six-membered rings. In the central siloxane bridge, one Si atom was replaced by a tetra-coordinated Al atom. The authors identified two different types of physisorption complexes for aldehydes. In both, there is a hydrogen bonding interaction between the hydrogen atom of the acid surface and the aldehydic carbonyl group. The aliphatic chain can be oriented perpendicular or parallel to the surface. In the first case, there is an additional van-der-Waals interaction between an adjacent siloxane bridge and the aldehydic hydrogen atom. In the second case, corresponding interactions occur between siloxane bridges and aliphatic hydrogen atoms of the aldehyde. The results presented in this work suggest that extended silica surfaces with *Brønsted* defects, such as pyrophyllite, can lead to a strong absorption of aldehydes (Iuga and Vivier-Bunge, 2008). In addition, Hakim et al. demonstrated that simple organic molecules, such as propanol or octanoic acid, form ordered adsorption layers on calcite where the layer thickness is defined by their character. They postulated that organic compounds from the surrounding environment would adhere to the surface (Hakim et al. 2017).

The formation of aldehydes is very interesting, since alkylamines can be formed very easily from aldehydes, e.g. via a reaction with ammonia in the presence of formic acid (similar to a Leuckart-Wallach reaction) or alkylcarboxylic acids via an acid-catalyzed reaction. Both ammonia and various acids occur in higher concentrations in hydrothermal fault zones (Schreiber et al. 2012).

Mayer et al. proposed a mechanism of periodic vesicle formation which is expected to occur in fault zones filled by water and CO₂ (Mayer et al. 2015). At a depth of approximately 1 km, pressure and temperature conditions induce a local phase transition between supercritical CO₂ (scCO₂) and subcritical gaseous CO₂ (gCO₂). The presence of amphiphilic compounds, such as alkyl amines or alkyl carboxylic acids, under these conditions inevitably leads to the transient formation of coated water droplets in the gas phase and corresponding vesicular structures in the aqueous environment. The process of periodic formation and destruction of vesicles simultaneously provides a perfect environment for molecular evolution in small compartments and for the emergence of protocells. The basic process of vesicle

formation was experimentally reproduced using a mixture of octadecylamine and octadecanoic acid with a mass ratio of 1:1 in a water/CO₂ system (Mayer et al. 2018).

Conclusion

The investigations of fluid inclusions in calcite samples show the same results as the analyses of the 3 billion years old Archean quartz samples from Western Australia: a homologous series of long-chain aldehydes. Returning to the question of whether the hydrothermal conversion of inorganic (CO, H₂, H₂O) to organic compounds, such as aldehydes, is an ongoing process that continues to occur today, we believe that our results demonstrate that this is in fact the case.

Materials And Methods

The drill core sample

A drill core from the Wehr caldera (N 50° 43', E 7° 22' East Eifel volcanic field, Germany) was taken out of the depth of 950 m to 968 m. The Wehr volcano erupted twice, the most recent eruption took place 150,000 years ago (Bogaard 1989). The eruption formed a collapse structure in the Devonian basement (Wörner 1988), which was filled up with various layers of tephra and sediments. The hydrothermal processes are closely linked to volcanic activity, which is also evident in the Devonian basement. In particular, Quaternary hydrothermal calcite built is of relevant importance here. There are fluid inclusions inside the hydrothermal calcites which have to be examined for organic compounds. For this purpose, samples of hydrothermal calcite from a drilling core were taken. A hydrothermal solution has penetrated into cracks in the fine-grained Devonian rocks (clayish shale, silt stone, fine-grain sandstone) and led to the crystallization of idiomorphic calcites (Figure 2). Further information about the geological framework is shown in the Supplementary Information A.

Sample Preparation

The calcite samples were taken mechanically by fragmentation of the drill core under protected conditions. After the cleaning procedure of the sample surface with hexane, the washed calcite sample was cooled with liquid nitrogen and mortared with a mortar and a pestle and extracted three times with 10 mL hexane. The solution of the grinding step was collected in a Teflon tube. The mortared calcite (powder) was immediately scraped out of the mortar with the pestle and collected in the same Teflon tube. Then it was centrifuged with an Eppendorf Centrifuge 5804 R for 10 min with 3000 rpm at 14°C. The supernatant was enriched with a Büchi Syncore apparatus for seven hours with 200 rpm at 40°C to 1000 µL. Then 500 µL of the 1000 µL sample were taken for the analysis of volatile organic compounds with higher vapour pressure (HVOCs). The residual 500 µL were enriched to dryness and resolved with 100 µL hexane for the analysis of volatile organic compounds with a lower vapour pressure (LVOCs).

Finally, the samples were analyzed with a GC QTOF MS. Further information about the sample preparation is shown in the Supplementary Information B.

Determination Of Organic Compounds By Gc Qtof Ms

GC QTOF MS analyses were performed on a gas chromatograph equipped with a quadrupole time-of-flight gas chromatography (7890B GC/Q-TOF coupled to an accurate mass analyzer (MS 7250), both from Agilent (Santa Clara, CA, United States)). Ionization was performed by electron impact (70 eV). Data Acquisition was performed and processed by MassHunter software version B.08.00 from Agilent Technologies (Santa Clara, CA, United States) and NIST Library (Gaithersburg, MD, United States). Further information about the chromatographic and mass spectrometric setup is listed in the Supplementary Information C.

Declarations

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Methodology: YG OJS

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Figures

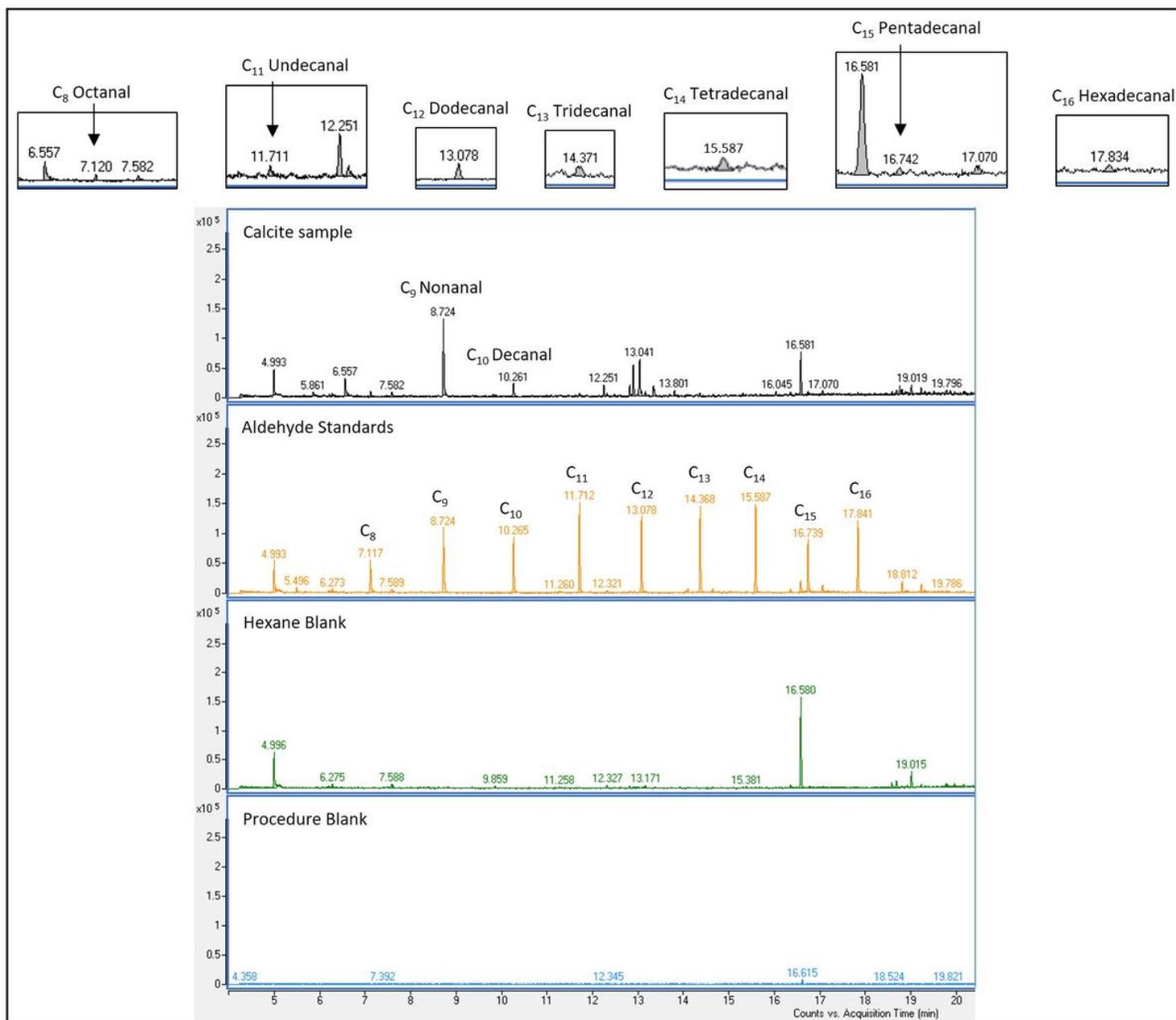


Figure 1

EIC of the $[C_7H_{11}]^+$ aldehyde fragment with m/z 95.0864 of the calcite sample in comparison to the aldehyde standards, hexane and procedure blank.



Figure 2

Light-colored calcite in the fragmented drill core.

Supplementary Files

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